

Final Technical Report

Growth of Single Crystals and Fabrication of GaN and AlN Wafers

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ABSTRACT

Gallium nitride (GaN) thin films were grown on (0001) sapphire substrates at 1050°C by controlled evaporation of gallium (Ga) metal and reaction with ammonia NH₃. The feasibility of the growth process was demonstrated and discussed. One of the biggest challenges of working in the Ga–NH₃ system was the instability of molten Ga under NH₃ atmosphere at elevated temperatures, especially between 1100–1200°C.

In the first part of the study, transport of Ga species from the source-to-substrate during the GaN growth process and the influence of ammonia–liquid Ga reaction on Ga transport were investigated. Experimental results under different conditions were studied and compared to theoretical predictions to quantify the mechanism of transport in the vapor growth technique. In presence of NH₃, Ga transport far exceeded the predicted upper limit for the vapor phase transport. Visual observations confirmed that a significant amount of Ga left the source in a cluster rather than atomic form.

A novel Ga source design was employed in an effort to obtain a stable and high vapor phase transport of Ga species at moderate temperatures. In this design, pure N₂ was flowed directly above the molten Ga source. This flow prevented the direct contact and reaction between the molten Ga and NH₃ and prevented Ga spattering and GaN crust formation on the source surface. At the same time, it significantly enhanced Ga evaporation rate and enabled control of Ga transport and V/III ratio in the system.

Growth characteristics were described by a mass transport model based on process parameters and experimentally verified. The results showed that the process was mass transport limited and the maximum growth rate was controlled by transport of both Ga and reactive ammonia species to the substrate surface. A growth rate of 1.4 μm/h was

obtained at 1050°C, 800 Torr, 3 slm of ammonia flow rate, and 1250°C Ga source temperature at a 24 mm source-to-substrate distance. It was found that the process required a more effective supply of active NH₃ to the substrate in order to increase the crystal quality and growth rate.

The surface morphology of the deposited layers was examined by optical and scanning electron microscopies. XRD analysis was used to determine the crystallinity of deposited films and revealed a full-width at half-maximum (FWHM) of 0.6 deg. for the (0002) GaN peak. EDX analysis was employed for the chemical characterization of the samples and showed that the deposited material contained only Ga and N elements. Room temperature PL spectrum demonstrated the optical quality of the grown samples.

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1. INTRODUCTION

GaN and related materials have already demonstrated high-power and low-noise microwave performance. But the most exciting potential is probably in optoelectronic applications. In fact, GaN and its alloys with AlN and InN recently became the basic materials for short-wavelength optoelectronics [1]. This was mainly due to their direct energy gaps covering the whole visible spectrum and a large part of the UV range. A blue laser pioneered by Nakamura and efficient blue-, green-, amber-, and white-light-emitting diodes that emerged from initiative work of Pankove hold promise of future solid-state replacements for inefficient and unreliable incandescent bulbs dating to Thomas A. Edison time in terms of their basic design. These materials, besides being a prime object of interest of device engineers, are also an exciting subject of research for a physicist due to their outstanding position in the III-V family of semiconductors. GaN (also InN and AlN), unlike other III-V materials are hard, partially ionic semiconductor compounds of highly chemical and thermal stability.

The major drawback in the development of GaN-based technology is the lack of high quality, large-area native substrates in large quantities, and most of the epitaxial research and device development rely upon heteroepitaxial growth. Growth of GaN crystals is a challenge due to the low decomposition temperature of GaN relative to its high melting temperature, the low solubility of nitrogen in Ga and the high equilibrium vapor pressure of nitrogen on GaN at moderate temperatures. Among the widely studied techniques for GaN bulk or thick film growth, many advantages of growth via vapor phase transport have been reported including the use of the simplest source species and high yield of

synthesis. In the vapor phase transport technique, evaporated Ga metal or sublimed GaN powder is transported in NH_3 atmosphere to the reaction zone in which deposition occurs. High growth rates of up to 0.5 mm/h were obtained with this method; however, the crystals grew only for short periods of time. In the best case, needle or platelet crystals were obtained but their lateral size was far from desirable. Earlier observations in the growth of GaN from the vapor phase and the problems encountered are briefly reviewed in Section 2.3.

The major challenges in the growth of GaN from elemental Ga and NH_3 have been:

- 1) limited growth time, mainly due to the poor stability of Ga source in NH_3 atmosphere,
- 2) formation of Ga droplets resulting in high but uncontrolled GaN growth rate only for short periods of time, and
- 3) formation of GaN crust over the Ga source that reduced the outgoing Ga flux and eventually caused cessation of the growth.

The goal of this study is to understand and possibly solve the challenges associated with the growth of single-crystalline GaN from metallic Ga and to develop a cost-effective, stable and reproducible process for Ga crystal growth.

In Chapter Two, a literature review covering III-N devices, the motivation for the growth of GaN, the properties and growth of GaN and the process of growth from the vapor phase are presented. The growth system where the crystal growth experiments were conducted is described in Chapter Three.

Studies on the transport of gallium (Ga) species in a vapor growth system under different conditions are presented in Chapter Four. Comparison between the experiential data and the theoretical calculations indicated that Ga species transport involves some additional mechanisms to the vapor transport which was revealed to be caused by

ammonia hitting the molten Ga surface. The modification applied to the Ga source in an effort to form reproducible and stable growth conditions and effects of this modification on the Ga transport and the stability of the Ga source are also discussed in this chapter.

Chapter Five presents the growth of GaN on sapphire with a modified sandwich growth method which is a variation of vapor phase transport process. Optimum growth conditions and the quality of the grown crystals are discussed. Necessary process improvements are pointed out which would increase the crystal quality and the growth rate.

A one-dimensional model is developed to describe and understand the observed experimental dependencies in the growth of GaN from elemental Ga and NH_3 in the modified sandwich growth technique. Results regarding the growth conditions and the validity of the model are discussed in Chapter Six. It was observed that the maximum growth rate is controlled by the transport of both Ga species and reactive ammonia to the substrate surface. Gallium nitride films were grown on (0001) sapphire substrates at 1050°C by controlled evaporation of Ga and reaction with NH_3 . It was found that the process required an effective supply of NH_3 to the substrate in order to increase both the quality and the growth rate of the crystals.

A summary of this study and a description of possible future work are presented in Chapter Seven. The main drawback of the system was found to be the supply of sufficient amount of active nitrogen species to the substrate. Calculations of the catalytic NH_3 decomposition rate revealed that the extent of decomposition strongly dependent on the growth conditions and the equipment. It was concluded that the process needs further improvement that allows the supply of sufficiently cold, in other words, not decomposed

NH₃ to the substrate region in order to achieve the growth of high quality thick GaN layers at high growth rates.

Numerical simulations of the temperature distribution and velocity streamlines within the reaction tube based on total reactor pressure, gas flow rates and heater temperatures are presented in Appendix A. These results assisted in understanding the GaN growth characteristics in the system.

In Appendix B, energy of the overall reaction involved in the GaN growth from Ga and NH₃ is calculated and compared to that of other vapor phase epitaxy processes. Studies on the catalytic decomposition of NH₃ are presented in Appendix C. Error calculations are described in Appendix D. The hardware and software safety interlocks of the reactor, and the procedures for safe and stable operation are covered in Appendix E.

2. BACKGROUND

GaN (also AlN and InN) attracted tremendous interest for applications in blue/green diode lasers and LEDs, high temperature electronics, high-density optical data storage, and electronics for aerospace and automobiles. They cover an energy bandgap range of 0.7 to 6.2 eV which covers a spectral range from infrared to ultraviolet. There is also great scientific interest in GaN and related materials because they appear to form the first semiconductor system in which extended defects do not severely affect the optical properties of devices [3].

Although GaN has been studied far more extensively than other group-III nitrides, the level of understanding necessary for technological improvements has not been achieved. Synthesis of p-type GaN has led to the demonstration of LEDs in the violet, blue, green and yellow bands of the visible spectrum with brightness suitable for outdoor displays, CW lasers, power Modulation Doped Field Effect Transistor (MODFETs), and UV detectors [4]; however, a large amount of work remains to be done in the determination of the fundamental physical properties of GaN. Information available in the literature regarding many of the physical properties of GaN is still in the process of evolution, and controversial. This is a result of measurements being made on samples of widely varying quality.

2.1. Properties of GaN

2.1.1. Chemical Properties of GaN

The elements B, Al, Ga, and In are forming compounds with N having the composition MN, where M represents a III-metal. The chemical bond of these compounds is predominantly covalent. Because of the large differences in electronegativity of the two constituents, there is a significant ionic contribution to the bond which determines the stability of these compounds. Hence, GaN is a highly stable compound even at elevated temperatures. It has a bond energy of 8.92 eV/atom. In fact, this chemical stability combined with its hardness makes GaN an attractive material for protective coating. Currently, the majority of GaN researchers are interested in semiconductor device applications. While the thermal stability of GaN allows freedom of high temperature processing, the chemical stability of GaN results in a technological challenge. Conventional wet-etching techniques used in semiconductor processing have not been very successful for GaN device fabrication. It has been reported that GaN is insoluble in H₂O, acids, or bases at room temperature, but it dissolves in hot alkali solutions at a very slow rate [5]. Low-quality GaN has been etched at reasonably high rates in NaOH, H₂SO₄, and H₃PO₄ by several researchers [7]. Although, these etches are useful for identifying defects and estimating their densities in GaN films, they are not very successful for fabrication of devices. Well established chemical etching processes are required for the device-technology development. Status of wet etching of group III nitrides was summarized by Pearton and Shul [8]. More promising possibilities are the

various dry-etching processes under development and reviewed by Mohammad et al. [7]. In addition, photo-enhanced current (PEC) and reactive ion etching (RIE) have been investigated as potential tools for patterning of III-N heterostructures [9].

2.1.2. Crystal Structure

The group III nitrides GaN, AlN and InN can crystallize in the following three crystal structures: (1) wurtzite, (2) zinc-blende (3) rock-salt. At ambient conditions, the thermodynamically stable phase is the wurtzite structure. A phase transition to the rock-salt structure takes place at high pressure. In contrast, the zinc-blende is metastable and may be stabilized by heteroepitaxial growth on substrate reflecting structural compatibility.

The wurtzite structure has a hexagonal unit cell with $c/a = \sqrt{8/3} = 1.633$. As shown in Figure 1, this structure is composed of two hexagonal close packed (hcp) sublattices shifted with respect to each other along three-fold c axis by the amount of $u = 3/8$ in fractional coordinates. Both sublattices are occupied by one atomic species only, resulting in four atoms/unit cell. Locally, every atom of one kind is surrounded by four atoms of the other kind which are arranged at the edges of tetrahedron. The symmetry of the wurtzite structure is given by space group $P6_3mc$, and the two inequivalent atom positions are $(\frac{1}{3}, \frac{2}{3}, 0)$ and $(\frac{1}{3}, \frac{2}{3}, u)$. For the III-nitrides, the axial ratios deviate considerably from the ideal value for hexagonal close packing of spheres, and hence, each lattice is distorted from the ideal geometry to some extent.

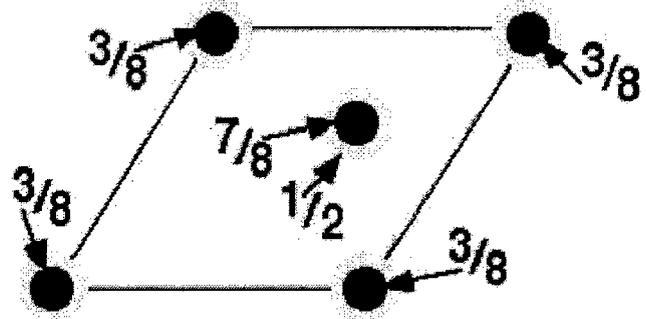
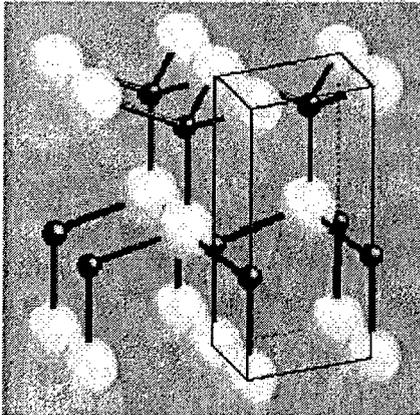


Figure 1. The wurtzite crystal structure: atoms of one specie are black, those of the other specie are white; and the unit cell is marked.

Table 1. Properties of wurtzite GaN

Property	Wurtzite GaN
Bandgap (eV)	3.4
Coefficient of Thermal Expansion ($\times 10^{-6} \text{ K}^{-1}$) [12]	\parallel -5.59 (300 - 900 K) \perp -3.17 (300 - 700 K) 7.8 (700 - 900 K)
Young's Modulus (GPa) [4]	$C_{11} = 396$ $C_{12} = 144$ $C_{13} = 100$ $C_{33} = 392$
Bulk modulus (GPa) [13]	176
Thermal conductivity ($\text{Wcm}^{-1}\text{K}^{-1}$)	1.3
Melting point (K) [14]	2791

2.1.2.1. Lattice Constant

In general, single crystals are required for accurate measurements of the lattice constant and related properties. Because of the lack of GaN single crystals, most of the data acquired to date stems from nitride layers grown heteroepitaxially. As a result, lattice constants and associated thermal and mechanical properties were found to exhibit significant scattering [10]. The most widely accepted lattice constants of wurtzite GaN are $a = 3.189 \text{ \AA}$ and $c = 5.186 \text{ \AA}$ [11]. Some of the relevant properties of wurtzite GaN are listed in Table 1.

2.1.2.2. Polarity of the Structures

None of the three structures in which GaN crystallizes possess a center of inversion. This means that, when defining an atom position on a closed packed plane with coordinates (x,y,z) , it is not invariant to the position $(-x, -y, -z)$ since inversion results in replacement of group III atoms by nitrogen atoms and vice versa. As a result of the lack of inversion symmetry, the crystal exhibits crystallographic polarity: $\{0001\}$ basal planes in the wurtzite structure differ from $\{000\bar{1}\}$ planes. A variety of properties of the material depends on its polarity, for example, growth, etching, defect generation, plasticity and piezoelectricity.

2.1.3. Electrical properties and defects in GaN

Electrical properties of unintentionally doped GaN were reviewed extensively by Strite and Morkoc [15]. They pointed out that widely varying results reflect the different crystal quality and purity of material. The room temperature electron mobility values in bulk GaN grown with HVPE to a thickness of 60 μm were reported to be 950 cm^2/Vs [16] and that reported for MOCVD grown layers was also in excess of 900 cm^2/Vs [17]. MBE grown films, however, produce much lower mobility values in the range of 100-300 cm^2/Vs [18]; this has been attributed to both high dislocation density and point defects [19]. Other important electrical properties were reviewed by Nakamura et al. [2].

Following the successful growth of GaN on sapphire substrates in the late 1960s [5], it became quickly obvious that doping and defects would play a vital role in future development of GaN. The early unintentionally doped GaN was invariably n-type, which at the same time was believed to be due to nitrogen vacancies. The high n-type background carrier concentration on the order of 10^{18} cm^{-3} proved to be difficult to reduce and the absence of a shallow acceptor dimmed the prospects of a production scale GaN based device effort [3]. The search of p-type GaN was not successful until Akasaki and Amano demonstrated this feat in 1989 [20]. This was achieved by reducing the background n-type carrier density in unintentionally doped GaN films by the use of a low temperature AlN buffer layer, and by Mg-doping followed by an ex situ low energy electron beam irradiation treatment [21, 22]. This remarkable discovery eventually led to the demonstration of a variety of bipolar devices. Even though much progress has been

made in doping of GaN, there still exist significant challenges; especially with p-type doping. The low hole mobility and low achievable free hole concentration result in large sheet resistance preventing the fabrication of reliable Ohmic contacts with low contact resistance. These material challenges have prevented the use of the AlGaIn/GaN system to its full potential in electronic applications. Furthermore, an immature p-type doping technology has led to lifetime problems and required that InGaIn laser diodes operate at a higher than expected bias voltage [3].

2.1.4. Optical Properties

Maruska and Tietjen [5] were the first to accurately measure the direct band gap energy of GaN to be 3.39 eV. Pioneering work in this area was performed by Pankove et al. who reported on the low temperature photoluminescence (PL) spectrum of wurtzite GaN. Since the primary interest in GaN is for its potential as a blue and UV light emitter, a lot of effort has been devoted to determine its optical properties. The available data were reviewed by Akasaki and Amano [24], and by Strite and Morkoc [15].

2.1.5. Summary of Properties

Table 2 summarizes the most important properties of wurtzite GaN.

Table 2. Wurtzite GaN parameters

Property [2]	Wurtzite GaN
Density (g/cm ³)	6.15
Melting Temperature (°C)	>1700
Lattice Constant, a (Å)	3.189
Lattice Constant, c (Å)	5.185
Electron mobility (cm ² /Vs)	1000
Hole mobility (cm ² /Vs)	30
Breakdown field (V/cm)	>5 x 10 ⁶
Thermal Conductivity (W/cmK)	1.5
Energy Gap (Γ Valley) (eV)	3.39
Effective Mass (Γ Valley) (m _e)	0.20
Deformation Potential (eV)	8.3
Polar Optical Phonon Energy (meV)	91.2
Piezoelectric Constant e ₁₄ (C/m ²)	0.375
Static Dielectric Constant	8.9
High Frequency Dielectric Constant	5.35
Transverse Constant (C _t) (dyn/cm ²)	4.42 x 10 ¹¹
Longitudinal Constant (C _l) (dyn/cm ²)	2.65 x 10 ¹²
Transverse Sound Velocity (cm/s)	2.68 x 10 ⁵
Longitudinal Sound Velocity (cm/s)	6.56 x 10 ⁵

2.2. Growth of GaN

The group of III-nitrides do not exist in nature and thus must be synthesized. Crystal growth of this group using standard methods (Czochralski, Bridgman) is extremely difficult because of the following factors:

1. the high melting temperature (~ 2800 K for GaN);
2. the relatively low sublimation – decomposition temperature compared to the melting temperature;
3. the very high equilibrium nitrogen vapor pressure (~ 40 kbar at melting temperature) and
4. low solubility in acids, bases, and most other inorganic elements and compounds.

