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14. ABSTRACT  
A report of the activity related to the research on the role of defects and electronic structure on the sensitivity of energetic materials to detonation.

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## **First and Second Year Report Period (Dec. 2005 to Dec. 2007)**

This report combines the first two years of the formal grant award period. Because of a Faculty Development Leave for the PI and the associated extreme difficulty in identifying and hiring personnel to work on the project, it was impossible to achieve the planned progress for the first year. This difficulty, as well as the inability to use the funds for the planned instrumentation (since a commercial package could not be obtained), led Dr. Goldwasser and the PI to conclude that a no-cost extension of the grant would be both inevitable and necessary.

### **2006**

The PI left for his scheduled year of Faculty Developmental Leave in January, 2006 beginning with five months as a visiting professor at the University of Milan. The PI returned in June, 2006. The last five months of the year was spent as a Senior Fulbright Fellow at the Wroclaw University of Technology. Because the departure for Milan was, for practical purposes, at the inception of the grant, the search for postdoctoral associates had to be initiated from abroad. A great deal of time was taken in trying to locate suitable postdoctoral associates over that time. However, few with the necessary background were identified and those that were, either had prior commitments, or were not interested in pursuing the position due, in no small part, to visa problems. This proved to be quite time-consuming even though unsuccessful. However, with the graduation of one of the PI's graduate students, who had worked on energetic materials for his PhD, it appeared that it could be possible to make some progress on the research. The individual, J. J. Haycraft, agreed to work on the project. He began to devise an experimental protocol for measurement and introduction of defects in the energetic materials. The initial work was focused on design and construction of an apparatus for injecting defects into the crystals using PZT ceramics. Unfortunately, Dr. Haycraft was made an extremely good job offer arising from a prior application. He left in August for his new position.

During this time, the PI was engaged in working with his collaborator, Prof. Gavezzotti at the University of Milan, on trying to develop a calculational approach to understanding the role of defects in solids of energetic materials. In the course of this work, it was found that the density of crystals of energetic materials are abnormally high compared to crystals of other organic molecular crystals. This was investigated by calculation and it was found that, contrary to commonly held belief, the intermolecular forces in these solids were significantly more strong than the Coulombic interactions. The results of these lattice energy investigations were submitted to the *J. Phys. Chem.* and subsequently published in that Journal. (C. J. Eckhardt, A. Gavezzotti, *J. Phys. Chem. B* **111** (2007) 3430.)

### **2007**

With Dr. Haycraft's departure, the search for the second postdoctoral associate had to be expanded to a search for two associates. Dr. Himansu Mohapatra was hired to continue with the work begun by Dr. Haycraft. In addition, Dr. Mohapatra was charged with the growth of suitable crystals of the less common polymorphs of HMX and RDX. While they could be obtained, they were not of the optical quality needed for the proposed experiments. Dr. Mohapatra spent the bulk of his time in trying to devise suitable methods of growth of the

polymorphs of the two energetic materials. The search continued for the second postdoctoral associate. The search focused more on scientists in the US in order to avoid the problems with foreign workers. Further, it became evident that the criteria for the positions would have to be severely modified. This permitted identification of several potential people. One, in particular, was almost perfectly suited to the position but declined the offer in favor of another. This pattern repeated. However, Dr. Haidong Zhang, with a new PhD from the University of Florida, agreed to join the group in August, 2007. This was particularly welcome since Dr. Zhang had some familiarity with working at synchrotrons. The light source for the proposed vacuum ultraviolet (VUV) spectroreflectometer is to be the CAMD synchrotron light source in Baton Rouge.

Upon arrival, Dr. Zhang began immediately to work on the design of the VUV spectroreflectometer. This is not a simple task since the sample, an organic, is inherently unstable in a vacuum. Nevertheless, as of December, 2007, Dr. Zhang had made very good progress in identifying the problems inherent with this difficult experiment. By December, 2007, he was identifying what components were viable and had design drawings for the instrument. It is anticipated it will be installed in the summer of 2008. Dr. Zhang's appointment runs to August, 2008 and he has agreed to stay another year to complete the research.

Near the end of 2007, a second postdoctoral associate with experience in Brillouin scattering was identified. Dr. Zhando Utegulov, formerly at NIST, agreed to join the project and work on the defect problem. Dr. Mohapatra had indicated that he would rather have another project and he has been able to successfully help Dr. Utegulov move into the defect studies of the proposed research. It is expected that the proposed studies will be underway by the summer of 2008.

Because of the PI's Faculty Development Leave and the extreme difficulty in staffing for this project, it has been effectively delayed by at least a year and perhaps somewhat more. Now, with full staffing, we are desperately trying to make up for lost time. The project is now well underway and significant results are expected by the end of 2008.

\* The PI applied for, and was granted, the Faculty Leave of Absence before the particular program for this ONR research was announced.

## Computer Simulations and Analysis of Structural and Energetic Features of Some Crystalline Energetic Materials

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A database of 43 literature X-ray crystal structure determinations for compounds with known, or possible, energetic properties has been collected along with some sublimation enthalpies. A statistical study of these crystal structures, when compared to a sample of general organic crystals, reveals a population of anomalously short intermolecular oxygen–oxygen separations with an average crystal packing coefficient of 0.77 that differs significantly from 0.70 found for the general population. For the calculation of lattice energies, three atom–atom potential energy schemes and the semiempirical SCDS-PIXEL scheme are compared. The nature of the packing forces in these energetic materials is further analyzed by a study of the dispersive versus Coulombic contributions to overall lattice energies and to molecule–molecule energies in pairs of near neighbors in the crystals, a partitioning made possible by the unique features of the SCDS-PIXEL scheme. It is shown that dispersion forces are stronger than Coulombic forces, contrary to common belief. The low abundance of hydrogen atoms in these molecules, the close oxygen–oxygen contacts, and the high packing coefficients explain the observation that, for these energetic materials, crystal densities are anomalously high compared to those of most organic materials. However, an understanding, not to mention prediction or control, of the deeper mechanisms for the explosive power of these crystalline materials, such as the role of lattice defects, remains beyond present capabilities.

### I. Introduction

Weighing the advantages and dangers of developing more efficient energetic materials is a difficult matter that involves ethics and politics rather than pure science, but it requires due consideration of the fact that the preponderant use of explosives is for non-military applications. There are two overriding concerns for those who use these materials: safety and explosive power. Energetic materials that are resistant to detonation can be handled more safely, although often at the expense of explosive power. These matters are amenable to the usual studies of structure–property relationships, hence the present work.