Thus, novel growth techniques must be employed for III-nitrides. GaN was first synthesized by Johnson et al. [25] in the early 1930s by passing NH_3 gas over hot metallic Ga. Two decades later, Grimmeiss et al. [26] employed the same technique to produce small crystals of GaN for the purpose of measuring their photoluminescence spectra. Another decade later, Maruska and Tietjen used a chloride transport vapor technique to form an epitaxial layer of GaN on a sapphire substrate [5]. All these and, the employment of nucleation buffer layers by Amano et al. [21], Yoshida [27] and Moustakas et al. [23] are considered to be pivotal improvements in the growth of nitrides.

Nearly every crystal growth technique, substrate type and orientation has been tried in an effort to grow high quality group III-nitride thin films [28].

Due to the lack of native GaN substrates, GaN based nitride semiconductors have been grown hetero-epitaxially. Efforts to produce GaN [29, 30] and AlN [31, 32] bulk materials for investigating their properties as well as potential use as substrates for epitaxy are at their early beginnings. Despite this limitation, remarkable progress in the growth of high quality epitaxial III-nitride films by a variety of methods such as Hydride Vapor Phase Epitaxy (HVPE) [5, 33], Metal Organic Vapor Phase Epitaxy (MOVPE) [21, 34], Reactive Molecular Beam Epitaxy (MBE) [35], and bulk or thick crystal growth by vapor phase transport [36, 37] and solution techniques at high pressure and high temperature have been achieved [38].

The most desirable methods of growing high purity single crystal Group III nitrides would be those in which these nitrides themselves are used as starting material. This has been a goal for many researchers. Slack and McNelly [39] suggested a technique in which high quality AlN powder from the Al metal using AlF_3 as an intermediate product can be produced. The AlN powder can be converted to a single crystal by sublimation in a closed tungsten crucible or in an open tube with a gas flow. Over the years, physical vapor transport (PVT), which is the most popular technique for growth of bulk AlN, has been intensively explored. So far, the best results for the growth of bulk AlN were achieved using the PVT method [32, 40-42]. GaN is harder to grow due to a much higher nitrogen overpressure required to stabilize GaN at high growth temperatures. High quality, transparent and colorless platelets of GaN were grown from the in liquid gallium

at high temperatures (1400-1600°C) and high N₂ pressures of up to 20 kbar; however, this method is not scaleable [38].

Since Johnson et al. [25] synthesized polycrystalline GaN deposits by the reaction of NH₃ with molten Ga, several researchers have employed similar methods to grow GaN single crystals. Ejder [36] studied the growth of GaN from the vapor phase by the reaction of Ga and NH₃ to investigate the effect of the Ga partial pressure on crystal morphology. Callahan et al. [44] observed the formation of high aspect ratio crystals on the upper surface of the GaN crust, where the Ga vapor concentration presumably is low. Beneath the crust, growth was preferentially along the basal plane of GaN. The decomposition of a solid gallium tribromide-ammonia complex (GaBr₃:4NH₃) in an ammonia, argon and hydrogen atmosphere was used for the deposition of gallium nitride on SiC substrates [45]. The deposition rate of gallium nitride in argon was considerably less than in ammonia. Pichugin and Yaskov [46] obtained needles of 5 x 0.3 mm by flowing ammonia over the Ga in a quartz boat for several hours at 1 bar pressure and 1150°C. A more detailed study aimed at determining the optimum conditions for the growth of GaN crystals was conducted by Elwell et al. [37, 47]. The largest crystal, 2.5 mm long, was nucleated on a silicon carbide rather than on a GaN seed. Manchon et al. [48] grew hexagonal needles up to 2 mm in length by heating GaN powder at 1150-1200°C in pure ammonia for 16 hrs. Zetterstrom [49] grew rather large crystals of up to 5 mm x 1 mm by heating pre-synthesized GaN above 1100°C in an ammonia flow for about 72 hours. The needles were colorless as long as the temperature did not exceed 1200°C. A more detailed series of studies of the morphology of GaN crystals formed primarily by sublimation was undertaken by Aoki and Ogino [50]. They compared the

reaction between the ammonia and gallium in an open quartz boat with that in a partly closed quartz ampoule. Sakai et al. [51] have obtained GaN needles by sublimation of GaN powder in a 50 sccm flow of NH_3 via deposition onto sapphire substrates at a substrate temperature of 1050°C . However, the crystal size was limited to small needles.

Balkas et al. [52] were successful in obtaining high quality GaN crystals up to 1 mm in length via sublimation with a growth rate of 0.5 mm/hr in the direction perpendicular to the basal plane. They also found that the direction of the fastest growth and thus the crystal shape were controlled by changing the Ga/ NH_3 flux ratio and the growth temperature. Kamler et al. [53] investigated GaN sublimation in a horizontal quartz tube reactor. They reported that the sublimation process was slower but more stable in terms of reactant transport as compared to Ga as the source material.

Growth by a modified vapor phase transport process called the sublimation sandwich method (SSM) was reported by Wetzell and co-workers [54, 55]. The GaN films were grown from the metallic Ga and activated ammonia on (0001) 6H-SiC, using a modified sandwich method originally described in [56]. A quartz reactor contained several deposition cells in a row, which shared the same ammonia gas stream. Each cell consisted of a Ga source and substrate, separated by a gap of only 5 mm. The ammonia flow rate through the gap was 25 to 50 slm at atmospheric pressure. Under these conditions, there was an effective mass transport of Ga vapor and activated nitrogen towards the substrates. At growth temperatures between 1170 and 1270°C , single crystal layers of GaN were obtained at growth rates up to $300 \mu\text{m/h}$. Vodakov et al [57] succeeded in growing GaN on 6H-SiC substrates by using the SSM technique with deposition rates up to $200\text{--}1100 \mu\text{m/h}$. GaN films with thicknesses between 70 and $100 \mu\text{m}$ showed near-

bandgap photoluminescence at 3.48 eV with a FWHM of 8 meV. The FWHM of the (0002) ω -scan was measured to be in the range of 270–300 arcsec, indicating good structural quality for films grown at very high rates.

2.3. GaN growth from the Vapor Phase

Growth of crystalline films from the vapor phase onto solid substrates, involves the physical transport of gas phase precursors to a solid surface where film constituents are incorporated and gaseous byproducts are released. The fundamental steps that occur in the process involve the transport of reactants to the reaction zone, which decompose or react with one another depositing a film on the substrate and transport of the reaction by-products away from the reaction zone.

2.3.1. Sublimed GaN vs. Evaporated Ga as Ga Source

It is generally accepted that the process is called “sublimation” when the source material is pre-synthesized GaN and “evaporation-condensation” when pure Ga is used [51]. The sublimation process was reported to be slower but more stable in terms of transport of reactants [53]. However, the reasoning behind this is still not clear. In spite of its high growth rate and simplicity in apparatus, the sublimation technique has an inherent difficulty in that pre-synthesized GaN powder or crystals must be continuously supplied. On the other hand, the use of pure Ga as the source material provided the maximum growth rate, but resulted in poor stability [58, 59]. The high GaN growth rate is thought

to originate from Ga droplets generation due to instability of the Ga source [59]. Elwell et al. [37] reported the formation of a GaN crust on the molten Ga during the GaN crystal growth by the reaction between Ga and NH_3 . Subsequent studies revealed the cessation of growth is due to the crust formation [44]. Shin et al. [58] observed that the violent reaction at the molten Ga surface caused spattering of Ga and the contact of Ga droplets with already formed GaN crystals changed the growth kinetics. They also reported that the increase in the flow rate to 325 sccm and above resulted in the formation of a thin crust of GaN on the source surface and the cessation of the violent reaction. With a controlled and stable Ga source continuous growth could be achieved within an appropriate processing window.

2.3.2. Epitaxial Growth

Attempts to grow free-standing GaN crystals in the Ga – NH_3 system without seed or substrate were not successful [60]. In the best case, only mm size needle or platelet crystals were obtained.

With the discovery of X-ray diffraction in 1912, the knowledge of internal crystal structure developed very rapidly and by the time Royer (1928) carried out his extensive and systematic studies of a wide variety of overgrowths, it was possible to examine the effect of the geometry of crystal structure on orientation. Royer introduced the term *epitaxy* (“arrangement on”) [61]. Based on his results, he established the requirements for epitaxy. The most important is that oriented growth occurs only when it involves the parallelism of two lattice planes that have identical or quasi-identical structure and

closely similar atomic spacing. The experiments indicated that epitaxy occurred only if the misfit was no more than 15%. The misfit was defined as $100(b-a)/a$, where a and b are the corresponding spacings in the substrate and overgrowth, respectively.

Over the past decades, epitaxy has left the laboratory and assumed crucial importance in the processing of electronic and optoelectronic devices. Two epitaxial processes are commonly applied: 1) homo-epitaxy and 2) hetero-epitaxy. In the homo-epitaxy the growth takes place on the substrate of the same material (native substrate). In the optimal conditions, the growth is two-dimensional and replicates the surface layer-by-layer with the same atomic arrangement as the substrate crystal. During the hetero-epitaxy however, single crystalline materials are grown on non-native substrates. In this case, the growth of the first layer takes place on a different crystal from the growing layer. The lattice mismatch and induced stress on the substrate surface cause the adatoms distribution to vary from the crystal structure, which is being grown. The first layer of adsorbed atoms on the surface tends to occupy those spots that give minimal surface stress [62]. During the growth of the first few layers, misfit dislocations are formed [63]. Buffer layers, grown at lower temperatures, have been routinely used in order to relax the stresses during the nucleation stage [64, 65].

2.3.3. Substrates for Epitaxy

The unavailability of native single crystalline substrates for homo-epitaxy has been the biggest problem in the growth of GaN and other III-N thin films. Therefore, the hetero-epitaxy was a practical necessity. The use of ammonia and operating temperatures

above 1000°C additionally limited the choice of non-native substrates for growth of GaN in the Ga – NH₃ system. Sapphire has been by far the most popular choice owing to its low price, availability of large-area, high-quality crystals, its transparent nature and high temperature stability. Sapphire is not the best choice in terms of lattice mismatch (17% mismatch with GaN). However, it has been observed that the dislocations formed during GaN growth due to lattice mismatch with the substrate are not detrimental to the electrical and optical properties although the dislocation density is in the range of 10⁹-10¹⁰/cm². Both c-plane (0001) and r-plane (1 $\bar{1}$ 02) sapphire have been studied extensively [66]. GaN films grown on c-plane (0001) substrates grow (0001)-oriented. The low surface energy of this plane results in poor hetero-nucleation, resulting in rough film morphology, usually dominated by the sapphire's surface features or contaminants. In addition, due to a large lattice mismatch between the sapphire and GaN, the stress relaxation occurs in a region close to the substrate-crystal interface and misfit dislocations form at a very high density [64, 67, 68]. The thermal expansion mismatch also introduces stresses during the cooling process, which may result in delamination and cracks of the layers grown on sapphire.

2.3.4. Species Transport

Fluid mechanics, heat transfer, and species transport are critical in determining both the transport of precursors to the growth interface and the degree of gas phase reactions occurring prior to the species participation in the surface growth reactions. The movement of precursors to the growth surface is determined by the geometry of the

system involved as well as the pressure, temperature, and type of gas under consideration. It is therefore crucial to understand and control factors governing transport processes underlying crystal growth from the vapor phase.

At low gas densities, free molecular flow occurs. The chambers of high vacuum evaporators operate within the molecular flow regime. In this case the mean distance between molecular collisions is large compared to dimensions of the system. The kinetic theory of gases provides an accurate picture of molecular motion under such conditions. At higher pressures, however, the mean distance between successive molecular collisions is reduced, and they dominate relative to molecule – chamber wall collisions. The gas is said to be in a viscous flow regime. Compared to the molecular flow, viscous flow is quite complex. Criteria for distinguishing between the flow regimes are based on the magnitude of the Knudsen number, Kn , which is defined by the ratio of the gas mean free path (λ_{mfp}) to a characteristic dimension D_p of the system, i.e., $Kn = \lambda_{mfp} / D_p$. At the atmospheric pressure, λ_{mfp} of the gas is as small as several microns that the flow of the gas is limited by its viscosity. At very low pressures, where λ_{mfp} is much larger than the dimensions of the vacuum enclosure, the flow is molecular and $Kn > 1$. In between, at intermediate, where λ_{mfp} of the molecules is similar to the dimensions of the vacuum enclosure, the gas flow is governed by the viscosity as well as molecular phenomena. In the state where the gas is viscous, the flow can be laminar or turbulent. At low gas velocities, the flow is laminar and layered, parallel gas flow lines may be imagined. Under these conditions, the laminar flow velocity is zero at the walls of a tube, but it increases to a maximum at the tube axis. For higher flow velocities, the gas layers are no longer parallel but swirl. In this turbulent flow range, cavities of lower pressure develop

between layers. The limit between the turbulent and laminar flow is defined by the value of Reynold's number. The Reynold's number is a dimensionless quantity expressed by $R_e = \rho v D / \eta$ where ρ is the gas density, v the velocity, η the viscosity, and D the diameter of the tube.

Models of crystal growth from the vapor consist of nonlinear, coupled partial differential equations that represent the conservation of momentum, energy, total mass, and individual species [70]. The general derivation and form of these equations is given in standard references in transport phenomena textbooks [71 - 73] and also explained in more detail in Appendix A. The general equations in the continuum regime apply for typical reactors operating at sufficiently high pressures and with large characteristic dimensions such as reactor height and substrate spacing, so that the mean free path is small relative to these characteristic dimensions.

2.3.5. Thermodynamics and Kinetics of GaN Growth

Thermodynamics addresses a number of important issues with respect to growth from the vapor. These calculations can provide information on the partial pressures of the involved gaseous species and the direction of transport in case of reversible reactions. Importantly, it provides an upper limit of what to expect and an estimate of the extent of departure from equilibrium under specific conditions. Epitaxial growth must take place under relatively low gas-phase supersaturations so that the reactant partial pressures do not deviate excessively from the equilibrium values. If the arrival rate of the reactant

species greatly exceeds the nucleation and lattice incorporation rate at the crystal surface, poor crystal perfection will result [74].

Thermodynamics, on the other hand, does not address questions related to the speed of the reaction and the resulting growth rates. Indeed, processes that are thermodynamically possible frequently proceed at low rates because of both transport and reaction kinetics [75]. Generally, it is impractical to operate a vapor phase epitaxy process with so low flow rates that permit complete equilibration of the incoming reactants with the growing layer. Thus, a portion of the reactants must pass unreacted, and the overall epitaxial deposition rate may be limited by the rate of mass transfer of reactants from the main gas stream to the substrate surface where they undergo reaction. This is the classic diffusion or mass transport controlled process. A convenient working model assumes a hypothetical boundary layer or stagnant gas region near the substrate surface where the gas stream velocity approaches zero. Outside this boundary layer the reactant concentrations are at their bulk or mainstream values. This subject is discussed in more detail in Chapter 6.

Thurmond and Logan [76] calculated the equilibrium pressure of N_2 above GaN. More recently, Karpinski et al. [77] used the partial decomposition of GaN in a high pressure apparatus to measure the equilibrium pressure of N_2 over GaN, over a wide range of temperatures. The equilibrium N_2 pressure values were reported to be around 100 atm at 1100°C. The nitrogen molecule with its highly stable N-N triple bond (9.8eV) is extremely difficult to break into its atomic constituents that are necessary for GaN formation. While this is not a barrier to growth of GaN crystals in high-temperature, high-pressure solution growth, it makes the growth of GaN impossible at pressures below

atmospheric pressure. This high stability necessitates dissociation of N_2 using various processes such as those utilized in plasma assisted molecular beam epitaxy (MBE), electron cyclotron resonance plasma assisted MBE, and radio frequency plasma MBE. All these approaches require modifications to the growth reactor and necessitate a significant investment. Even with the use of highly energetic nitrogen species, the thermal energy on the growing surface is important for the uptake of both N and Ga [78]. While all of the above techniques have demonstrated GaN growth, they are inadequate for bulk growth of GaN because of the low growth rates.

Because of these very high equilibrium nitrogen pressures, most crystal growth experiments have used ammonia as the source of nitrogen. Thermodynamically, almost all the NH_3 is decomposed into N_2 and H_2 at temperatures higher than $300^\circ C$. Calculation of the free energy of the reaction



showed that the formation of GaN is favored over the temperature region of interest. However, the growth rate of GaN is negligible at $900^\circ C$, showing that the chemical reaction is kinetically controlled. Calculations of free energy in several GaN growth processes and their comparison with respect to GaN growth from Ga and NH_3 can be found in Appendix B. It is also well known that NH_3 decomposition is a catalytic process [79] and the extent of the decomposition strongly depends on the growth conditions and the equipment. A detailed study on NH_3 decomposition is presented in Appendix C.

The growth kinetics in growth from vapor phase is dependent on a number of factors associated with transport of reactants to the substrate surface and subsequent heterogeneous surface reactions that lead to crystal formation. Intimate microscopic

details of these reactions are often unknown, thus growth kinetics is frequently modeled in macroscopic terms.

A growth model has been proposed for the formation of GaN from Ga and NH_3 and is illustrated in Figure 2 [80]. In a first step, and in order to form the GaN molecule, the reaction barrier must be reduced either by an increase in the energy of the reactant molecules or through a feasible chemical reaction path. The first step of the proposed reaction sequence includes the adsorption of NH_3 on a Ga site, either within the GaN surface or an adsorbed Ga atom on the substrate surface. The bonding between Ga and N is through the lone pair electrons of NH_3 . The second and third steps involve the formation of Ga-ammonia intermediate molecules coupled with the migration of H from the adsorbed ammonia to the surface and further reaction of Ga with the ammonia to replace H with Ga. The final step involves the formation of GaN and liberation of H_2 from the surface. The concept of on surface cracking (OSC) of ammonia has been introduced by Kamp and co-workers [81, 82].