For decades, explosive power has been correlated with the energy density of the material and, in particular, with the mass density because of the dependence of the magnitude of the detonation velocity and of the detonation sensitivity to this simple physical property.<sup>1</sup> For crystals comprised only of atoms of elements of the first and second periods, the densities of energetic materials tend to be anomalously large. This may have implications regarding defects, since they usually lower the density. Another important measure of explosive power has been the oxygen balance that is taken to be optimal when the amount of oxygen in the energetic material is just sufficient for conversion of all oxygen to only carbon dioxide and water.<sup>2</sup>

Most energetic materials are organic molecules of rather irregular shape, and thus a question of interest is how these

molecules pack in crystals to obtain the most efficient space occupation and hence a high density for a given chemical composition. For such oxygen-rich substances, it would normally be expected that closely packed molecules that have many peripheral oxygen atoms would find difficulty in avoiding close and, presumably, repulsive contacts. A crucial matter then is the nature and magnitude of the intermolecular forces that bind the molecules in these crystals. A related topic is crystal polymorphism, because different polymorphs exhibit varying sensitivity to detonation, thus confirming that solid-state properties must have significant influence on the mechanism of detonation.<sup>1a</sup> A proper understanding of crystal packing forces may help in controlling the growth of favorable polymorphs. Further, the study of “perfect” crystals is a first step in the analysis of the defects which disrupt the crystal lattice and presumably have a significant influence on the performance of these materials.<sup>3</sup>

For these reasons, we have undertaken a detailed packing analysis and examination of the binding forces of a series of crystals of energetic materials by a combination of recently developed statistical techniques and molecular simulation methods.<sup>4</sup> Using crystallographic databases, some geometrical features of the crystal packing of energetic materials are compared with those of a general collection of organic compounds. For the calculation of intermolecular energies, the available empirical atom–atom force fields are critically examined, but the results are also compared with those obtained by more accurate semiempirical and *ab initio* methods. The

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prospects for an accurate and reliable simulation of these materials are then examined.

## II. Experimental Data

**1. Heats of Sublimation.** The sublimation enthalpy,  $\Delta H_{\text{sub}}$ , is a vital piece of information in crystal-structure analysis and force-field calibrations, because its negative should be equal to the calculated lattice energy. The NIST thermodynamic database<sup>5</sup> was searched for thermodynamic properties of the crystalline materials of interest here, in particular for sublimation enthalpies of compounds whose crystal structure also has been determined by X-ray diffraction. The primary literature<sup>6</sup> was also searched independently, and often checked to help in making a choice between different  $\Delta H_{\text{sub}}$  values stored in the NIST database. In particular, the  $\Delta H_{\text{sub}}$  for HMX (1,3,5,7-tetranitro-1,3,5,7-tetraazacyclooctane, refcode OCHTET) has been assessed from a Clausius-Clapeyron treatment of the original data.<sup>6c</sup> The reliability of thermochemical data is somewhat variable, and discrepancies of up to 20–30% among different values for  $\Delta H_{\text{sub}}$  of the same crystal are quite common (for an extreme instance, the reported  $\Delta H_{\text{sub}}$ 's for dimethylnitramine range from 41 to 70 kJ mol<sup>-1</sup>). The experimental data come from different techniques, or often refer to high temperature, as is appropriate for the measurement of vapor pressures of nonvolatile compounds, while calculated lattice energies derive from structures determined at room temperature. Assuming a typical range of values for gas-crystal  $\Delta C_p$ 's for organic crystals of 20–40 J K<sup>-1</sup> mol<sup>-1</sup>, one can estimate a 2–4 kJ mol<sup>-1</sup> increase in the  $\Delta H_{\text{sub}}$  for a 100 K lowering of the temperature. In view of the inherent uncertainties, the comparison between experimental  $\Delta H_{\text{sub}}$ 's and calculated lattice energies is no more than a guideline in the calibration of crystal potentials, and only an overall consideration of trends over many data can be significant; parametrization based on only a few experimental data can sometimes be misleading.

**2. Crystal Structures.** The Cambridge Structural Database<sup>7</sup> (CSD) and the primary literature were searched for X-ray crystal structure determinations of materials with promise for high-energy content, especially molecules containing several nitro groups. The final database of 43 selected crystals (see Tables 1 and 2) is representative rather than exhaustive, but proved to be sufficient for the present purposes. The crystal structure determinations are of variable accuracy as a consequence of changeable crystal quality and the other usual experimental limitations of the X-ray work. This hardly affects the atom–atom force-field calculations, whose potentials are isotropic and radial, and rather insensitive to minor structural variations. However, even a minor inaccuracy in atomic positions significantly affects the ab initio quantum chemical calculations or the PIXEL calculations (see below) that take into account all of the electron density distribution and are much more sensitive to structural details. For example, unrealistic intramolecular energy differences are calculated ab initio for the same molecule extracted from different crystal structure determinations because of geometrical distortions due to different accuracies in atomic positional parameters, which, however small, have large consequences for the intramolecular binding geometry. The effect is less pronounced with intermolecular energies.

## III. Statistical and Theoretical Methods

**1. Atom–Atom Contact Distributions.** The structural database of Table 1 was used to analyze atom–atom contact densities. Distance distribution functions (DDF) are defined as follows:<sup>8</sup> All atom–atom distances between atomic species  $i$

**TABLE 1: Compound Designator Key: Cambridge Structural Database REFCODE, Explosive Acronym (if Extant), and Common Chemical Name**