Several other studies reveal that, following physisorption to the surface, NH_3 instantaneously chemisorbs and releases a hydrogen atom at temperatures above 200K. At moderate temperatures, NH_2 and NH are formed, and H_2 is released from the surface. With increasing temperatures, the remaining H is stripped off and nitrogen forms a nitride with the surface [83-84].

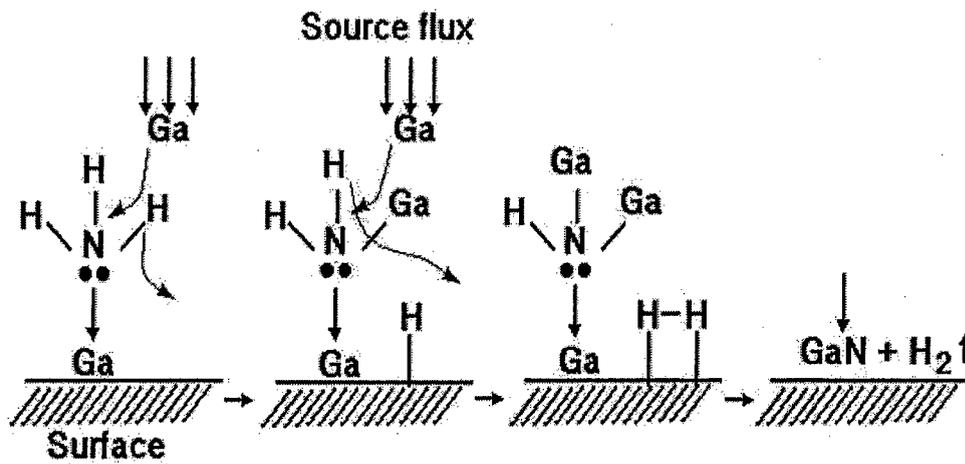


Figure 2. Schematic illustration of GaN formation on a solid surface starting with Ga and NH₃ precursors.

3. EXPERIMENTAL SETUP

Crystal growth experiments were conducted in a computer controlled growth system, custom designed for GaN growth. The growth system is shown schematically and pictorially in Figs. 3 and 4, respectively. The complete system consists of a stainless steel growth chamber, a chemical pump to maintain the growth pressure in the reactor tube, a mechanical pump to evacuate the outer chamber, a control tower including an in-situ monitoring facility, a gas delivery and control system, and a chiller for the outer chamber.

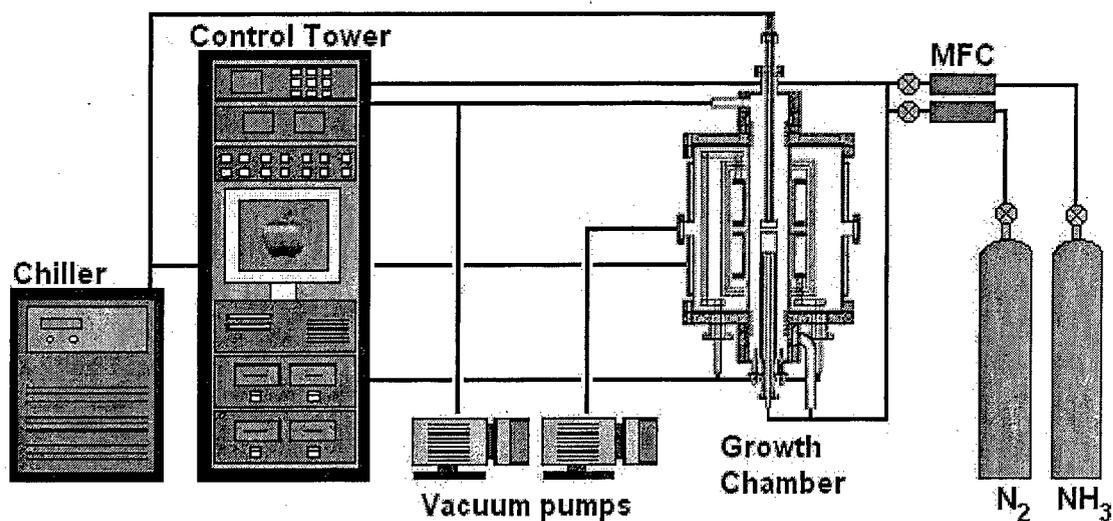


Figure 3. Schematic diagram of the GaN growth system

The growth furnace was operated using a PID controller programmed with LabView. A variety of safety features were employed. An interlock system was programmed to shut down the whole system and the flow of all gases in case of overheating, overpressure, or loss of cooling water.

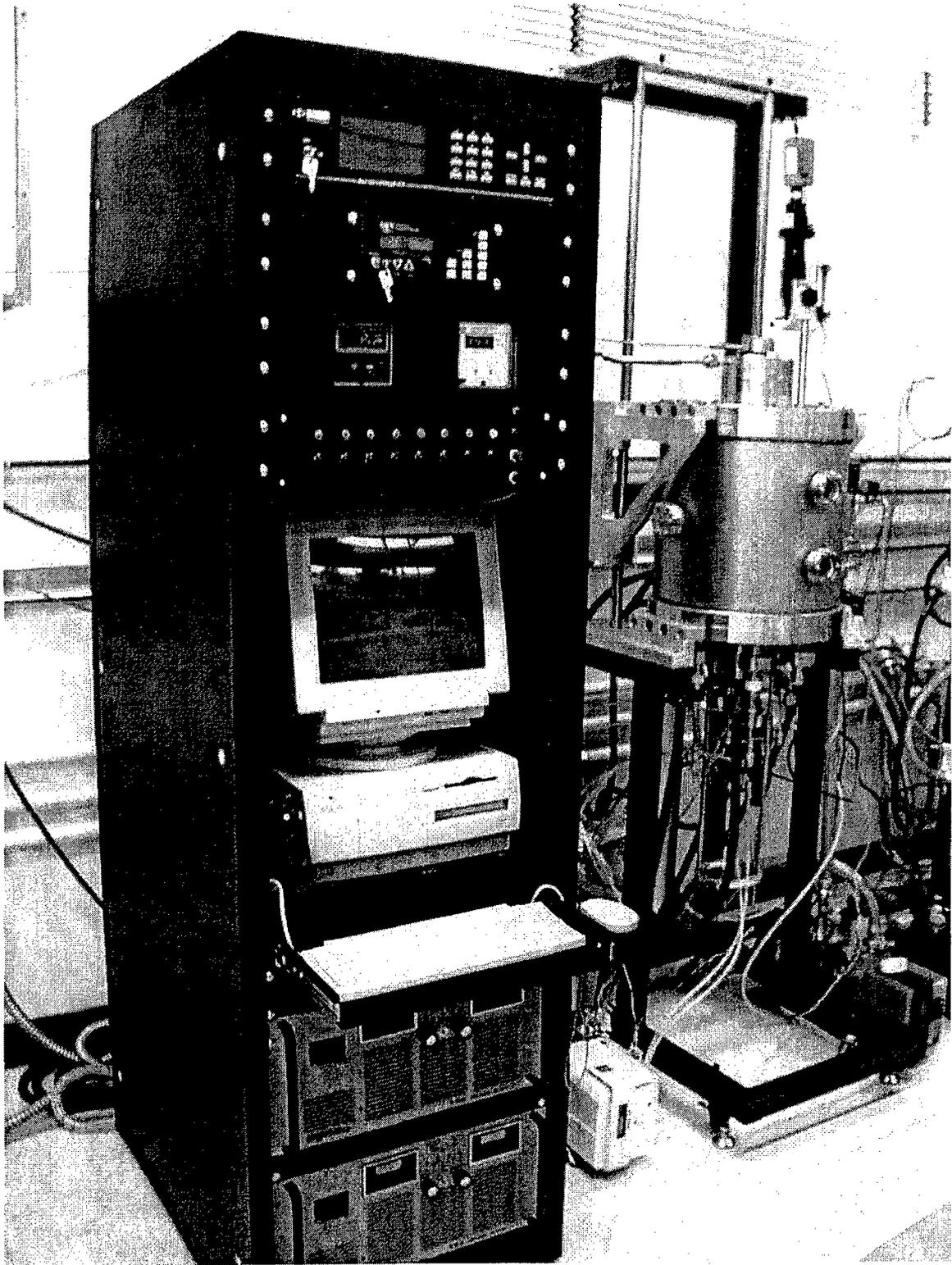


Figure 4. Picture of the GaN growth system

Fig. 5 shows the system cross-section. The 16.5"-diameter, double-walled, water-cooled outer stainless steel growth chamber contained a 4"-diameter alumina reaction tube surrounded by two graphite heaters. The heaters were designed to have specific resistances at high temperatures to comply with the outputs of power supplies and were fabricated from a high purity grade of graphite (>99%, 2020 grade graphite from Carbone of America). Both graphite heaters are supported by and connected to water-cooled electrical feedthroughs. Radiation shields made of molybdenum sheets were placed around the two heaters to reflect the heat effectively toward the center of the alumina reactor tube. To monitor the temperature, two type-C thermocouples were inserted through small holes in the molybdenum radiation shields and placed 2 mm away from the center of each heater. The thermocouples were not in a direct contact with the heaters to avoid an abrupt change in reading in case of detachment of the thermocouple from the heater.

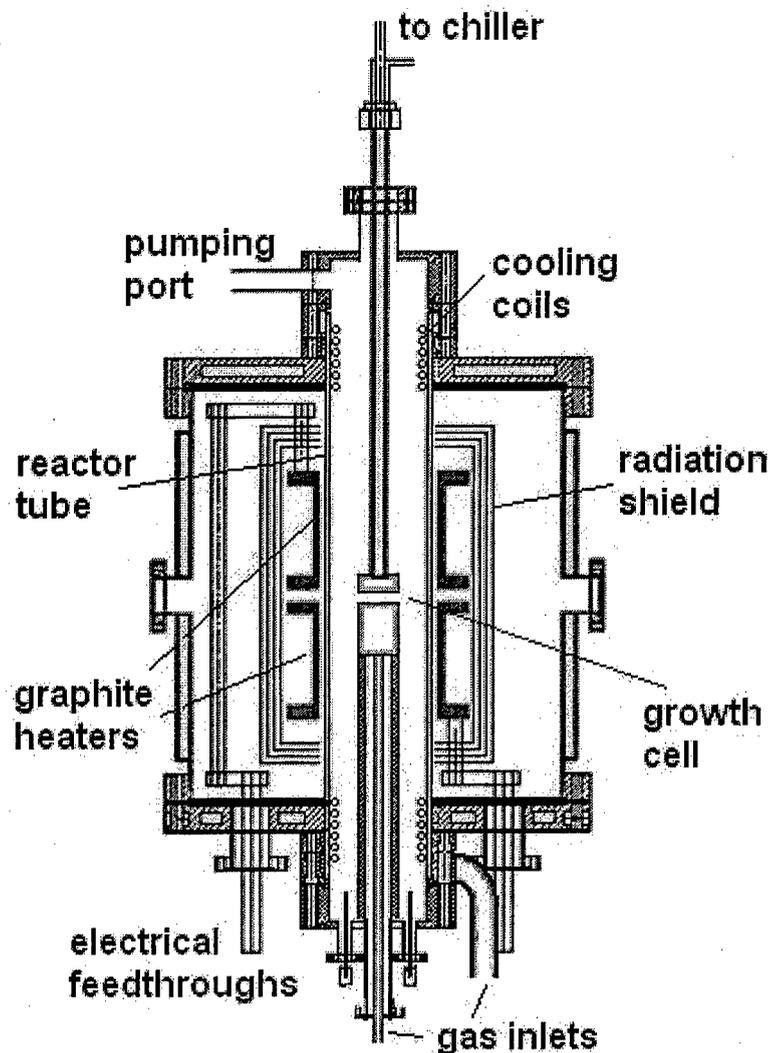


Figure 5. Schematic diagram of the GaN growth chamber.

The outer chamber was continuously evacuated by a mechanical pump to minimize the oxidation of graphite heaters and to prolong their lifetime. The pressure in the outer chamber was maintained below 3×10^{-3} Torr even at a heater temperature of 1750°C . The outer chamber was equipped with a water jacket around the outside as well as top and bottom flanges. Two viton O-rings were used to seal the gap between the alumina

reaction tube and the outer chamber. To avoid O-ring deterioration at high temperatures, a water-cooled copper coil separately cooled the O-rings. A pneumatic cylinder was used to lift either the top flange or the chamber for maintenance.

All the fixtures inside the alumina reaction tube were manufactured from molybdenum (Mo) and hot pressed, binderless boron nitride (BN) due to high temperature stability and inertness of these materials to NH_3 . A chemical pump was used in tandem with a butterfly valve to maintain the desired pressure in the alumina reaction tube.

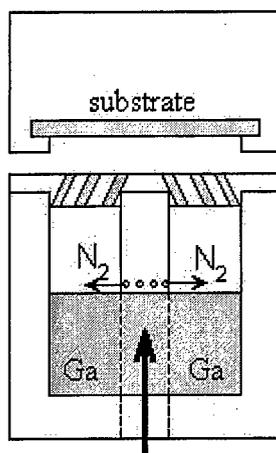


Figure 6. Schematic diagram of the growth cell with N_2 flow over the molten Ga source.

The growth cell consisted of a Ga source and a substrate in a sandwich configuration as shown in Fig. 6. UHP-grade nitrogen (N_2) was flowed directly above the molten Ga source with a controlled flow rate to prevent contact between the molten Ga source and NH_3 and to gain enhanced and controlled Ga evaporation and transport to the substrate. The Ga source holder was closed using a cap with holes to balance the gas flow velocities

of N_2 and NH_3 . The temperatures of each part of the growth cell were measured with type-C thermocouples. The temperature distribution can be seen in Appendix A.

4. GROWTH OF GAN CRYSTALS VIA PHYSICAL EVAPORATION: ON THE TRANSPORT OF GALLIUM

4.1. Outline

The transport of gallium (Ga) species in a vapor growth system for gallium nitride (GaN) growth was investigated under different conditions. A simple mass-transport model based on total reactor pressure, gas flow rates and source temperature was developed and verified. Comparison between the experiential data and theoretical calculations indicated that Ga species transport in presence of ammonia (NH_3) is not solely by evaporation and diffusion. The reaction between the molten Ga and ammonia caused spattering of the Ga source and transport of Ga clusters. Subsequently, pure nitrogen (N_2) was flowed directly above the molten Ga source to prevent direct reaction between the molten Ga and NH_3 , to enhance Ga evaporation rate and, at the same time, to enable control of Ga transport to the substrate.

4.2. Introduction

Widely studied techniques in GaN bulk or thick film growth are high-pressure, high-temperature crystal growth [29, 89], hydride vapor phase epitaxy [90, 91] and growth via vapor phase transport [53, 92, 58]. Among these techniques, many advantages of growth via the vapor phase transport have been reported including high growth rate, high crystal quality, use of the simplest source species, and high yield of the growth components [92].

In vapor phase transport technique, evaporated Ga metal or sublimed GaN powder is transported in NH_3 atmosphere to the reaction zone in which deposition occurs.

Advantages and disadvantages of the use of both pure GaN powder and Ga metal are discussed in Chapter two, section 2.3.1. In previous studies, high GaN growth rates for limited times accompanied with crust formation and instability of the Ga source were observed [59]. In fact, Shin et al. [58] observed that a violent reaction at the molten Ga surface caused spattering of Ga and the contact of Ga droplets with already formed GaN crystals changed the growth kinetics. They also reported on the formation of a GaN crust on the Ga source surface and subsequent cessation of the growth reaction. Hence, understanding the transport mechanism of Ga in the ammonia-liquid Ga system is paramount for the development of a stable and reproducible growth process.

Therefore, in this chapter, transport of Ga species during the GaN growth process and the role of ammonia-liquid Ga reactions on transport were investigated.

Experimental results were compared to theoretical predictions to quantify the mechanism of transport in vapor growth technique. The modification applied to the source holder in an effort to form reproducible and stable growth conditions was discussed.

4.3. Experimental Procedure

Experiments were conducted in the system described in Chapter 3. Two different sets of experiments were performed. The first set was employed to measure the amount of Ga transport under different conditions using the growth cell shown in Fig. 7 (a and b) where Ga source holder is a BN open cylindrical cup with 1" inner diameter. In some of the

experiments, Ga source surface was heated with a hot tungsten filament with the purpose of cracking NH_3 before reaching the source as shown in Fig. 7 (b) where the filament temperature was 1350°C . It is well known from the early surface science studies that tungsten is an effective catalyst for ammonia decomposition [83]. All the experiments were conducted at a constant source temperature of 1250°C . Four different experimental conditions were: 1) N_2 atmosphere, 2) NH_3 atmosphere, 3) vacuum (5mTorr), and 4) in NH_3 with hot filament above the source. The reactor pressure and flow rates of gases were 750 Torr and 700 sccm, respectively in all experiments except the ones under vacuum. The source surface area was 5 cm^2 , and the source-to-substrate distance was constant at 2.5 cm.

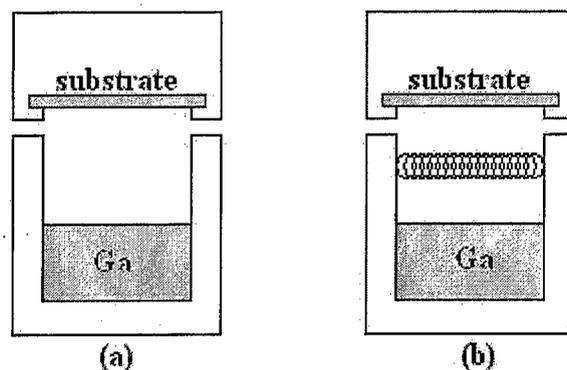


Figure 7. Schematic diagram of the growth cell. (a) BN cylindrical source holder. (b) Growth cell with the hot tungsten filament.

In the second experimental design, pure nitrogen (N_2) was flowed directly above the molten Ga source with a controlled flow rate to prevent the contact between the molten Ga and ammonia and to gain enhanced and controlled Ga evaporation rate and transport to the substrate. The source holder was closed using a cap with holes to balance

the gas flow velocities of N_2 and NH_3 and to avoid any entry of NH_3 gas inside of the source holder. A schematic figure of this growth cell can be seen in Fig. 6.

4.4. Theoretical calculations

The general phenomenon of the vapor phase transport is complicated. The evaluation of the upper and lower limits of the vapor phase transport enabled the analysis of the transport mechanism in the GaN growth system. Steady state diffusion of Ga through a stationary gas (N_2 or NH_3) forms the lower limit of mass transport in the system while free evaporation into vacuum forms the upper limit.

A one-dimensional diffusion model was used to calculate the lower limit of Ga transport. At the source surface, the liquid Ga is in equilibrium with its vapor, and the Ga partial pressure P_{Ga} is the equilibrium vapor pressure at a given temperature. This pressure decreases toward the substrate surface until it reaches P_{Ga}^* . This provides the concentration gradient which is the driving force for diffusion of Ga. As shown in Figure 8, diffusion occurs in the x direction due to the pressure decrease $P_{Ga} - P_{Ga}^*$ in the x direction over a distance L. The gradients are assumed to be linear.

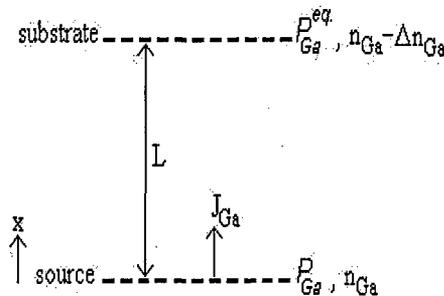


Figure 8. Illustration of mass transport by diffusion.

The diffusion flux of Ga illustrated in Figure 8 is represented by [93]

$$J_{Ga} = -D_{Ga-i} \frac{\Delta P_{Ga}}{\Delta x} \quad (1)$$

where $\Delta P_{Ga} = P_{Ga} - P_{Ga}^*$ and D_{Ga-i} is the diffusivity of Ga in N_2 or NH_3 . At the source surface, $x = 0$, the pressure of Ga is simply equal to the vapor pressure of Ga at a given temperature and at the substrate level, $x = L$, assuming the complete depletion, the pressure of Ga is equal to zero.. The result from the kinetic theory, when corrected for the Maxwellian distribution of velocities and for velocity persistence, is [94]

$$D_{Ga-i} = \frac{8 \left(\frac{2kT}{\pi} \right)^{1/2} \left(\frac{1}{m_{Ga}} + \frac{1}{m_i} \right)^{1/2}}{3\pi (n_{Ga} + n_i) (d_{Ga} + d_i)^2} \quad (2)$$

If the concentrations are replaced by P/kT and a nitrogen or ammonia rich gas phase is assumed, diffusivity is approximately proportional to $T^{1.5}$ and P^{-1} . It is usual to represent the diffusivity in gases by

$$D_{Ga-i} = D_{0,Ga-i} \left(\frac{P_0}{P} \right) \left(\frac{T}{T_0} \right)^n \quad (3)$$

where n is an empirical exponent in the range of 1.66 – 2.0 [93, 95, 96]. $D_{0,Ga-i}$ is the diffusivity of Ga in N_2 or NH_3 at standard temperature $T_0 = 273$ K and pressure $P_0 = 1$ atm. Typical values for D_0 span the range of 0.1-10 cm^2/s [95]. Whenever measured values of diffusivities are available, they should be used. Typically, the measurement errors are less than those associated with the predictions [96]. For rough estimates,

diffusivity values can be calculated via Eq. (2) with better than a factor of two accuracy [93]. D_0 and n values were measured for the Ga-N₂ system to be 0.131 cm²/s and 1.7, respectively, using a measurement set-up shown in Figure 9. The set-up is simply a closed Ga container with a small opening of a known geometry at the top, through which diffusion takes place.

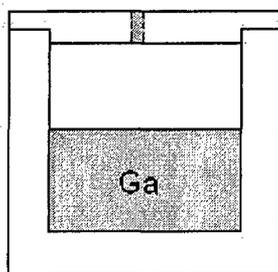


Figure 9. Schematic of diffusivity measurement set up.

In order to analyze the upper limit, Ga transport under vacuum was compared to the theoretical data obtained from Langmuir's mass evaporation rate calculation which estimates the evaporation from a free surface. Hertz stated that a liquid has a specific ability to evaporate and can not exceed a certain maximum evaporation rate at a given temperature, even if the supply of heat is unlimited [96]. Furthermore, the theoretical maximum evaporation rates are obtained only if as many evaporant molecules leave the surface as it would be required to exert the equilibrium pressure on the same surface while none of them must return. The latter condition means that a hydrostatic pressure of $P = 0$ must be maintained. Langmuir investigated the evaporation of tungsten from filaments in evacuated glass bulbs and assumed that the evaporation rate of a material at a pressure below 1 Torr is the same as if the surface were in equilibrium with its vapor and

thereby excluded the recondensation of evaporated species [97]. He derived the maximum mass evaporation rate per unit area as

$$\Gamma = \left(\frac{m}{2\pi kT} \right)^{1/2} P^* \quad (5)$$

where P_{Ga} is the equilibrium vapor pressure of the material at temperature T . Therefore, the mass evaporation rate of Ga in vacuum, where the ambient hydrostatic pressure acting upon the Ga in the condensed phase was 5×10^{-3} Torr, was calculated based on Langmuir's mass evaporation studies. Ga vapor pressure data previously measured by Honig and Kramer [98] were used in the calculations.

4.5. Results and discussion

The lower limit of the mass transport was calculated from the transport rate via pure diffusion of Ga in stagnant N_2 gas using the experimental parameters and geometry, and it was found to be 6.3×10^{-3} g/h. On the other hand, upper limit was calculated to be 1.80 ± 0.41 g/h using Langmuir's mass evaporation rate. Transport data measured in the first set of experiments, where the Ga source surface was open to contact with NH_3 as shown in Fig. 7(a), are tabulated in Table 3. The data show that Ga transport in NH_3 gas, without a filament, was higher than that in N_2 by a factor of 50. It was even higher than the transport under vacuum by a factor of 2.5. When this value is compared to the theoretical calculations, it is seen that there is a 2 orders of magnitude difference between the experimental and the calculated value. In the experiments with the hot filament above the molten Ga, as shown in Fig. 7(b), the measured Ga transport under the NH_3 flow was

approximately the same as that measured in N₂ and below the value obtained under vacuum. It was also observed that under these conditions Ga droplet formation diminishes, source stays stable and the formation of the GaN crust over the source is avoided.

Table 3. Measured Ga mass transport rate under different conditions

Experiment ^x	Ga mass transport rate (g/h)
in N ₂ atmosphere	7.2x10 ⁻²
in NH ₃ atmosphere	3.6
in NH ₃ atmosphere and filament assisted	7.6x10 ⁻²
under vacuum	1.5

^xsource surface area= 5 cm², T=1250°C, flow rate=700 sccm

The difference between the measured values and theoretical calculations for the Ga transport in N₂ gas can be due to a number of factors such as deviations from the assumed conditions in the theoretical modeling and the effect of flow rate of 700 sccm on the diffusion length, L. During the measurements, it was necessary to flow ammonia to be able to maintain a constant ammonia concentration, and in the measurements under N₂, the same gas flow rate was employed in order to obtain comparable data. With the intention of investigating this effect further, similar mass transport measurements were carried out with the same set-up and parameters but under stagnant N₂ atmosphere where Ga mass transport rate was measured to be 2.6x10⁻² g/h. The difference between this value and that measured under nitrogen and ammonia flows with the filament assisted source show the effect of 700 sccm gas flow rate on the mass transport rate.

One drawback of heating the source surface with a filament is that due to a small source-to-substrate distances the filament affects the temperature of the substrate surface significantly, which in turn affects the crystals quality.

In order to obtain a stable and high transport of Ga species at moderate temperatures, a source holder design was employed as explained in the experimental procedure section. The results of Ga mass loss rate measurements under different flow rates of N_2 and NH_3 gases are presented in Fig. 10. In all experiments 1 slm of N_2 was flowed over the molten Ga source. After each experiment a clean Ga source surface, free of GaN crust, was obtained, indicating that the NH_3 – liquid Ga surface reactions were suppressed. An enhanced evaporation rate of about 40 mg/h was obtained at a source temperature of 1250°C. Ga mass loss rate under different flow rates of N_2 show the consistency of the collected data and stability of the experimental set-up.

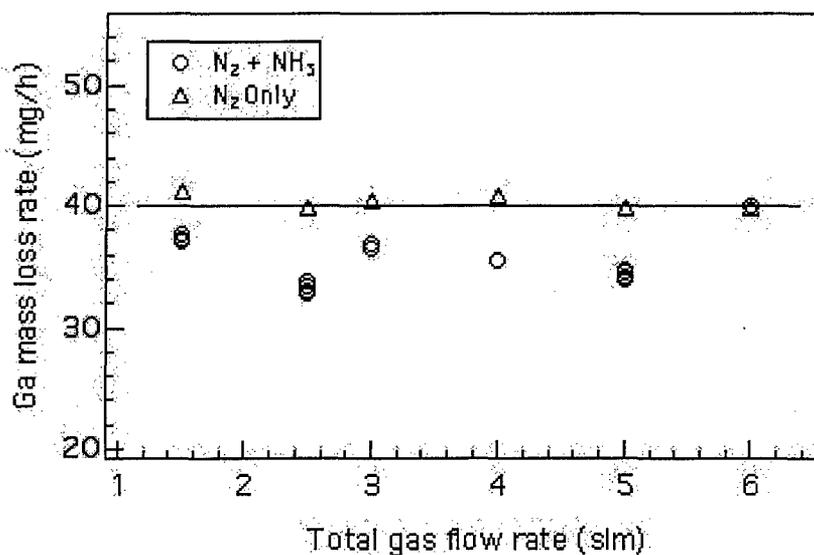


Figure 10. Ga mass loss rate as a function of total flow rate

When the combination of N_2 and NH_3 gases was flowed, lower Ga mass loss rate with some deviations was observed. The lower Ga mass loss rate obtained in this case could be a result of GaN deposition on the source holder and its effect on the mass loss rate measurements. The deviations could be controlled through keeping the total flow rate constant by using a gas mixture of appropriate amounts of NH_3 and N_2 in consecutive experiments. Results of these experiments are shown in Fig. 11. The plot shows that the Ga mass loss rate and in turn the transport of Ga species could be controlled and predicted for different flow rates of NH_3 when the total flow rate of gases was kept constant. This allows the control of V/III ratio in the reactor in different experiments.

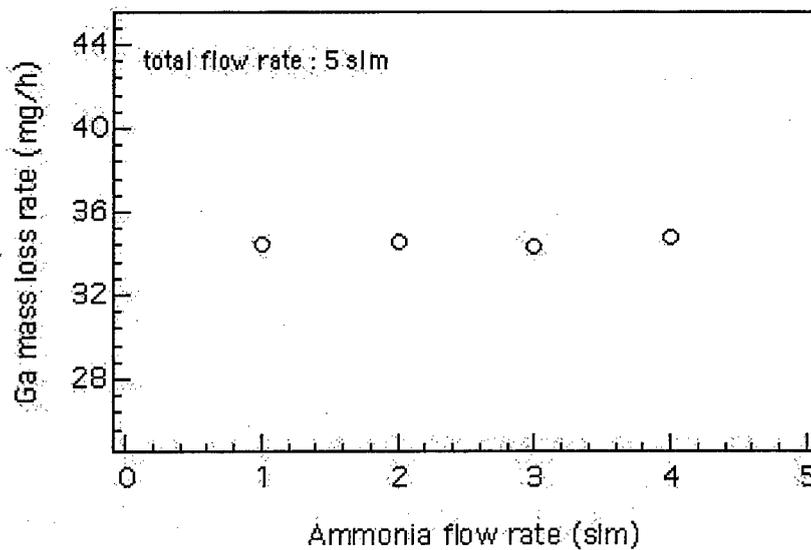


Figure 11. Ga mass loss rate as a function of NH_3 flow rate at constant total flow

4.6. Summary

The transport of Ga species from source-to-substrate in a vapor growth system under different conditions was studied and experimental results were compared with respect to theoretical calculations. Despite the simplifications used in the theoretical estimations, calculated results show good agreement with the values measured experimentally. It was observed and quantified that under the NH_3 atmosphere, Ga transport does not occur via the simple vapor phase transport. The use of a hot filament above the molten Ga resulted in suppression of Ga droplet formation, a stable source and source surface free from GaN crust. It was concluded that in order to obtain a steady vapor phase transport of Ga species, the contact between the molten Ga and NH_3 must be avoided. It was shown experimentally that the novel source design prevented NH_3 – molten Ga reaction, and Ga spattering and GaN crust formation on the source surface. A significantly enhanced Ga evaporation rate was achieved and it became possible to control Ga transport and V/III ratio in the system.

5. GROWTH OF GaN FROM ELEMENTAL GALLIUM AND AMMONIA VIA A MODIFIED SANDWICH GROWTH TECHNIQUE

5.1. Outline

Gallium nitride (GaN) thin films were grown on (0001) sapphire substrates at 1050°C by controlled evaporation of gallium (Ga) metal and reaction with ammonia (NH₃) at a total reactor pressure of 800 Torr. Pure nitrogen (N₂) was flowed directly above the molten Ga source. This flow prevented direct reaction between the molten Ga source and ammonia, which causes Ga spattering and GaN crust formation, substantially enhanced Ga evaporation rate and, at the same time, enabled control of Ga transport to the substrate. A growth rate of 1.4 μm/h was obtained at 1050°C, 800 Torr, 3 slm of ammonia, and 1250°C Ga source temperature with a 24 mm source-to-substrate distance. XRD measurements revealed a full-width at half-maximum (FWHM) of 0.6 deg. for the (0002) GaN peak. Experimental variables and improvements in the overall growth process regarding the transport of species and the quality of the grown crystals are presented.

5.2. Introduction

GaN and its alloys with AlN and InN recently became the basic materials for short-wavelength optoelectronics, as described in Chapter 2, however, the development of GaN-based technology is still strongly limited by difficulties in obtaining high-quality, large-area native substrates in large quantities. Due to the challenges in the growth of

GaN crystals, novel techniques were employed as described in Section 2.2. Among the widely used techniques in GaN bulk or thick film growth, many advantages of the sandwich growth method have been reported and summarized in Chapter 2.

In this chapter, the growth of GaN on sapphire with a modified sandwich growth method was investigated. Results regarding the stability of the growth process and the quality of grown crystals are discussed.

5.3. Experimental

Experiments were conducted in the system described in Chapter 3. Crystal growth of GaN was performed by evaporating high purity metallic Ga (99.999%) in a stream of UHP-grade N₂ and NH₃. In all experiments, the growth temperature was 1050°C while the total pressure and NH₃ flow rate were varied from 50 to 800 Torr and 1 to 4 slm, respectively. In order to obtain stable and high Ga vapor phase transport at moderate temperatures, a novel source design was employed, as explained in Chapters 3 and 4. In all experiments 1 slm of N₂ was flowed over the molten Ga source. Experimental conditions are summarized in Table 4.

Table 4. Summary of experimental conditions for GaN growth

total flow rate (slm)	NH ₃ flow rate (slm)	N ₂ flow rate (slm)	N ₂ (source) flow rate (slm)	P _T (Torr)	T _{substrate} (°C)	T _{source} (°C)
5	1 - 4	4 - 1	1	50 - 800	1050	1250

Optical and scanning electron microscopes (SEM) were used to characterize the crystal surface morphology. X-ray diffraction (XRD) analysis was used to determine crystallinity of the deposited films. X-ray energy dispersive analysis (EDX) was

employed for chemical characterization and Photoluminescence (PL) technique was used to evaluate the optical crystal quality.

5.4. Results and Discussion

5.4.1. Stable Ga source

In all experiments, 1 slm of N_2 was flowed over the molten Ga source. After each of these experiments a clean Ga source surface, free of GaN crust, was obtained indicating that the NH_3 – liquid Ga surface reactions were suppressed [101]. An evaporation rate of about 35 mg/h was obtained at a source temperature of 1250°C. The Ga mass loss rate was kept constant by keeping the N_2 flow rate over the source constant.

Ga mass loss rate in the system depended on the source temperature, N_2 flow rate above the molten Ga and the total pressure. The main temperature dependence is due to the exponential dependence of the equilibrium Ga vapor pressure on temperature. The effect of N_2 (source) flow rate and the total pressure on the Ga mass flow rate is similar to that in precursor delivery systems used in CVD processes to introduce vapors of source reagents to the CVD tool. It was found that the Ga mass loss rate was directly proportional to the N_2 (source) flow rate, but inversely proportional to the total pressure, P_T [102, 103]. This behavior is plotted in Figure 12.