CATJIQ	1,3,3,5,7,7-hexanitro-1,5-diazacyclo-octane
CIWMEA	TNAZ, 1,3,3-trinitroazetidine
CTMTNA	RDX, cyclotrimethylene-trinitramine
DATNBZ	1,3,5-trinitro-2,4-diaminobenzene
DIMNAN	N,N-dimethyl-4-nitroaniline
DNBENZ	1,3-dinitrobenzene
DNITBZ	1,4-dinitrobenzene
DNPMTA	dinitrophenatmethylenetetramine
GEMZAZ	1,4-dinitroglycoluril
GIMBOT	2,2,4,4,6,6-hexanitrostilbene
HIHHAH	3-nitrobenzaldehyde
JEDSUG	1,3-dinitro-1,3-diazacyclopentan-2-one
JEHLAJ	2,5,7,9-tetranitro-8-oxo-2,5,7,9-tetra-azabicyclo(4.3.0)nonane
JEXLUT	cis-1,3,5,7-tetranitro-1,3,5,7-tetra-azadecalin
JEXMAA	TNBI 1,1,3,3-Tetranitro-4,4-bi-imidazolidine
JEXMEE	TNSD, 1,3,7,9-Tetranitro-1,3,7,9-tetra-azaspiro(4.5)decane
JEXMII	trans-1,4,5,8-tetranitro-1,4,5,8-tetra-azadecalin
KEMTIF	2,4,8,10-tetranitro-2,4,8,10-tetra-azaspiro(5.5)undecane
KOFKAR	1,3-dinitro-1,3-diazacyclohexane
KOFKEV	1,3-dinitro-1,3-diazacycloheptane
KOFKIZ	1,5-dinitro-3-nitroso-1,3,5-triazacycloheptane
METNAM	N,N-dimethylnitramine
MTNANL	TETRYL, 2,4,6-trinitro-N-methyl-N-nitroaniline
MNTDMA	N,N-dimethyl-3-nitroaniline
NACPON	N,N-dimethyl-4-nitroaniline
NITOLU	4-nitrotoluene
NOHTET	HMX, 1,3,5,7-tetranitro-1,3,5,7-tetra-azacyclo-octane
NOHTAZ	2,4,6-trimethyl-1,3,5-trinitrohexahydro-1,3,5-triazine
PUBMUU	HNIW 2,4,6,8,10,12-hexanitro-2,4,6,8,10,12-hexa-azatetracyclododecane, hexanitrohexa-azaisowurtzitane
SECVOL	2,2-bis(1,3-dinitrohexahydropyrimidine)
TATNBZ	1,3,5-triamino-2,4,6-trinitrobenzene
TNBENZ	1,3,5-trinitrobenzene
TNIOAN	1,3,5-trinitroaniline
TNOXYL	1,3,5-trinitro-2,4-dimethylbenzene
TNPHNT	1-ethoxy-2,4,6-trinitrobenzene
ZZZFYW	1,2-dinitrobenzene
ZZZMUC	2,4,6-trinitrotoluene
ZZZQSC	2,4-dinitrotoluene

and  $j$  in all crystal structures in the database are calculated up to a certain limit  $R_{\text{max}}$ . Let  $N_k(R)^{ij}$  be the number of atom–atom distances within the  $k$ th distance bin between  $R_k$  and  $R_k + dR$ . An inspection of the unnormalized distribution function reveals a separation below which no contacts are observed,  $R^{oij}$ . The normalization factor,  $F_N$ , and the DDF are then

$$F_{N,ij} = \sum N_k(R)^{ij} / [4\pi/3((R_{\text{max}})^3 - (R^{oij})^3)]$$

$$\text{DDF} = g_k(R)^{ij} = (1/F_{N,ij})N_k(R)^{ij}/dV_k$$

where  $dV_k$  is the volume of the spherical shell.  $F_N$  represents the condition of uniform distribution of the observed contacts over the available contact space. The DDF is formally similar to a radial distribution function, RDF, but the physicochemical meaning is different, because a RDF refers to a sampling over the same molecular species in a homogeneous system, while the DDF refers to sampling over different molecules and over many different systems, the crystal structures.

**2. Ab Initio Calculations.** Quantum chemical calculations on isolated molecules with a fixed geometry as extracted from the crystal structure determination were carried out using the GAUSSIAN package,<sup>9</sup> at the MP2 6-31G\*\* level with the “cube” option to produce an electron density for PIXEL (step of 0.08 Å), and the “pop = esp” option for the calculation of

TABLE 2: Crystal Lattice Energies (kJ mol<sup>-1</sup>)

CSD refcode	-E(UNI) <sup>a</sup>	ΔH(subl) <sup>b</sup>	-E(PIXEL) <sup>c</sup>	-E(SRT) <sup>d</sup>	E(coul+pol) <sup>e</sup>	E(displ) <sup>f</sup>
CATJQ	152			183		
CIWMEA	93		84	113	-61	-92
CTMTNA	109	112	116	130	-87	-97
DATNBZ	154	140	116		-75	-121
DIMNAN	108	103	105		-60	-118
DNBENZ	101	85	80		-40	-77
DNITBZ	105	96	89		-49	-80
DNPMTA	130		130	141	-87	-107
GEMZAZ	148		155		-157	-125
GIMBOT	205	180				
HIHHAH	101	110	88		-49	-85
JEDSUG	93		112	124	-86	-81
JEHLAJ	136			181		
JEXLUT	159			179		
JEXMAA	154			176		
JEXMEE	147			173		
JEXMII	155			169		
KEMTIF	160			184		
KOFKAR	101		101	117	-68	-82
KOFKEV	108		106	120	-67	-92
KOFKIZ	117			131		
METNAM	59	70	72	70	-41	-56
MTNANL	134	134	118	150	-77	-106
MNTDMA	104	93	94		-49	-105
NACPON	100					
NITOLU	86	79	71		-30	-73
NOHTAZ	127			141		
OCHTET00	127			179		
OCHTET03	124			168		
OCHTET12	146	162	157	180	-91	-127
PUBMUU00	156			175		
PUBMUU01	165			181		
PUBMUU02	158			187		
SECVOL	152			188		
TATNBZ	190	168	156		-113	-137
TNBENZ10	105	107				
TNBENZ13	107	107	86		-54	-94
TNIOAN	132	120	100		-62	-97
TNOXYL	132	130	103		-58	-104
TNPHT	119	121	104		-63	-100
ZZZFYW	101	87	88		-46	-76
ZZZMUC	119	105				
ZZZQSC	108	98	89		-46	-91

<sup>a</sup> Lattice energy with UNI atom-atom force field. <sup>b</sup> Heats of sublimation. <sup>c</sup> Total PIXEL lattice energy. <sup>d</sup> Lattice energy reported in ref 17, with atomic charges from highest-level quantum chemical calculation. <sup>e</sup> Sum of Coulombic and polarization PIXEL terms. <sup>f</sup> PIXEL dispersion energy.

electrostatic-potential-derived atomic point charges. Alternatively, these were also calculated by the so-called rescaled-EHT method,<sup>10</sup> based on a Mulliken population analysis on an Extended Hückel calculation with modified valence-orbital ionization potentials. This produces atomic point charges very similar to those from a Mulliken population analysis on the MP2 wavefunction, and can be routinely applied even to very large molecular systems because of the virtually null computational requirements of an EHT calculation versus a full molecular orbital MP2 calculation. These atomic point charges were used only for the calculation of cell dipoles in the convergence correction, or for comparison with PIXEL Coulombic energies which include penetration energies (see below).