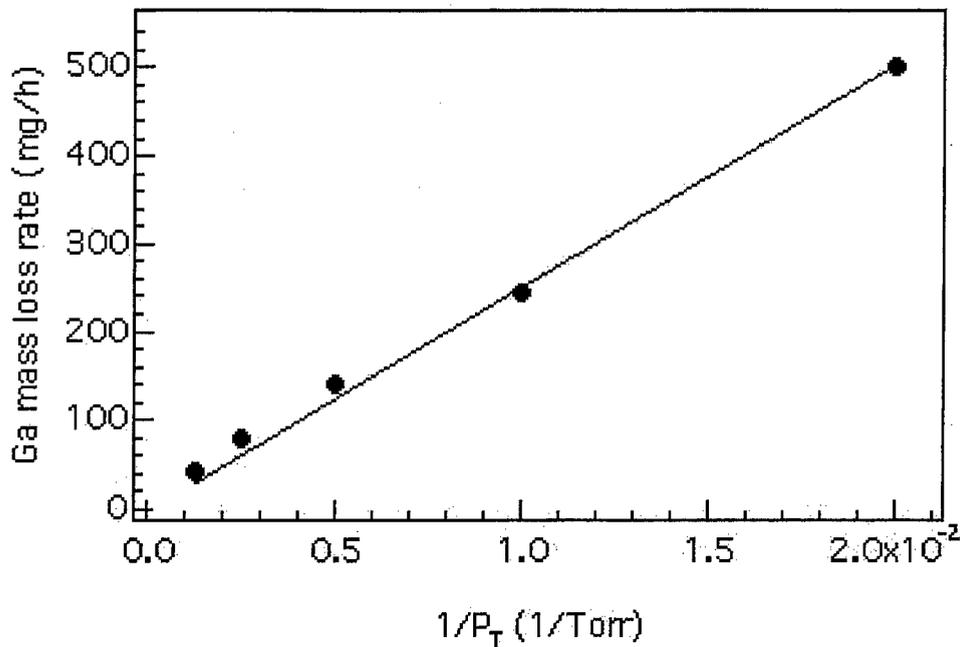


Figure 12. Plot of Ga mass loss rate vs. $1/\text{total pressure}, 1/P_T$

5.4.2. Growth Morphology and Kinetics

Figure 13 shows the influence of the ammonia flow rate which represents the effect of V/III ratios on crystal growth. GaN deposited under NH_3 flow of 1 slm consist of hexagonal islands. These islands tend to grow more laterally with increasing V/III ratio, compare Figures 13(b) and 13(c). A further increase of the V/III ratio results in growth of polycrystalline GaN.

A growth rate of $1 \mu\text{m/h}$ was obtained at 1050°C , 800 Torr, 3 slm of ammonia, 1250°C Ga source temperature, a V/III ratio of 10000 and a 24 mm source-to-substrate distance. The growth rate increased slightly with increasing NH_3 flow rate, as shown in Fig. 14.

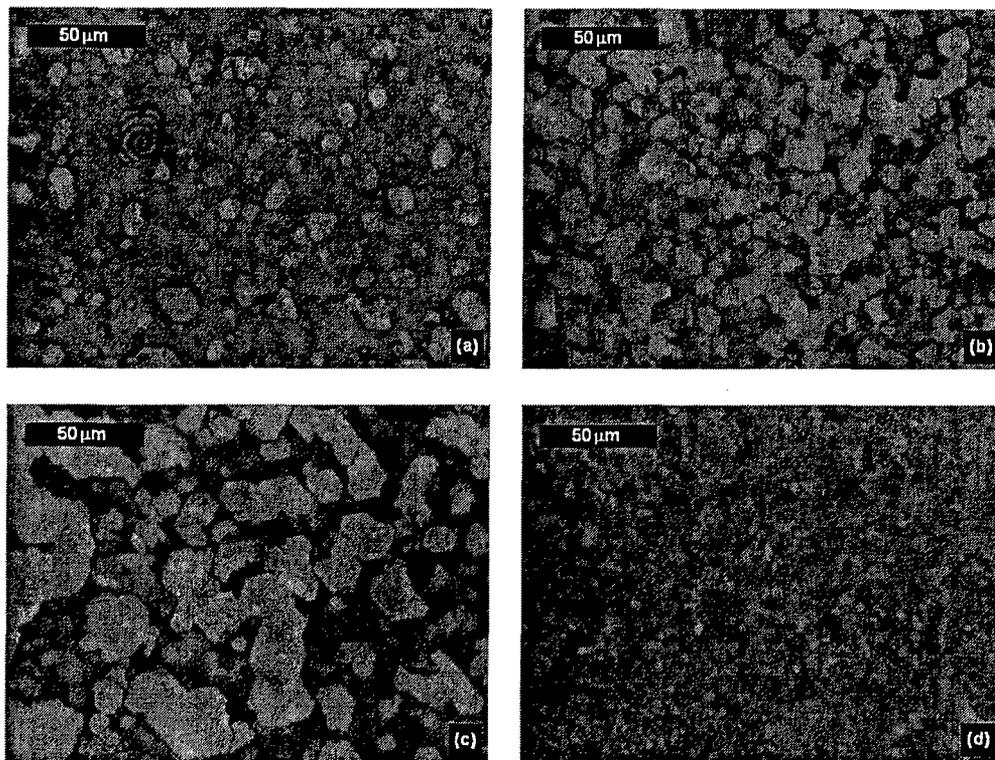


Figure 13. Surface morphologies of GaN deposited directly on sapphire under different NH_3 flow rates (a) 1 slm, (b) 2 slm, (c) 3 slm, (d) 4 slm.

The growth rates obtained at different total pressures are presented in Fig. 15. The deposition rates up to $10 \mu\text{m/h}$ were obtained at a total reactor pressure of 50 Torr. It was also observed that the ratio of deposition rate to Ga mass loss rate for each experiment remained constant and evaporation normalized growth rate was found to be independent of the total reactor pressure. A detailed study on this subject together with the analysis of growth characteristics are covered in Chapter 6.

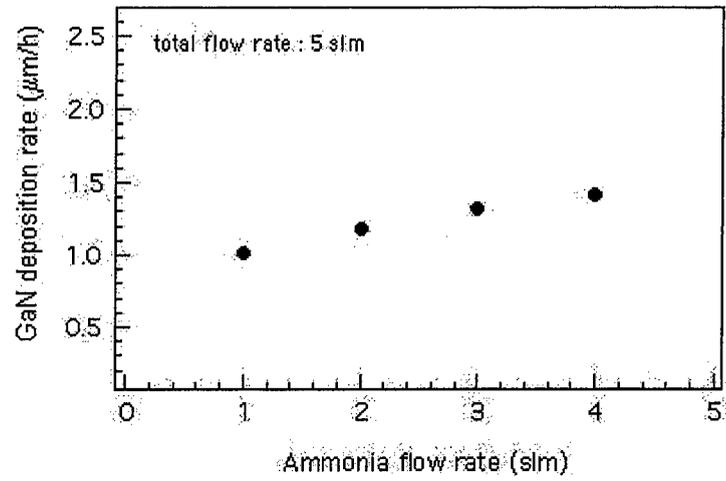


Figure 14. GaN deposition rate versus NH_3 flow rate

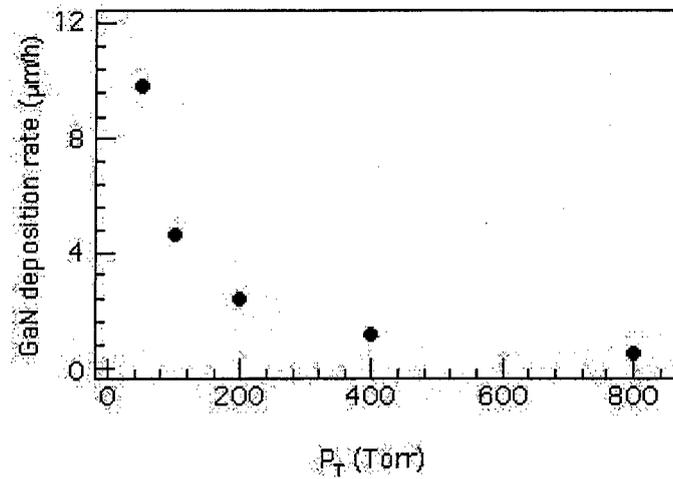


Figure 15. Plot of GaN deposition rate versus total pressure, P_T

5.4.3. Crystal Characterization

XRD analysis of the grown samples was performed using CuK_α line and a representative result is presented in Fig. 16. The degree of in-plane crystallographic orientation was measured by the width of the diffraction peak in a ω -scan. A FWHM of 0.6° was obtained for the diffraction peak in a χ -scan, Fig. 17.

The EDX chemical analysis result presented in Fig. 18 shows only Ga and N peaks; the contamination is below the detection limits of the technique, 0.1 at. % [104].

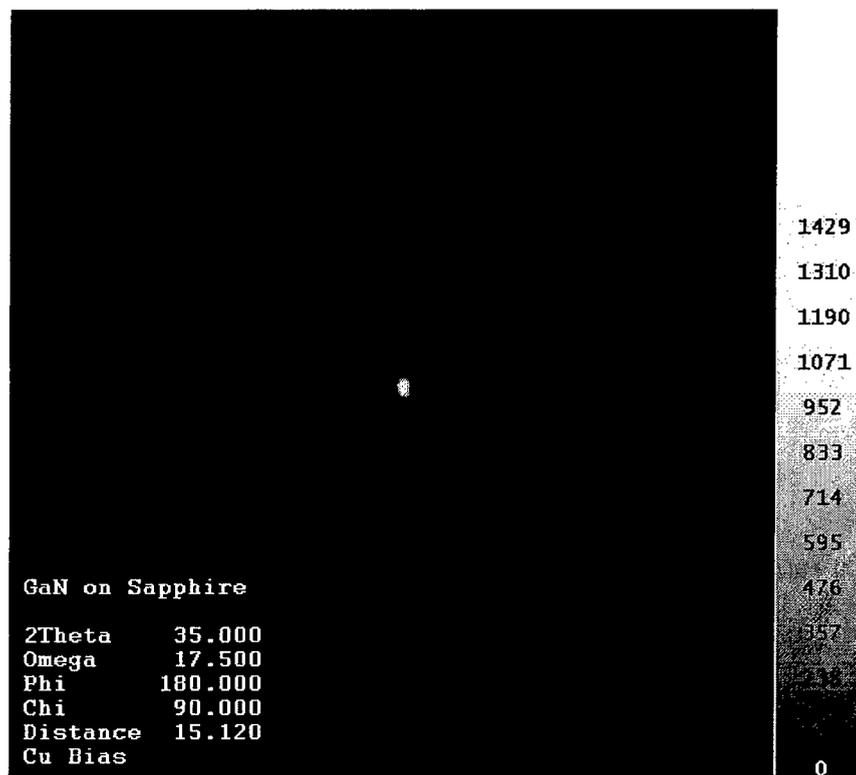


Figure 16. XRD analysis of a (0002) GaN epilayer deposited at 1050°C on *c*-plane sapphire

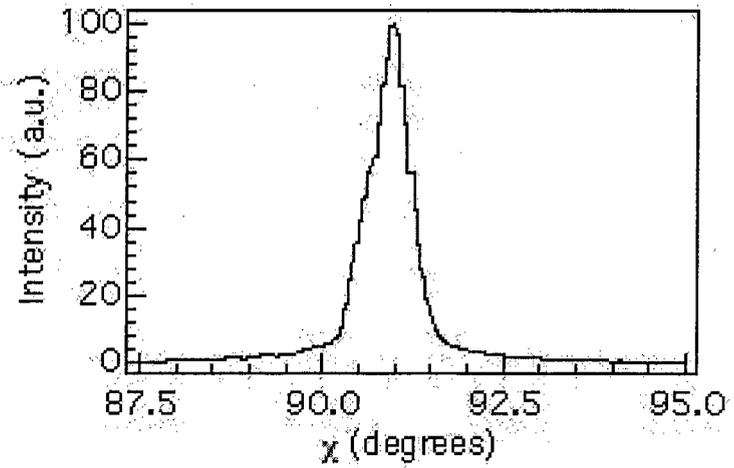


Figure 17. χ -scan X-ray diffraction of (0002) GaN epilayer deposited at 1050°C on c-plane sapphire

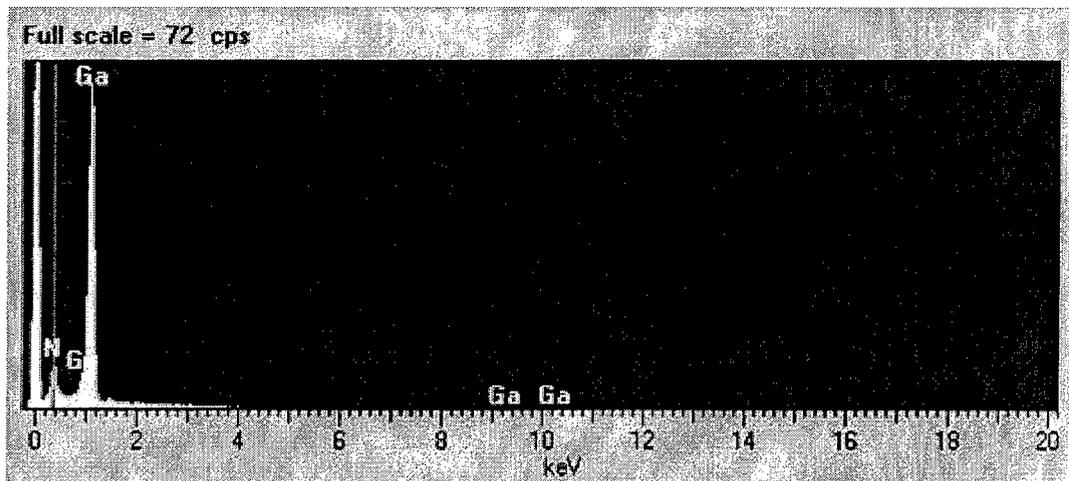


Figure 18. EDX of GaN epilayer deposited at 1050°C on c-plane sapphire.

Fig. 19 shows a typical room temperature PL spectrum of GaN samples. A HeCd laser with a wavelength of 325 nm was used to excite the samples. Strong emission at 367 nm (~ 3.4 eV) and an FWHM of 28 nm (255 meV) was observed. The yellow luminescence band is thought to be related to non-stoichiometric point defects [105].

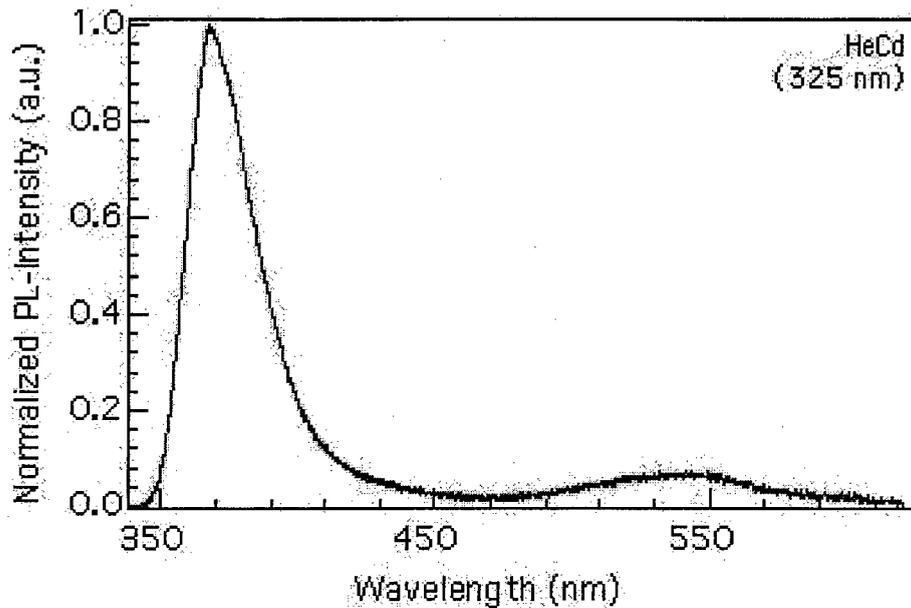


Figure 19. Photoluminescence spectrum of GaN epilayer deposited at 1050°C on c-plane sapphire

5.5. Summary

Gallium nitride films were grown on (0001) sapphire substrates at 1050°C by controlled evaporation of Ga metal and reaction with NH_3 . Surface morphology of the deposited layers was examined by optical and scanning electron microscopy. XRD analysis was used to determine crystallinity of the deposited films and revealed that the

grown material was single crystalline and epitaxially grown on the sapphire wafer. EDX analysis was employed for the chemical characterization of the samples and showed that the deposited material contained only Ga and N elements. Room temperature PL spectrum demonstrated the optical quality of the grown samples. The yellow band luminescence was attributed to the non-stoichiometric point defects such as N vacancies.

This work demonstrates the feasibility of the process for the growth of GaN epilayers. However, the process needs further improvement which would allow more efficient delivery of NH_3 to the substrate region in order to increase the crystal quality and the growth rate. A more detailed discussion on this subject can be found in Chapter six.

6. MASS TRANSFER IN GaN GROWTH FROM ELEMENTAL GALLIUM AND AMMONIA VIA A MODIFIED SANDWICH GROWTH TECHNIQUE

6.1. Outline

A one-dimensional mass transport model based on total reactor pressure, gas flow rates and source temperature was developed for the growth of GaN with a modified sandwich growth technique and verified with growth results. Theoretical calculations as well as growth rate measurements for different ammonia flow rates and source-to-substrate distances showed that the maximum growth rate was controlled by transport of both Ga species and reactive ammonia to the substrate surface. A growth rate of 1.4 $\mu\text{m/h}$ was obtained at 1050°C, 800 Torr, 3 slm of ammonia, and 1250°C Ga source temperature with a 24 mm source-to-substrate distance.

6.2. Introduction

The lack of III-Nitride substrates for homoepitaxial growth has been and remains to be the main technological issue in the development of III-Nitride based devices. For the bulk growth of GaN, high crystal growth rates and controlled crystal growth over a large area need to be achieved.

Among the various bulk/thick crystal growth techniques, vapor phase transport methods have been reported to be advantageous due to the use of the simplest source species, high yield of the growth components and high growth rates [92]. In the vapor

phase transport methods, understanding of the growth process is important for the control of the operating conditions as well as the cost of the process which makes the growth modeling an important tool. Significant progress in mathematical modeling of the vapor phase transport methods for bulk SiC growth have been reported in the literature [106-109]. Similar growth methods for bulk growth of AlN became subject of recent studies and many transport and thermodynamic studies on the sublimation growth of AlN have been reported [64, 110-115]. Encouraged by these achievements, analogous growth methods were employed to the GaN growth with slight modifications as explained in Section 2.2. Baranov et al. [60] have studied the gas dynamics and the supply of the main gaseous species into the growth cell in a reactor for GaN crystal growth by a sublimation sandwich technique (SST). More recently, Wu et al. [116] reported on the physical modeling of GaN growth using SST.

In this chapter, GaN growth characteristics and species transport properties were investigated in order to determine the GaN growth rate and consequently optimize the experimental parameters to obtain higher growth rates. Results regarding the growth conditions and the validity of the calculations were discussed.

6.3. Growth Characteristics

In the growth process, described in Chapter 3, Ga metal evaporates and travels towards the substrate by a forced flow of N_2 (source) as shown in Figure 20.

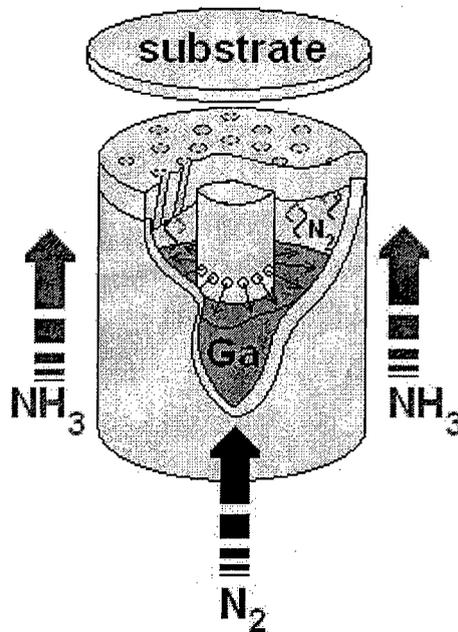


Figure 20. Schematic illustration of the modified sandwich configuration

Experimental results at different total pressures are presented in Fig. 21(a) and (b). It was observed that the ratio of deposition rate to Ga mass loss rate in each experiment remained constant. This showed that the process was mass transport limited.

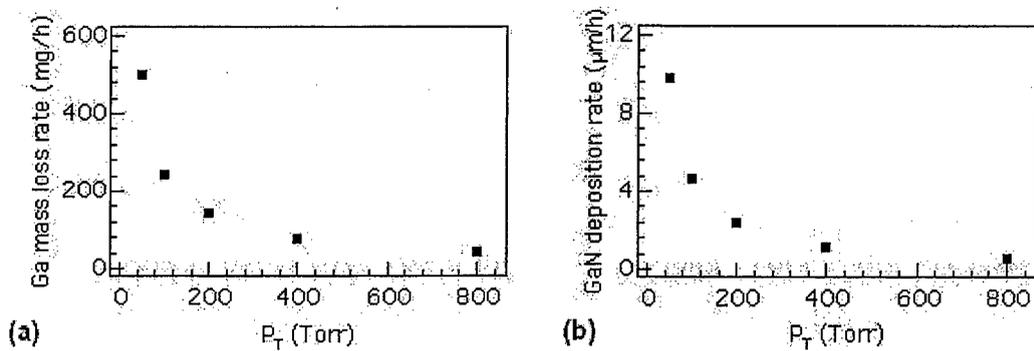


Figure 21. (a) Plot of Ga mass loss rate versus total pressure, P_T (b) GaN deposition rate versus total pressure, P_T

Characteristic number calculations according to the experimental parameters and chamber geometry revealed that the state of the gas was viscous and the flow laminar (see Appendix A). In the viscous flow regime, the flow velocity drops rapidly to zero across the boundary layer near the substrate surface. According to the boundary layer approximation, the flow is nearly stagnant within the boundary layer. The mixing of gases is assumed to occur beyond the boundary layer, in the main gas stream. At the boundary layer edge, the reactant partial pressure is equal to P_i , where the reactants are Ga and NH_3 , and decreases toward the surface until it reaches an equilibrium value. This provides a concentration gradient $\Delta P_i / \delta_o$ that is the driving force for the diffusion of the i^{th} specie. Consequently, the growth rate \dot{G}_{GaN} in the system can be formulized in a linear approximation as

$$\dot{G}_{\text{GaN}} \approx K(T) D_i \frac{\Delta P_i}{\delta_o} \quad (6)$$

where D_i is the diffusivity of reactants, δ_o the boundary layer thickness and $K(T)$ includes temperature dependent terms. Diffusivity in the boundary layer is inversely proportional to the total pressure P_T and the boundary layer thickness depends on the kinematic viscosity and gas velocity as

$$\delta_o = \left(\frac{\nu}{u} \right)^{1/2} \quad (7)$$

$$\text{Kinematic viscosity: } \nu = \left(\frac{\eta}{\rho} \right) \quad \eta: \text{dynamic viscosity} \quad (8)$$

ρ : gas density

$$\text{Gas velocity: } u = \frac{Q}{P_T A}, \quad (9)$$

where Q is the throughput, P_T the total reactor pressure and A the cross sectional area. Since both the kinematic viscosity and the gas velocity are inversely proportional to the total pressure P_T , it can be predicted from Eq (7) that the boundary layer thickness is independent of the total pressure. Assuming complete depletion of the reactants at the substrate surface, i.e. $P_i^{eq} = 0$ and $\Delta P_i \cong P_i$, and relating the velocity of the gas and the partial pressure of the different species to the flow rates, the growth rate becomes

$$\dot{G}_{GaN} \approx K(T) \frac{f_i}{\sqrt{f_T}} \quad \text{or} \quad \approx K(T) \frac{P_i}{P_T} \sqrt{f_T} \quad (10)$$

where f_i is the flow rate of the i^{th} specie, f_T is the total flow rate and $K(T)$ includes temperature dependent terms. An evaporation rate-normalized growth rate defined as in Eq. (11) reveals that at a constant growth temperature and total flow rate, the normalized growth rate is independent of total pressure. The experimental verification of this behavior is presented in Fig. 22.

$$\bar{G}(P) \approx \frac{\dot{G}_{GaN}}{f_i} \approx \frac{1}{\sqrt{f_T}} \quad (11)$$

The model predicts that the maximum growth rate can be controlled by the transport of both Ga species and reactive ammonia to the substrate surface. However, the key problem in improving the crystal growth rate in present reactor was found to be the supply of sufficient amount of active nitrogen species to the substrate. As known, decomposition of NH_3 is a mildly endothermic process and thermodynamically almost all the NH_3 decomposes into N_2 and H_2 at temperatures higher than 300°C [117].

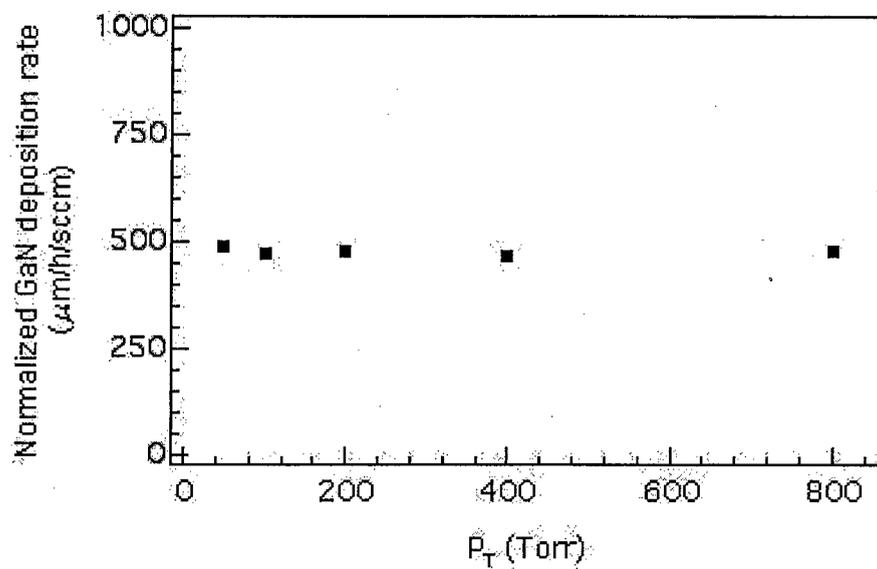


Figure 22. Plot of evaporation normalized growth rate versus total pressure, P_T

It is also well known that NH_3 decomposition is a catalytic process [118] and the extent of the decomposition strongly depends on the growth conditions and the equipment. In the present growth system, alumina tube walls, BN and Mo-parts act as catalysts for the NH_3 decomposition until it reaches the substrate. They change the activation energy of decomposition by providing an alternate pathway for the decomposition reaction. The decomposition rate has an exponential dependence on temperature determined by the Arrhenius equation. A heater temperature of 1650°C was employed to reach a source temperature of 1250°C . This resulted in a temperature of around 1400°C on the inner surface of the alumina tube at which the extent of NH_3 decomposition is enormous [119]. Optical microscopy characterization of the deposited films revealed that as the deposition rate increases by increasing reactant fluxes to the substrate at a constant V/III ratio, crystal color becomes darker, and in severe cases,

results in Ga rich needle formation. A more detailed study on ammonia decomposition in the growth system can be found in Appendix C.

6.4. Summary

Gallium nitride films were grown on (0001) sapphire substrates at 1050°C by controlled evaporation of Ga metal and reaction with NH₃ via a modified sandwich technique. Growth characteristics were described by a mass transport model based on process parameters and experimentally verified. Results showed that the maximum growth rate was controlled by the transport of both Ga species and reactive ammonia to the substrate surface. A growth rate of 1.4 μm/h was obtained at 1050°C, 800 Torr, 3 slm of ammonia, and 1250°C Ga source temperature at a 24 mm source-to-substrate distance. It was found that the process required a more effective supply of NH₃ to the substrate in order to increase the crystal quality and the growth rate.

7. SUMMARY AND FUTURE WORK

This work mainly focused around the process development for the growth of GaN crystals from elemental Ga and NH_3 . One of the biggest challenges of working with Ga – NH_3 system was the instability of molten Ga at elevated temperatures, especially between 1100 – 1200°C. Preliminary experiments revealed that the eruptions on the molten Ga surface caused spattering of the Ga and the formation of macro-sized Ga droplets on the growth surface. The contact of a spattered Ga droplet with an already formed GaN crystal disturbed the crystal growth. In addition, it was observed in the same experiments that the growth eventually ceased due to the formation of GaN crust on the molten Ga source surface. Therefore, in the first part of this study, transport characteristics of Ga species from source-to-substrate during the GaN growth process and the role of ammonia-liquid Ga reaction on the transport were investigated. Experimental results under different conditions were studied and compared to theoretical predictions to quantify the mechanism of transport in the vapor growth technique. It was observed and quantified that under the NH_3 atmosphere Ga transport occurred predominantly by cluster rather than vapor phase transport.

In this work, the majority of the effort has been spent on a growth cell design that could prevent the direct reaction between the molten Ga and NH_3 and enable reproducible and stable growth conditions. It was observed that in order to obtain steady vapor phase transport of Ga species the contact between the molten Ga and NH_3 must be avoided. The use of a hot filament above the Ga resulted in suppressed Ga droplet formation, stable source and source surface free of GaN crust, which also supports the

hypothesis. One drawback of this approach was that due to small source-to-substrate distances, the filament significantly affected the temperature of the substrate surface, which in turn affected the crystal quality.

Therefore, a novel Ga source design was employed in an effort to obtain a stable and high vapor phase transport of Ga species at moderate temperatures. In this design, pure N_2 was flowed directly above the molten Ga source. This flow prevented the direct contact between the molten Ga and NH_3 , Ga spattering, and GaN crust formation on the source surface. A significantly enhanced Ga evaporation rate was achieved and it became possible to control Ga transport and V/III ratio in the system.

Gallium nitride films were grown on (0001) sapphire substrates at $1050^\circ C$ by controlled evaporation of Ga metal and reaction with NH_3 . Experimental results under different ammonia flow rates demonstrated the effect of V/III ratio on the crystal growth. It was observed that the hexagonal GaN islands obtained under 1 slm of NH_3 flow rate tended to grow more laterally with increasing NH_3 flow rate at a constant Ga evaporation rate. Further increase in V/III ratio resulted in polycrystalline growth of GaN. The surface morphology of the deposited layers was examined by optical and scanning electron microscopies. XRD analysis was used to determine crystallinity of the deposited films and revealed that the grown material was single crystalline and epitaxially grown on the sapphire wafer. EDX analysis was employed for the chemical characterization of the samples and showed that the deposited material contained only Ga and N elements. The room temperature PL spectrum demonstrated the optical quality of the grown samples. The observed yellow band luminescence was attributed to the non-stoichiometric point defects such as N vacancies.

Growth characteristics were described by a mass transport model based on process parameters and experimentally verified. The results showed that the process was mass transport limited and the maximum growth rate was controlled by the transport of both Ga species and reactive ammonia to the substrate surface. A growth rate of 1.4 $\mu\text{m/h}$ was obtained at 1050°C, 800 Torr, 3 slm of ammonia flow rate, and 1250°C Ga source temperature at a 24 mm source-to-substrate distance. It was found that the process required effective supply of NH_3 to the substrate in order to increase the crystal quality and the growth rate.

This work demonstrates and discusses the feasibility of the process for growth of GaN from elemental Ga and NH_3 . The key challenge in improving the crystal growth rate in this system was found to be the supply of sufficient amount of active nitrogen species to the substrate. Calculations of the catalytic NH_3 decomposition rate revealed that the extent of decomposition strongly depended on the growth conditions and equipment geometry. Alumina tube walls, BN and Mo-parts act as catalysts for NH_3 decomposition. The decomposition rate had an exponential dependence on temperature determined by the Arrhenius equation. In the growth process a heater temperature of 1650°C was employed to reach a source temperature of 1250°C. This resulted in a temperature of around 1400°C on the inner surface of the alumina tube at which the extent of NH_3 decomposition is enormous.

Optical microscopy characterization of the deposited films revealed that as the deposition rate increased by increasing reactant fluxes to the substrate at a constant NH_3/Ga ratio, crystal color became darker until a Ga rich needle formation was observed. The process needs further improvement which would allow the supply of active NH_3 to

the substrate region in order to achieve the growth of high quality thick GaN layers of high growth rates.

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APPENDIX A: REACTOR MODELLING

A.1. General Assumptions

Based on the operating conditions and the system geometry presented in Fig. 23, the characteristic numbers were defined.

A.1.1. State of the Gas

The gas in a vacuum system can be in viscous state, in molecular state or in a state which is intermediate between these two. When a system is brought from the atmospheric pressure to “high vacuum”, the gas in the system goes through all these states. The mean free path λ of the gas molecules is small at atmospheric pressure [69] so that the flow of gas is limited by its viscosity. At low pressures, where the mean free path of the molecules is similar to the dimensions of the vacuum enclosure, the flow of gas is governed by molecular phenomena; this is the intermediate flow. At very low pressures where the mean free path is much larger than the dimensions of the vacuum enclosure, the flow is molecular.

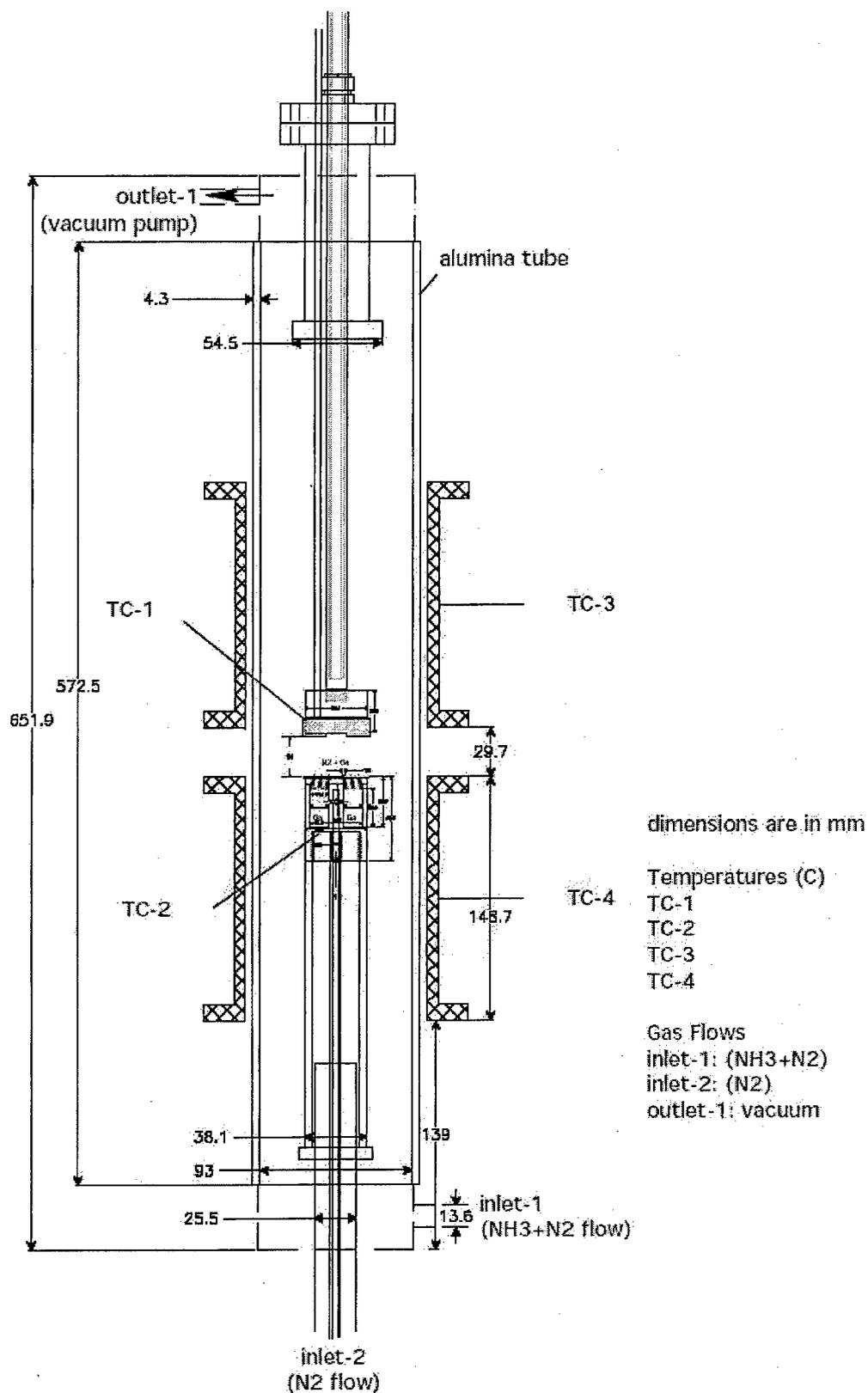


Figure 23. Drawing of the alumina reactor tube

The state of gas is described by the Knudsen number described as

$$K_n = \frac{\lambda}{L} \quad (12)$$

$$\lambda = \frac{1}{\sqrt{2\pi\alpha^2 n}}, \quad (13)$$

where L is the characteristics distance, λ mean free path, α molecular diameter and n the number of molecules per unit volume. The viscous flow regime is operative when gas transport occurs at pressures above 20 Torr in deposition reactors of typical size [75]. For the GaN growth system studied in this work, the mean free path was calculated as 0.3 μm ; for the source-to-substrate distance of 2.5 cm as the characteristic dimension of the system, The Knudsen number was estimated as $\sim 10^{-5}$ revealing that the state of gas was viscous.