### 3. Lattice Energies and Molecule-Molecule Energies.

Crystal structures were taken as is from the X-ray diffraction determination without optimization of lattice energies or structure relaxation, but hydrogen atom positions were recalculated as usual.<sup>10</sup> Force-field calculations were carried out in the atom-atom approximation with the chargeless UNI force field,<sup>10</sup> which uses a simple  $A \exp(-BR) - CR^{-6}$  functional form in the atom-atom distance  $R$ . Point-charge Coulombic terms were then evaluated separately using atomic point-charge parameters obtained as described above. For more accurate

calculations, the SCDS-PIXEL approach<sup>11</sup> was applied. The molecular electron density is first calculated by standard quantum-chemical methods giving the electron distribution on a large number (10 000–20 000 for our molecules) of charge pixels. The Coulombic energy is then calculated by sums over pixel-pixel, pixel-nucleus, and nucleus-nucleus Coulombic terms. A local polarizability is then assigned to each pixel, the electric field generated by pixels and nuclei in surrounding molecules is calculated, and the linear polarization energy is evaluated. An empirical damping function, using one disposable parameter, is introduced to avoid singularities. The overlap between molecular densities is calculated, and the exchange repulsion energy is evaluated as proportional to the overlap integral. Dispersion energies between two molecules A and B are calculated as a sum of pixel-pixel terms in a London-type expression involving the above-defined distributed polarizabilities and an "oscillator strength" energy,  $E_{OS}$ :

$$E_{DISP,AB} = (-3/4) \sum_{i,A} \sum_{j,B} E_{OS,i} f(R) \alpha_i \alpha_j / [(4\pi\epsilon^0)^2 (R_{ij})^6]$$

where the damping function is  $f(R) = \exp[-(D/R_{ij} - 1)^2]$  (for  $R_{ij} < D$ ), where  $D$  is an adjustable empirical parameter. Standard

PIXEL theory uses  $D = 3.0 \text{ \AA}$ , but a preliminary survey showed that all lattice energies of crystals considered here were underestimated with respect to the  $\Delta H_{\text{sub}}$ . Since the parameter is fully adjustable and the effect is systematic, a correction of this parameter to reduce the discrepancy was considered acceptable.  $D = 2.6 \text{ \AA}$  is used here, which increases the calculated dispersion energies by 5–10% with respect to the standard parametrization.

$E_{\text{OS}}$  may be approximated by the molecular ionization potential for small molecules or as the energy of the highest occupied molecular orbital (HOMO), since the interacting electrons are peripheral ones and hence are roughly at the HOMO energy level. A more refined approach considers each pixel as a separate oscillator, with a formal ionization potential  $I_i$ , which in turn is a function of the ionization potential,  $I^\circ$ , pertaining to the atom to whose basin the pixel belongs, and of the distance between the pixel and the atomic nucleus,  $R_i$ :

$$E_{\text{OS}} = (I_i I^\circ)^{1/2}$$

$$I_i = I^\circ \exp(-\beta R_i)$$

The parameter  $\beta$  is a function of the atom type. This “variable-ionization” form of the theory amounts, in fact, to using different dispersion energy coefficients according to the different kinds of interacting atomic basins, and yields more accurate lattice energies in organic crystals.<sup>8</sup>

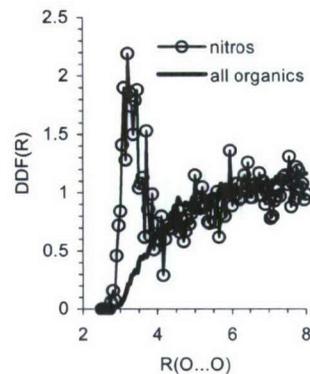
PIXEL calculations were applied to obtain separate Coulombic, polarization, dispersion, and repulsion contributions to the lattice energies, except for very large molecules, where the calculation of electron densities is too expensive, or for crystal structures with more than one molecule in the asymmetric unit, due to a technical difficulty with the evaluation of polarization contributions. For crystal structures in polar space groups, when the angle between the molecular dipole direction and a polar axis is small, a convergence correction should be applied to the lattice sums using the van Eijck–Kroon method;<sup>12</sup> for the very few cases when they were needed, however, these corrections were around 1–2 kJ mol<sup>-1</sup>, so they were considered to be within the computational noise and were neglected. More information on separate driving forces in crystal packing (the “structure determinant” packing analysis) was then obtained from PIXEL interaction energies between near-neighbor pairs of molecules in the crystal.

#### IV. Results and Discussion

##### 1. Density Distribution Functions and Packing Efficiency.

Figure 1 shows that in the database of nitro-derivative crystal structures, the oxygen–oxygen intermolecular distance distribution function (DDF) shows a well-developed peak at about  $R(\text{O}\cdots\text{O}) = 3.2 \text{ \AA}$ , a peak that does not appear when the same survey is conducted on a sample of ~30 000 general organic crystals. We initially interpreted this unexpected fact as meaning that, due to the relevant Coulombic interactions present in these crystals, particularly those due to stabilizing interaction between nitro-nitrogen sites and nitro-oxygen sites, the peripheral oxygen atoms are forced into a close intermolecular contact that is counterintuitive in terms of simple atom–atom Coulombic interactions. The explanation actually requires a subtler analysis including penetration energies, i.e., those attractive energies which arise from overlap of diffuse electron clouds and their proximity to positively charged nuclei in the intermolecular partner.

A similar, but considerably smaller, peak is observed in the DDF of  $\text{O}\cdots\text{H}$  distances. This is not unexpected in view of the

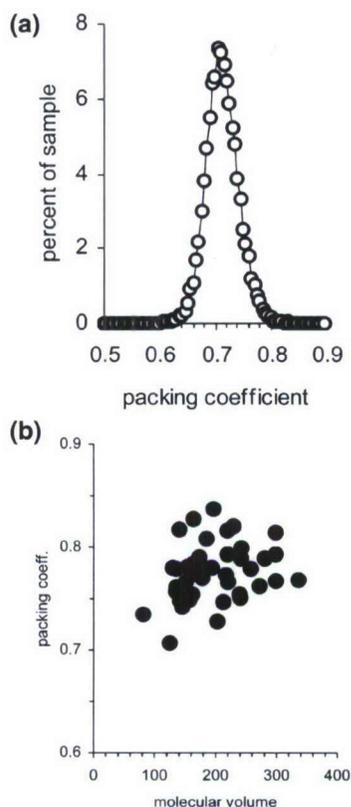


**Figure 1.** The density distribution function of  $\text{O}\cdots\text{O}$  separations in the sample of crystals of energetic materials of Table 1 (“nitros” line) and in a sample of 27 461 crystals of middle-size, general organic molecules containing only C, H, N, O, and halogen atoms, from the Cambridge Structural Database.

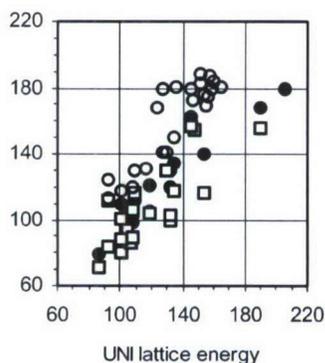
opposite sign of the molecular potentials around hydrogen atoms and oxygen atoms in nitro compounds, and a similar trend had been previously noted in a systematic study of nitrobenzenes.<sup>13</sup> These results show the difficulty of defining intermolecular bonds in terms of short atom–atom intermolecular distances. Thus the postulate of existence of a  $\text{C}-\text{H}\cdots\text{O}$  crystal bond on the basis of atom–atom contacts alone requires explanation why the same analysis does not lead to the postulate of some kind of  $\text{O}\cdots\text{O}$  “bond” which would be in conflict with elementary chemical reasoning. Otherwise, supplementary reasoning must be produced for the distinction between the two cases, an alternative that is also not straightforward.

The energetic performance of materials is often related to high crystal density. This property stems from two unrelated factors: (a) molecular mass, or the number of protons and neutrons and (b) intermolecular close packing, related to electronic interactions. Overall,<sup>14</sup> crystal density is determined mainly by molecular mass, an obvious property. For the design of materials for better transmission of heat and mechanical energy, factor b above, control of the relative efficiency by which space is occupied by constituent molecules of identical stoichiometry, is more challenging. A comparison between the packing coefficients<sup>15</sup> for the present database and for a sample of general organic crystals (Figure 2), in spite of the large difference in statistical significance due to the large difference in sample size, is convincing in this respect. This observation is undoubtedly related to the relatively low number of hydrogen atoms in these molecules, which permits a higher interaction-energy density and parallels the previous observation of an unusual distribution of short contacts between peripheral oxygen atoms

**2. Lattice Energies.** An extensive set of lattice energy calculations using an empirical atom–atom force field for energetic materials with high quality ab initio atomic point charges has been carried out by Sorescu, Rice and Thompson (SRT).<sup>16</sup> We use the lattice energy calculated by the universal UNI force field<sup>10</sup> as a guideline, and Figure 3 plots the results obtained by the PIXEL scheme and by the SRT scheme, as well as the  $\Delta H_{\text{sub}}$ . These data are collected in detail in Table 2. The best apparent agreement between experiment and calculation is that of the UNI potentials, with SRT potentials overestimating the lattice energies and PIXEL values showing an acceptable performance except for a few patent outliers (TNOXYL, TNIOAN, DATNBZ, HHHHAH, see also Table 2) where the calculated lattice energy is too small. One should recall, however, that the comparison is hampered on the one hand by



**Figure 2.** (a) Histogram of the percent distribution of packing coefficients in the general sample (see caption to Figure 1). The peak value is 0.70. (b) The distribution of the packing coefficients in the sample of crystals of energetic materials of Table 1 with average value of 0.77.



**Figure 3.** The UNI force-field lattice energy is the abscissa: on the ordinate are the sublimation enthalpies (black dots), the PIXEL lattice energies (squares), and the SRT lattice energies (circles); see Table 2 for the definitions. Units in  $\text{kJ mol}^{-1}$ .

the uncertainties of experimental values, as previously discussed, and, on the other hand, by the neglect of intramolecular rearrangement energies on going from the crystal to the gas-phase conformation. Accordingly, the overestimation of rigid-body lattice energies might be reasonable, where the correct sublimation enthalpy results after subtraction of the intramolecular energies.

The issue of accurate reproduction of experimental thermochemical data is complex, as previously discussed, but we argue that the added value of the PIXEL method consists in the energy partitioning (Table 2), which offers some insight into the nature of the energetic texture of crystalline materials not apparent in atom-atom schemes. PIXEL Coulombic energies are parameter-

free and are as accurate as the wavefunction is, while PIXEL dispersion energies depend, to some extent, on the parametrization. However, since Coulombic energies are exact and total lattice energies are quite comparable to experimental heats of sublimation, the dispersion contribution cannot be grossly off the mark. Therefore, the message in Table 2 is unequivocal: all lattice energies have a substantial Coulombic-polarization (what is commonly called "Coulombic") contribution, but dispersion energies are invariably equal to, or larger than, the Coulombic ones. This result, together with the previous analysis of oxygen-oxygen contacts, issues a warning against modeling schemes that rely too much on Coulombic energies, especially when they are estimated by atomic point-charge schemes, even for the compounds considered here that are commonly thought to possess a strong permanent polarization of their electron clouds.

**3. Recognition Modes.** Significant, further information on packing modes and packing forces in organic crystals comes from a calculation of interaction energies between pairs of close neighbor molecules in the PIXEL approach, the "structure determinant" analysis.<sup>17</sup> To do this, one reference molecule is chosen and crystal symmetry is used to generate those molecules whose centers of mass are closest to the center-of-mass of the reference molecule. Then, each of these coordinated molecules is designated by (i) the distance between its center-of-mass and the reference molecule, (ii) the label of the symmetry operation connecting the two molecules, and (iii) the molecule-molecule interaction energy calculated by PIXEL and apportioned according to its various contributions: Coulombic, dispersion, and repulsion. A set of these geometrical and energetic values for a molecular pair is referred to as a "structure determinant," and a set of structure determinants is an unequivocal and accurate fingerprint of a crystal structure.