A.1.2. Flow Regime

In the state where the state of gas is viscous, the flow can be turbulent or laminar. The flowing fluid experiences two forces: inertial and viscous. The viscous forces tend to stabilize the unidirectional motion of the flow, while the inertial forces tend to disrupt the unidirectional motion. Consequently, the nature of the flow is determined by the relative magnitudes of the two forces. At low flow velocities the viscous forces dominate and the stabilizing influence causes laminar flow. However, with increasing velocity, tendency toward disruption of laminar flow increases.

The Reynold's number, the ratio of inertial forces to viscous forces, is a dimensionless quantity expressed by

$$\text{Re} = \frac{\rho v D}{\eta}, \quad (14)$$

where ρ is the density of gas, v the velocity, η the dynamic viscosity, and D the diameter of the tube. For Reynold's numbers larger than 2100, the inertial effects are larger and the flow is entirely turbulent, while for $\text{Re} < 1100$, the viscous effect is larger and the flow is entirely laminar. The exact value of Re for which the flow changes from turbulent to laminar depends upon the roughness of the surface of the tube and other experimental factors.

The expression of the Reynold's number can be related to the throughput Q which is defined as the quantity of gas flowing through a pipe, expressed in PV (pressure _ volume) units per unit time. Thus

$$Q = P v \frac{\pi D^2}{4} \quad \text{or} \quad v = \frac{4}{P \pi D^2} Q \quad (15)$$

The ideal gas law is an appropriate equation of state relating the temperature, pressure and density at typical vapor deposition conditions. Thus, the gas density is equal to

$$\rho = n m \quad (16)$$

where n , the number of molecules per unit volume, equals to

$$n = \frac{N_A P}{R T} \quad (17)$$

and m , the mass of a molecule, is

$$m = \frac{M}{N_A} \quad (18)$$

where N_A is the Avogadro's number, M the molecular weight and R the universal constant ($R=62.364$ Torr.liter/K.mole). From Eqs. 16 to 18, it results that

$$\rho = \frac{MP}{RT} \quad (19)$$

and the expression for the Reynold's number, Eq. 14, can be written as

$$\text{Re} = \frac{MP}{RT} \frac{4Q}{P\pi D^2} \frac{D}{\eta} = \frac{4M}{\pi RT\eta} \left(\frac{Q}{D} \right) \quad (20)$$

If Q is expressed in sccm, T in K, η in $\mu\text{Pa}\cdot\text{sec}$, D in meters, according to Eq. 20

$$\text{Re} = 0.2588 \frac{M}{T\eta} \left(\frac{Q}{D} \right) \quad (21)$$

For the operating conditions used in the experiments, the Reynold's number was estimated to be less than 25 confirming the laminar viscous characteristics of the gas flow in the system.

As a result of the characteristic number calculations, the following assumptions were made (which do not essentially limit the accuracy or the applicability of the model):

- (a) The gas mixture is treated as a continuum;
- (b) The gases are treated as ideal; and they obey the ideal-gas law and the Newton's law of viscosity;
- (c) The gas flow is laminar;
- (d) The gas mixture in the reactor is transparent for heat radiation; and

A.2. Basic Equations

Models of crystal growth from the vapor consist of nonlinear, coupled partial differential equations that represent the conservation of momentum, energy, total mass and individual species. The general derivation and form of these equations is given in

references [71-73]. Numerous modifications of these equations have been used in models ranging from simple boundary-layer-type descriptions to complete three-dimensional models. The general equations are applicable to most crystal growth systems operating in the continuum regime. Since the reactor geometry is axisymmetric, the governing equations of the problem were written in cylindrical coordinates and in two dimensions (r in the radial direction and x in the axial direction). The governing equations of mass, momentum, energy and species concentrations and the ideal gas law are described as follows:

Total Mass:

$$\frac{\partial}{\partial x}(\rho u) + \frac{1}{r} \frac{\partial}{\partial r}(r \rho v) = 0 \quad (22)$$

Momentum:

$$\begin{aligned} \frac{\partial}{\partial x}(\rho u u) + \frac{1}{r} \frac{\partial}{\partial r}(r \rho u v) &= \frac{\partial}{\partial x} \left(\mu \frac{\partial u}{\partial x} \right) + \frac{1}{r} \frac{\partial}{\partial r} \left(r \mu \frac{\partial u}{\partial r} \right) \\ &\quad - \frac{\partial P}{\partial x} + \frac{\partial}{\partial x} \left(\mu \frac{\partial u}{\partial x} \right) + \frac{1}{r} \frac{\partial}{\partial r} \left(r \mu \frac{\partial v}{\partial x} \right) \\ &\quad - \frac{2}{3} \mu \nabla \cdot \vec{V} + F_x \end{aligned} \quad (23a)$$

$$\begin{aligned} \frac{\partial}{\partial x}(\rho u v) + \frac{1}{r} \frac{\partial}{\partial r}(r \rho v v) &= \frac{\partial}{\partial x} \left(\mu \frac{\partial v}{\partial x} \right) + \frac{1}{r} \frac{\partial}{\partial r} \left(r \mu \frac{\partial v}{\partial r} \right) \\ &\quad - \frac{\partial P}{\partial r} + \frac{\partial}{\partial x} \left(\mu \frac{\partial u}{\partial x} \right) + \frac{1}{r} \frac{\partial}{\partial r} \left(r \mu \frac{\partial v}{\partial r} \right) \\ &\quad - \frac{2}{3} \mu \nabla \cdot \vec{V} - \mu \frac{2v}{r^2} + F_r \end{aligned} \quad (23b)$$

Energy:

$$\frac{\partial}{\partial x}(\rho u h) + \frac{1}{r} \frac{\partial}{\partial r}(r \rho v h) = \frac{\partial}{\partial x} \left(\frac{k}{C_p} \frac{\partial h}{\partial x} \right) + \frac{1}{r} \frac{\partial}{\partial r} \left(r \frac{k}{C_p} \frac{\partial h}{\partial r} \right) \quad (24)$$

Species:

$$\frac{\partial}{\partial x}(\rho u Y_i) + \frac{1}{r} \frac{\partial}{\partial r}(r \rho v Y_i) = \frac{\partial}{\partial x} \left(D_i \frac{\partial Y_i}{\partial x} \right) + \frac{1}{r} \frac{\partial}{\partial r} \left(r D_i \frac{\partial Y_i}{\partial r} \right) \quad (25)$$

Ideal gas law:

$$P = \rho R T \sum_i \frac{Y_i}{W_i} \quad (26)$$

The integro-differential radiative heat transfer equation:

$$\begin{aligned} (\Omega \cdot \nabla) I(r, \Omega) = & -(\kappa + \sigma) I(r, \Omega) + \kappa I_b(r) \\ & + \frac{\sigma}{4\pi} \int_{\Omega'=4\pi} I(r, \Omega') \Phi(\Omega' \rightarrow \Omega) d\Omega \end{aligned} \quad (27)$$

where u is the axial velocity, v the radial velocity, ρ the density, μ the dynamic viscosity, F_x and F_r the gravitational acceleration vectors in the x and r co-ordinates, respectively, P the pressure, h the enthalpy ($h=C_p T$), k the thermal conductivity, C_p the specific heat, D_i the binary diffusion coefficient of reactants in carrier gases, Y_i the mass fraction and W_i the molecular weight of the i^{th} specie. In the radiative heat transfer equation, Ω is the direction of propagation of the radiation beam, I the radiation intensity, κ the absorption coefficient, σ the scattering coefficient, I_b the intensity of the black body, Φ the phase function of the energy transfer from the incoming Ω' direction to the outgoing direction Ω .

The solution of the conservation equations for total mass (Eq. 22) and momentum (Eqs. 23a and 23b) gives the pressure and velocity distribution throughout the reactor enclosure. The temperature distribution is given by the solution to the energy balance (Eq. 24).

A.3. Boundary Conditions

Boundary conditions must be specified for the above equations, in order to simulate the reactor configuration. The following boundary conditions have been specified.

Walls of the reactor tube surrounded by the graphite heaters:

$$\begin{aligned}
 \vec{V} &= 0 \\
 h &= h_{\text{graphite}} \\
 \frac{\partial Y_i}{\partial r} &= 0, (i = 1-3)
 \end{aligned}
 \tag{28}$$

Reactor inlets:

$$\begin{aligned}
 \vec{V} &= \vec{V}_{in} \\
 h &= h_{in} \\
 P &= P_{ref} \\
 Y &= Y_{in}, (i = 1-3)
 \end{aligned}
 \tag{29}$$

Mass diffusion is cut at gas inlets by setting diffusion coefficient as $D=0$.

Reactor outlet:

$$\begin{aligned}\frac{\partial u}{\partial x} &= 0 \\ \vec{V} &= 0 \\ \frac{\partial h}{\partial x} &= 0 \\ \frac{\partial^2 Y_i}{\partial x^2} &= 0, (i = 1-3)\end{aligned}\tag{30}$$

Symmetry axis:

$$\begin{aligned}\frac{\partial u}{\partial x} &= 0 \\ \vec{V} &= 0 \\ \frac{\partial h}{\partial r} &= 0 \\ \frac{\partial Y_i}{\partial r} &= 0, (i = 1-3)\end{aligned}\tag{31}$$

Other solid walls:

$$\begin{aligned}\vec{V} &= 0 \\ \frac{\partial h}{\partial r} &= 0 \\ \frac{\partial Y_i}{\partial r} &= 0, (i = 1-3)\end{aligned}\tag{32}$$

A.4. Discretization and solution of the governing equations

The Finite Volume Method was used to divide the solution domain into a number of cells known as control volumes. The first order up-wind differencing scheme (70%) blending with the 3rd-order differencing scheme (30%) is used to discrete variables in the

transport equations. For each of the solved variables and for the auxiliary variables, the under-relaxation method is used to maintain the stability of the coupled, nonlinear system of equations.

The grid number ($N_x \times N_r$) of the calculated domain is 191_63. The power law is used to dense the grid near the gas inlets and substrate surface. The calculation of the problem typically requires 3000 iterations (approximately 30 minutes CPU time using a Dell P4-1.7GHz computer). A finer grid size has been used and grid independent results have been obtained.

A.5. Results and Discussion

Fig. 24 (a) and (b) show the calculation domain and the grid system of the GaN growth reactor employed in this study.

Figure 25 shows the temperature distribution inside the reactor.

Figures 26 (a) and (b) show the velocity distribution and the velocity streamline distribution inside the reactor.

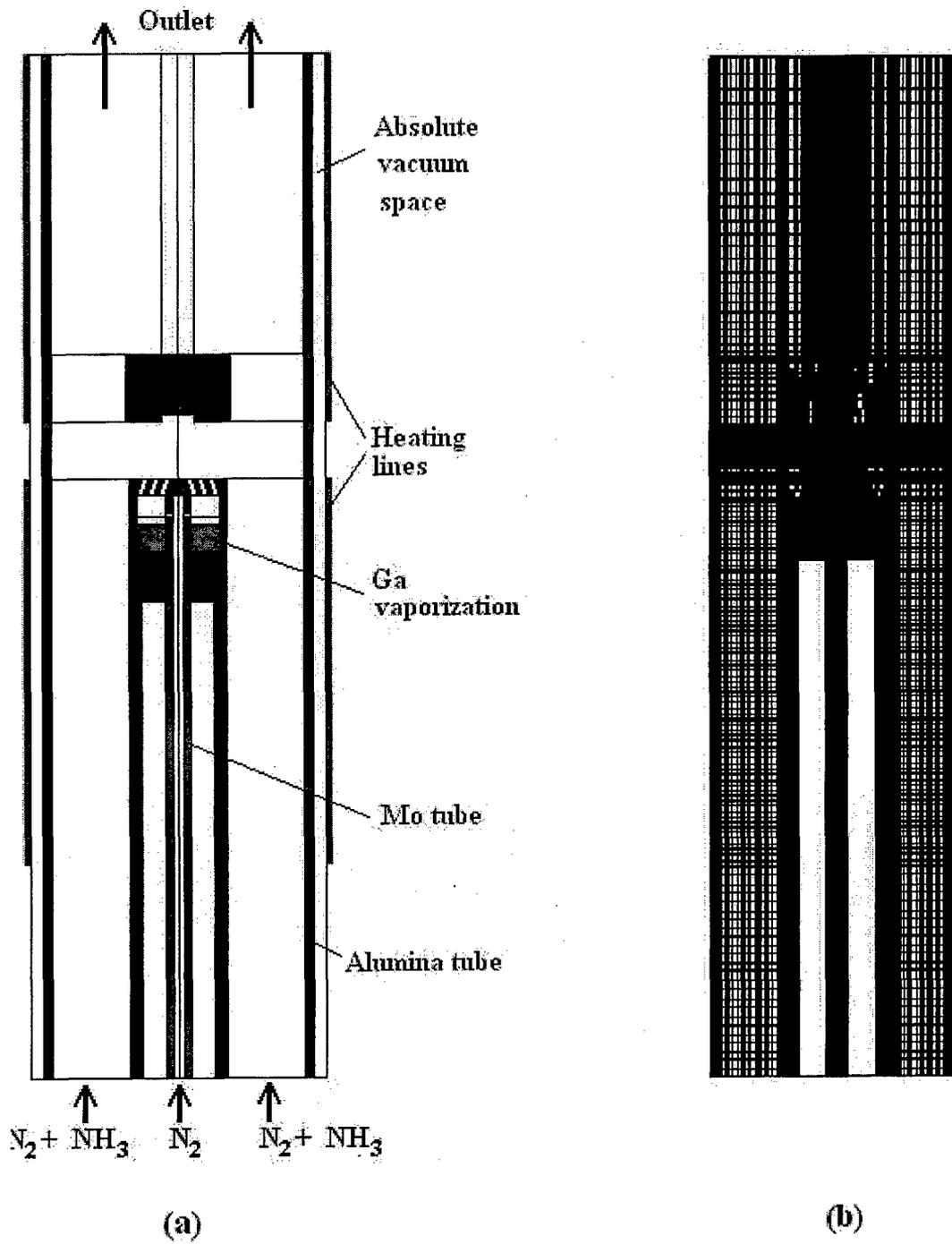


Figure 24. (a) Calculation geometry and (b) the grid system of the GaN growth reactor tube