Data from this procedure are more informative than overall lattice energies because the PIXEL energy partitioning quantitatively clarifies the relative importance of Coulombic and dispersion terms, and establishes a more immediate correlation between packing forces and molecular electronic structure. Table 3 collects these lists for three representative crystals, and Figures 4-6 show the corresponding structural diagrams.

In the methyl nitramine crystal, the two determinants along the screw axis, *A* and *C*, roughly correspond to antiparallel nitro groups, but an interpretation based only on opposing dipoles is not warranted, given that the Coulombic and dispersion contributions are roughly equivalent. The nitro-methyl *B* determinant has a predominant Coulombic character. The next three determinants complete the first coordination shell made of 12 nearest-neighbors as in the close packing of spheroids (recall that each *S* or *T* interaction has two partners to the reference molecule). They are rather weak and sometimes have a destabilizing Coulombic component, like the *E* interaction which presents a confrontation of methyl groups, but are stabilized by a counterbalancing dispersion term.

In the CTMTNA crystal, the strongest binding comes from an interaction over a centrosymmetric molecular pair with the nitro-oxygens in one partner pointing to the nitro-nitrogen in the other. Here the Coulombic term is strongly stabilizing, as revealed by the PIXEL calculation, in spite of the unavoidable proximity of oxygen atoms. Notice, however, that in this case, too, the predominating factor is dispersion. It would have been quite difficult to properly describe this dominating interaction in pure atom-atom terms, and only the PIXEL description is adequate. The next determinant, *B*, is a clear  $\text{O}\cdots\text{H}$  Coulombic interaction. However, it becomes increasingly difficult to assign

TABLE 3: Structure Determinants

determinant label,	$R(\text{c.o.m.}),$							approximate
	symmetry <sup>a</sup>	$E_{\text{coul}}^b$	$E_{\text{pol}}$	$E_{\text{disp}}$	$E_{\text{rep}}$	$E_{\text{tot}}$		description
METNAM								
A	3.86 S	-17	-3	-16	+7	-29		nitronitro antiparallel
B	6.13 T <sub>z</sub>	-18	-4	-9	+9	-21		nitro-methyl, double
C	5.36 S	-10	-3	-10	+6	-16		nitronitro antiparallel, offset
D	6.58 T <sub>x</sub>	-3	-1	-5	+3	-7		nitro-methyl, single
E	5.75 S	+3	-1	-10	+3	-5		methylmethyl offset
F	6.07 T <sub>xz</sub>	+4	-1	-8	+3	-1		nonspecific
CTMTNA								
A	4.41 I	-17	-5	-38	33	-28		nitro cups
B	7.29 G <sub>y</sub>	-18	-5	-17	14	-26		nitro-methylene
C	7.29 S <sub>z</sub>	-17	-3	-12	7	-25		adjacent cups
D	6.94 G <sub>x</sub>	-8	-4	-16	13	-15		oxygen-nitrogen
E	6.55 S <sub>y</sub>	-6	-2	-10	6	-12		nonspecific dispersion
F	6.45 G <sub>z</sub>	-5	-4	-23	19	-12		nonspecific dispersion
HMX								
$\beta$ A	6.54 T <sub>x</sub>	-37	-11	-36	29	-55		nested contact
$\alpha$ A	5.91 T <sub>z</sub>	-41	-11	-27	29	-51		triple O...H contacts
$\beta$ C	7.76 S	-13	-4	-20	13	-24		nonspecific
$\beta$ B	7.37 T <sub>xz</sub>	-10	-5	-20	12	-23		side contact
$\beta$ D	7.03 S	-11	-4	-18	15	-17		single O...H contact
$\alpha$ B	8.13 G	-9	-5	-23	24	-13		nonspecific dispersion
$\alpha$ C	7.22 -	+2	-7	-23	19	-10		nonspecific dispersion
$\beta$ E	10.9 T <sub>xz</sub>	+5	-1	-3	1	+2		destabilizing

<sup>a</sup> Label as in Figures 4–6, distance between centers of mass (com) (Å); label of the symmetry operator: S = twofold screw, T = translation, I = inversion center, G = glide plane. <sup>b</sup> PIXEL partitioned energies and total molecule–molecule energy, kJ mol<sup>-1</sup>.

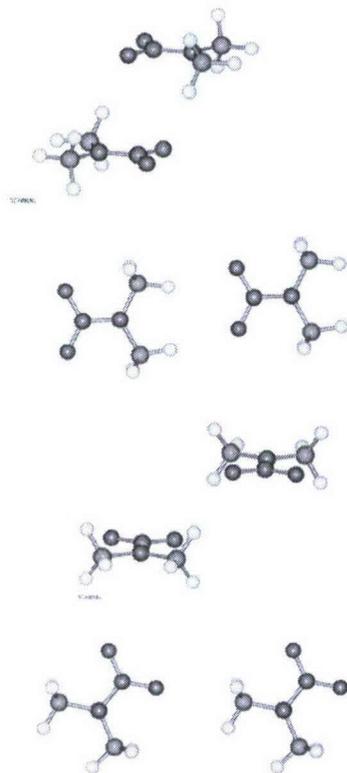


Figure 4. The structure determinants in the N,N-dimethylnitramine (METNAM) crystal structure. Top to bottom: A, B, C, and D (same labels as in Table 3).

an unequivocal label to the subsequent determinants, which comprise broad and diffuse mixtures of Coulombic and dispersion terms. The presence of these nonspecific structure deter-

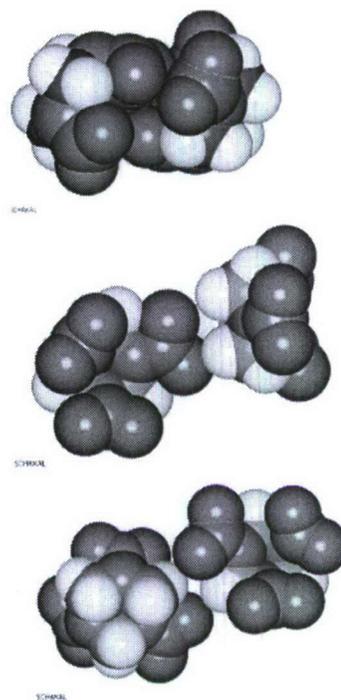


Figure 5. The structure determinants in the cyclotrimethylenetrinitramine (CTMTNA) crystal structure. Top to bottom: A, B, and C (same labels as in Table 3). Dark caps, nitro groups; white caps, hydrogen atoms.

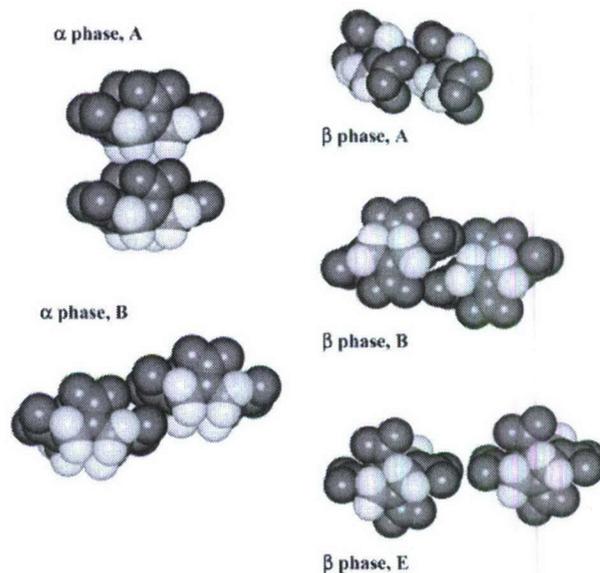
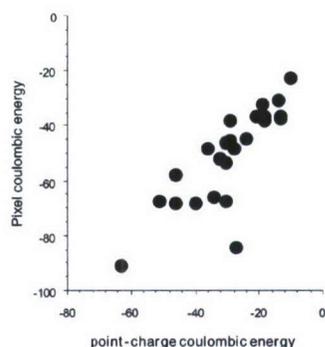


Figure 6. The structure determinants in the  $\alpha$  phase HMX (OCHTET00) and  $\beta$  phase HMX (OCHTET12). Same labels as in Table 3. Dark caps, nitro groups; white caps, hydrogen atoms.

minants is what makes crystal-structure analysis, not to mention prediction and control, so difficult in terms of intermolecular atom–atom bonding.

The same analysis on the A and D molecular pairs in the CTMTNA crystal has been carried out<sup>18</sup> with a methodology that in many respects resembles the PIXEL approach, using a HF/6-311G\*\* wavefunction. The use of HF instead of MP2 leads to higher Coulombic energies, as is well-known. On the other hand, a comparison between the results of the two methods at the same wavefunction level (HF) gives very close values



**Figure 7.** PIXEL exact Coulombic energy vs approximate point-charge Coulombic energy for the crystal structures listed in Table 2.  $\text{kJ mol}^{-1}$  units.

for the Coulombic + polarization term ( $-35.2$  and  $-16.8$ , PIXEL,  $-37.2$  and  $-18.8$ , ref 18).

The case with HMX is even more complex. Notice how the same molecule in two largely different conformations can build up crystal structure determinants of nearly equal intensity in two polymorphic crystals: the  $\alpha$  form by a clear confrontation of its nitro and methylene moieties with Coulombic stabilization ( $\alpha A$  determinant), and the  $\beta$  form by a less clear-cut "nested" contact ( $\beta A$  determinant), with an equal amount of Coulombic and dispersive character. In the  $\beta$  form, a near-neighbor pair ( $\beta E$ ) displays a sharp confrontation of opposing nitro groups where the distance prevents an overlap of the molecular electron densities and no penetration energy sets in. The resulting Coulombic interaction energy is mildly destabilizing, but a minor amount of dispersion from the overall molecular bodies is more than enough to wipe out the repulsion. Thus, the PIXEL analysis reveals the presence of nonspecific stabilizing or even "silent", slightly destabilizing partners whose role in crystal packing is difficult to assess but is by no means negligible, although they are usually neglected in standard crystal packing analyses based on short interatomic distances only.

**4. Point-Charge versus PIXEL Coulombic Energies.** A comparison between PIXEL Coulombic lattice energies and point-charge Coulombic energies is instructive. Figure 7 shows the result: the good news is that there is a rather consistent proportionality ratio by which the point-charge energy is about 60% of the real Coulombic energy. This could be useful in the rescaling of empirical, force-field parametrizations. There are, however, some patent outliers where the point-charge model is severely off the mark. The case of triamino-trinitrobenzene, with its extensive hydrogen-bonding network, is exemplary in this respect.

## V. Conclusions

The quantitative analysis of structures and interaction energies in crystals of energetic compounds is in many respects revealing. With their high number of nitro groups, these molecules have a low relative abundance of hydrogen atoms and in their crystals they must make do with an unusually high number of contacts between peripheral oxygen atoms. These are indeed numerous and anomalously short, as clearly appears in a quantitative study based on the newly introduced concept of Density Distribution Functions (DDF's). Packing coefficient values are unusually high, most likely as a consequence of a higher interaction energy density. Calculated lattice energies compare rather favorably with experimental sublimation enthalpies in all of the computational schemes employed here. The comparison is, however, much less compelling than is usually thought in molecular

simulation and force-field calibrations because of the large uncertainties in experimental values and because of the many approximations, especially the rigid-molecule approximation, which is particularly restrictive for many of the compounds examined here. The PIXEL analysis has an average performance in regard to total lattice energies but is superior in its interpretative power because it is based on complete electron distributions rather than on a few, localized, interaction centers. The PIXEL analysis reveals that (i) short oxygen–oxygen separations do not imply repulsive or destabilizing interactions, a probable conclusion arising from a superficial analysis in terms of atom–atom contacts, and (ii) dispersive contributions to the interaction energies are invariably higher, contrary to common belief.

The studies presented here use molecular simulations based on a perfect model, an infinite, homogeneous crystal without defects. Whether this model is adequate or not depends on which crystal properties are studied. We suggest that our results are significant concerning the general conclusions regarding structural features of these crystals: atom–atom distance distributions and packing coefficients, and the general nature of the crystal packing forces, especially their subdivision into Coulombic and dispersive components.

In any case, simulation of perfect crystals is expected to provide a paradigm by which the effect of imperfections can be discussed. Moreover, molecule–molecule interaction energies do not depend on a global model of the bulk crystal and are therefore certainly informative and reliable. A proper consideration of the role of defects will certainly be needed when tackling a molecular simulation for a more detailed investigation of subtle factors which determine the properties related to reactivity, decomposition and mass or heat transport of these materials. This is left for future work.

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**Supporting Information Available:** Atomic coordinate files, including renormalized hydrogen atom positions, for all the crystal structures considered. This information is available free of charge via the Internet at <http://pubs.acs.org>.

## References and Notes

- (1) (a) Freid, L. E.; Manaa, M. R.; Pagoria, P. F.; Simpson, R. L. *Annu. Rev. Mater. Res.* **2001**, *31*, 291–321. (b) Borne, L.; Patedoye, J.-C.; Spycykerelle, C. *Propellants Explos. Pyrotechn.* **1999**, *24*, 255–259.
- (2) Lothrop, W. C.; Handric, G. R. *Chem. Rev.* **1949**, *44*, 419–445.
- (3) (a) Hatano, T. *Phys. Rev. Lett.* **2004**, *92*, 085501/1–4. (b) van der Heijden, A. E. D. M.; Bouma, R. H. B. *Cryst. Growth Des.* **2004**, *4*, 999–1007. (c) Armstrong, R. W.; Ammon, H. L.; Elban, W. L.; Tsai, D. H. *Thermochem. Acta* **2002**, *384*, 303–313. (d) Kuklja, M. M.; Kunz, B. A. *J. Phys. Chem.* **1999**, *B103*, 8427–8431.
- (4) Gavezzotti, A. *Molecular Aggregation*; Oxford University Press, Oxford, 2006. Computer codes for the software used in this paper and reference manuals can be downloaded from <http://users.unimi.it/gavezzotti> as supplementary material to the book.
- (5) Afeefy, H. Y.; Liebman, J. F.; Stein, S. E. Neutral Thermochemical Data. In *NIST Chemistry WebBook*; NIST Standard Reference Database Number 69; Linstrom, P. J., Mallard, W. G., Eds.; National Institute of Standards and Technology: Gaithersburg, MD, 2003 (<http://webbook.nist.gov/chemistry>). Chickos, J. S. Heat of Sublimation Data. In *NIST Chemistry WebBook*; NIST Standard Reference Database Number 69; Linstrom, P. J., Mallard, W. G., Eds.; National Institute of Standards and Technology: Gaithersburg, MD, 2003 (<http://webbook.nist.gov/chemistry>).
- (6) (a) Cundall, R. B.; Palmer, T. F.; Wood, C. E. C. *J. Chem. Soc., Faraday Trans. 1* **1978**, *43*, 1339–1345. (b) Lyman, J. L.; Liau, Y.-C.;

Brand, H. V. *Combust. Flame* **2002**, *130*, 185–203. (c) Taylor, J. W.; Crookes, R. J. *Combust. Flame* **1973**, *21*, 391–393.

(7) Allen, F. H.; Kennard, O. *Chem. Des. Automation News* **1993**, *8*, 31–37.

(8) See ref 4, Chapter 12.

(9) Frisch, M. J.; Trucks, G. W.; Schlegel, H. B.; Scuseria, G. E.; Robb, M. A.; Cheeseman, J. R.; Montgomery, J. A., Jr.; Vreven, T.; Kudin, K. N.; Burant, J.C.; Millam, J. M.; Iyengar, S. S.; Tomasi, J.; Barone, V.; Mennucci, B.; Cossi, M.; Scalmani, G.; Rega, N.; Petersson, G. A.; Nakatsuji, H.; Hada, M.; Ehara, M.; Toyota, K.; Fukuda, R.; Hasegawa, J.; Ishida, M.; Nakajima, T.; Honda, Y.; Kitao, O.; Nakai, H.; Kiene, M.; Li, X.; Knox, J. E.; Hratchian, H. P.; Cross, J. B.; Adamo, C.; Jaramillo, J.; Gomperts, R.; Stratmann, R. E.; Yazyev, O.; Austin, A. J.; Cammi, R.; Pomelli, C.; Ochterski, J. W.; Ayala, P. Y.; Morokuma, K.; Voth, G. A.; Salvador, P.; Dannenberg, J. J.; Zakrzewski, V. G.; Dapprich, S.; Daniels, A. D.; Strain, M. C.; Farkas, O.; Malick, D. K.; Rabuck, A. D.; Raghavachari, K.; Foresman, J. B.; Ortiz, J.V.; Cui, Q.; Baboul, A. G.; Clifford, S.; Cioslowski, J.; Stefanov, B. B.; Liu, G.; Liashenko, A.; Piskorz, P.; Komaromi, I.; Martin, R. L.; Fox, D. J.; Keith, T.; Al-Laham, M. A.; Peng, C. Y.; Nanayakkara, A.; Challacombe, M.; Gill, P. M. W.; Johnson, B.; Chen, W.; Wong, M. W.; Gonzalez, C.; Pople, J. A. *Gaussian 03*, Revision A.1; Gaussian, Inc.: Pittsburgh, PA, 2003.

(10) Gavezzotti, A.; Filippini, G. *J. Phys. Chem.* **1994**, *98*, 4831–4837.

(11) (a) Gavezzotti, A. *J. Phys. Chem. B* **2002**, *106*, 4145–4154. (b) Gavezzotti, A. *J. Phys. Chem. B* **2003**, *107*, 2344–2353. (c) Gavezzotti, A. *J. Chem. Theor. Comp.* **2005**, *1*, 834–840. See ref 5, Chapter 12, for more details.

(12) van Eijck, B. P.; Kroon, J. *J. Phys. Chem.* **1997**, *B101*, 1096–1100.

(13) Demartin, F.; Filippini, G.; Gavezzotti, A.; Rizzato, S. *Acta Crystallogr., Sect. B* **2004**, *60*, 609–620.

(14) Dunitz, J. D.; Filippini, G.; Gavezzotti, A. *Tetrahedron* **2000**, *56*, 6595–6601.

(15) The packing coefficient is defined as the ratio of molecular volume to cell volume. These were calculated by numerical integration rather than by the Kitaigorodski spheres-and caps method (Kitaigorodski, A. I. *Organic Chemical Crystallography*, Consultants Bureau, New York 1961) which is rather inaccurate for molecules with nitro groups.

(16) Sorescu, D. C.; Rice, B. M.; Thompson, D. L. *J. Phys. Chem. A* **1998**, *102*, 8386–8392.

(17) Gavezzotti, A.; Filippini, G. *J. Am. Chem. Soc.* **1995**, *117*, 12299–12305.

(18) Politzer, P.; Ma, Y. *Int. J. Quantum Chem.* **2004**, *100*, 733–739.

(19) Keller, E. *SCHAKAL92, A Program for the Graphic Representation of Molecular and Crystallographic Models*; University of Freiburg: Freiburg, 1993.