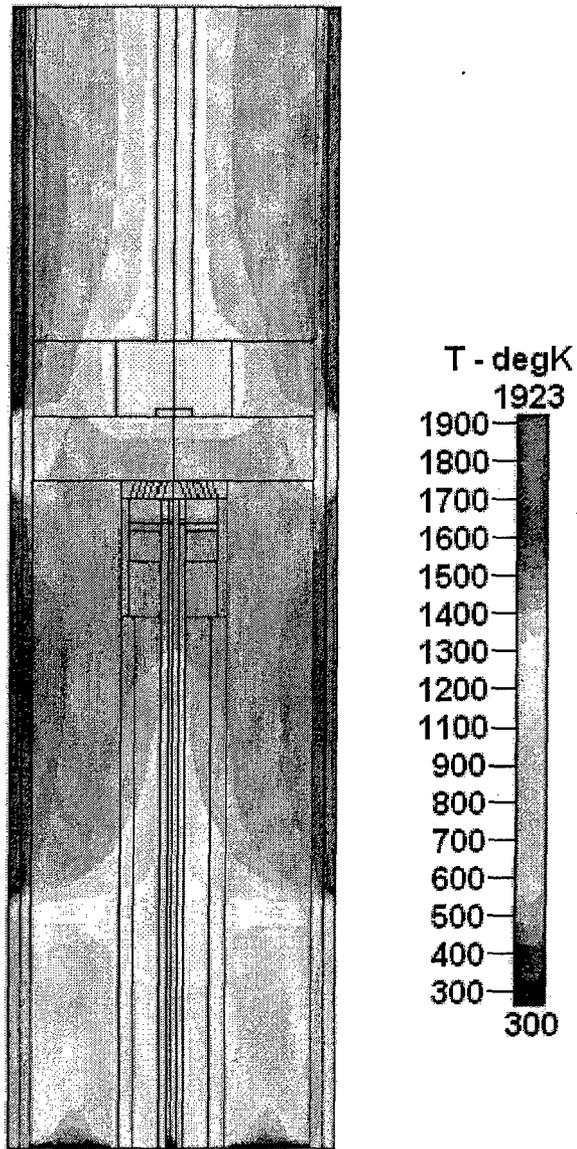
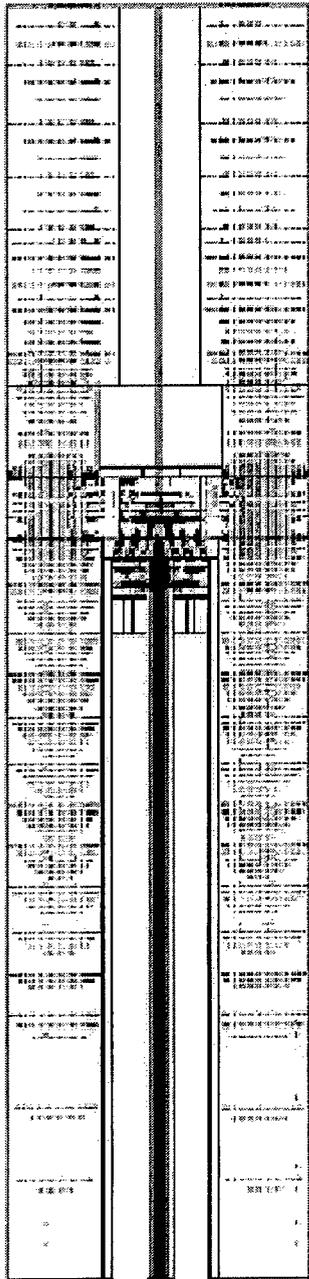
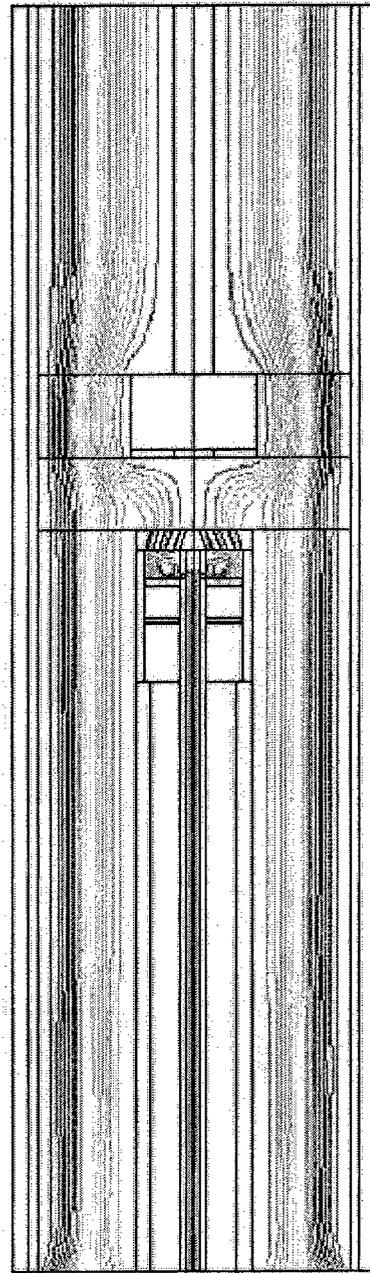


Figure 25. Temperature distribution in the reactor tube



(a)



(b)

Figure 26. (a) Velocity distribution and (b) Streamline distribution in the reactor tube

It can be seen in Fig. 25 that the achieved mixing gas temperature right below the substrate is about 1050°C (1323 K). The temperature on the Ga source surface is 1250°C (1523 K). It is observed that the temperature of the N₂ and NH₃ gases increases up to about 1350°C (1623K) before they reach the growth cell and mix with Ga vapor.

In the system there are three inlets and one outlet as shown in Fig. 23.

Inlet 1: N₂ + NH₃

Inlet 2: N₂

Inlet 3: Ga Vapor

For the streamline calculations, the reactor pressure was 800 Torr. The source-to-substrate distance was 24 mm. The Ga mass loss rate was 28 mg/h. The flow rates of NH₃ and N₂ at inlet 1 and N₂ at inlet 2 are listed in the Table 5. No reactions were considered. Fig. 26(b) shows that there was no reverse flow observed between the source and the substrate.

Table 5. Flow rates of NH₃ and N₂ at inlet 1 and N₂ at inlet 2

Case No.	Inlet 1 (slm)		N ₂ flow rate at inlet 2 (slm)
	NH ₃	N ₂	
1	3	1	1

APPENDIX B: REACTION ENERGETICS

Thermodynamics provides an upper limit of expected values under specified conditions; it implies that chemical equilibrium has been attained. Although this may occur in a closed system, it is generally not the case in an open or flow reactor where gaseous reactants and products are continuously introduced and removed. In general this type of systems may be treated empirically with the use of thermodynamic guidelines.

Thermodynamic analysis of potential vapor phase epitaxial processes provides an important estimate of the extent of departure from equilibrium or the supersaturation of the gas stream in the growth region under various conditions. Typically, deposition processes for coatings or similar applications operate with large departures from equilibrium in the deposition region in order to produce uniform coatings at maximum deposition rates. Epitaxial growth, on the other hand, must take place under relatively low gas phase supersaturations so that the reactant partial pressures do not deviate excessively from the equilibrium values. If the arrival rate of reactant species greatly exceeds the nucleation and lattice incorporation rates at the crystal surface, poor crystal perfection will result. High fluxes of reactants to the surface increase the probability of misoriented two-dimensional nucleation and, in severe cases, result in polycrystalline layers. In addition, excessive supersaturation may lead to three-dimensional nucleation in the gas phase with the resultant nuclei falling onto the substrate surface to produce misoriented regions within the layers.

Provided that the free-energy change, ΔG , can be approximated by the standard free-energy change, ΔG° , many simple consequences of thermodynamics with respect to

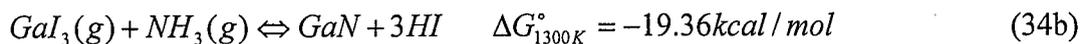
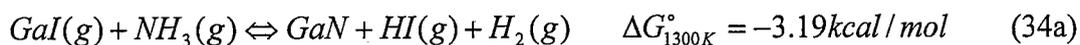
vapor phase deposition can be understood. For the epitaxial growth, it is essential that a single nucleus forms as an oriented seed for subsequent growth. According to elementary nucleation theory, a small negative value of ΔG , the chemical free energy per unit volume, is required to foster a low nucleation rate of large critical-sized nuclei. This in turn would require a ΔG value close to zero. When this happens, large amounts of reactants and products are simultaneously present. If ΔG were large and negative, however, the likelihood of a high rate of heterogeneous nucleation, or even homogeneous nucleation of solid particles within the gas phase, would be enhanced, as mentioned before. The large driving force for chemical reaction tends to promote polycrystal formation in this case.

The following is a comparison of free energy of the overall reaction involved in the GaN growth from Ga and NH_3 with that of other vapor phase epitaxy processes. Data used in the calculations were taken from Ref. [117].

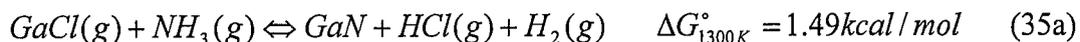
Formation of GaN from Ga and NH_3 :



Iodine-transport vapor phase epitaxy:



Chlorine-transport vapor phase epitaxy:





$$\Delta G_{1300K}^\circ = 96.45 \text{ kcal/mol} \quad (35b)$$



There are several different results for the chlorine-transport reaction in the literature and the accuracy of the results was discussed in detail by Newman [120].

Molecular Beam Epitaxy:

When NH_3 gas is used as the nitrogen source in MBE synthesis of GaN, the additional energy of Ga(g) is found to be important in driving the forward reaction [121, 122]. In this case, GaN is able to form in the unstable region of the GaN(s) and Ga(l)/ NH_3 (g) phase diagram. Fig. 27 illustrates the typical thin film growth conditions [120].

In the growth of GaN, the GaN decomposition reaction into Ga(l) and N_2 (g) may also play a role in limiting the growth parameters. Another competing reaction involves the decomposition of NH_3 . This turns out to be an important factor and is discussed in Appendix C.

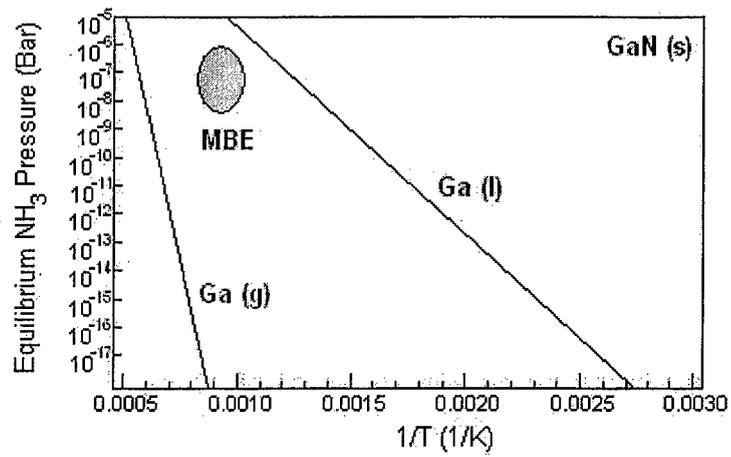
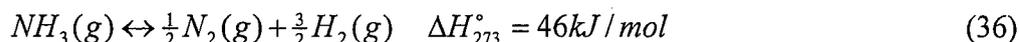


Figure 27. Change in Gibbs free energy and equilibrium pressure for the following reactions (1) $\text{Ga(l)} + \frac{1}{2} \text{N}_2(\text{g}) = \text{GaN(s)}$, (2) $\text{Ga(g)} + \frac{1}{2} \text{N}_2(\text{g}) = \text{GaN(s)}$

APPENDIX C: AMMONIA DECOMPOSITION

Despite the apparent simplicity of the GaN synthesis from elemental Ga and NH₃, the chemistry of the process is nontrivial due to one critical factor; ammonia is thermodynamically unstable at elevated temperature used in the growth. The ammonia decomposition reaction results in the production of N₂(g); a chemical specie which does not react with Ga(g) in the reactor environment. For this reason, the extent of ammonia decomposition in the reactor must be studied carefully to facilitate GaN synthesis.

Decomposition of ammonia is a mildly endothermic process, yielding hydrogen and nitrogen and it is accompanied by an increase in volume [79]:



The reaction is a reversible reaction, that is, it can proceed both in forward and backward directions. Fig. 28 illustrates the change in ΔG° and ΔH° with temperature [117]. The figure shows that ammonia is thermodynamically unstable and all the NH₃ is decomposed into N₂ and H₂ at temperatures higher than 300°C. However, it is well known that the decomposition rate is slow without a catalyst and the extent of decomposition strongly depends on the growth conditions and the equipment.

The rate of disassociation reaction is sluggish due to the kinetic barrier of breaking the N – H bond. Experimentally, Ban [123], Rocasecca et al. [124], Liu and Stevenson [125] also found that the decomposition rate increases with temperature, although with varying efficiency, depending on the catalyst. Tungsten, platinum and iron are commonly used for catalyzing the decomposition of ammonia [79]. The combination of Ga(l) or

GaN(s) is found to be surprisingly effective as a catalyst, although Ga(l) or GaN(s) alone were not found to increase the decomposition rate [120].

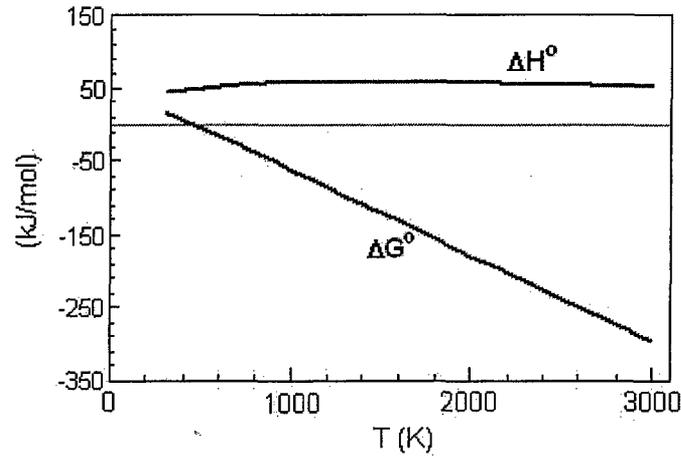


Figure 28. Gibbs free energy and enthalpy change versus temperature for ammonia decomposition reaction.

In the GaN growth chamber, alumina tube walls and Mo-parts act as catalysts for the NH_3 decomposition. These parts interact with NH_3 in a cyclic manner promoting perhaps many reactions at the molecular level. They change the activation energy, E_a , of decomposition by providing an alternate pathway for the reaction. The rate constant k of a reaction is related to E_a in the following way

$$k = Ae^{\left(\frac{-E_a}{RT}\right)} \quad (37)$$

where A is a constant related to collision rates and the rate is equal to $k \times$ function of concentration. Thus, a change in E_a changes the rate of a reaction. In this case, alumina tube walls and Mo-parts are said to be heterogeneous, or contact, catalysts because they are in a separate phase from the gas.

The interaction of ammonia with different material surfaces has been thoroughly investigated [126-135]. These studies have been undertaken primarily to acquire insights into the industrially important ammonia synthesis process. Figure 29 shows the energy profiles for ammonia synthesis in absence and presence of the iron catalyst [136]. The arrows in the diagram show the activation energies of each of the steps. The exact values vary depending on the catalyst.

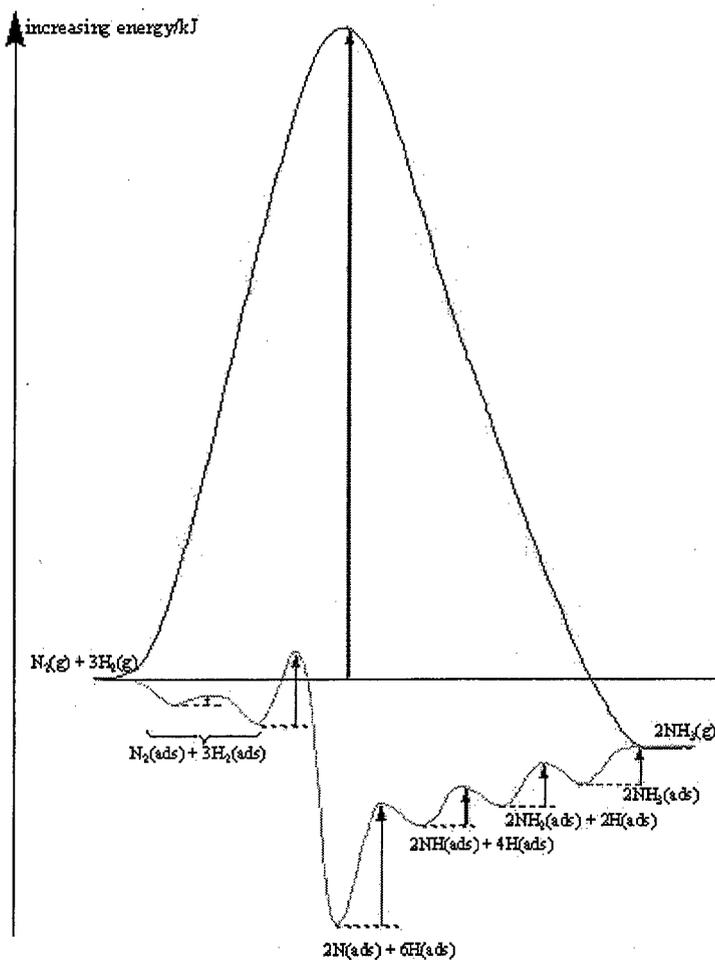
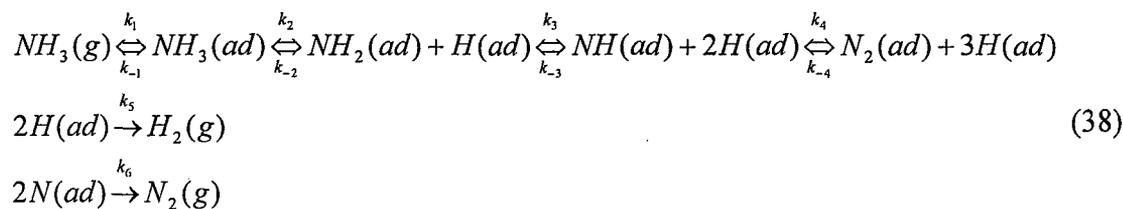


Figure 29. Energy changes in ammonia decomposition reaction

The reaction of NH_3 decomposition on a catalytic surface can be expressed by the following sequence of molecular steps [79].



The rates k_2 and k_3 are different on the individual surfaces. The effect of several metallic catalysts on the activation energy for the decomposition of ammonia can be seen in the Table 6.

Table 6. Activation energy of ammonia decomposition reaction on different surfaces

Catalyst	E_a (kJ/mol)
None ^[137]	335
Tungsten ^[137]	163
Alumina (Al_2O_3) ^[138]	127
Molybdenum ^[139]	116

Data are from different references where the conditions are slightly different

Ertl and Huber [119] studied the steady state decomposition reaction at $P_{\text{NH}_3} = 5 \times 10^{-6} \text{ Torr}$ on an iron (Fe) wire under flow conditions with a high pumping speed to keep the partial pressure of N_2 and H_2 so small that the reverse reaction could be neglected. The activation energy, E_a was determined from the Arrhenius plot for the rate of stationary N_2 production. Up to temperatures of 640 K, E_a was determined as 188 kJ/mole for elevated temperatures, it was found that E_a decreases continuously and the rate becomes nearly temperature independent ($E_a=0$) at around 850 K. The reaction order

with respect to P_{NH_3} at low temperatures is $n = 0$ and increases with temperature to $n = 1$.

If the temperature is high enough and/or the NH_3 pressure small, the delivery of NH_3 from the gas phase will be limiting. With the reaction sequence (38), the rate of decomposition is described by [79]

$$rate = \frac{k_1 k_2 P_{NH_3}}{(k_{-1} + k_2)} \quad (39)$$

Since $k_2 \gg k_{-1}$ [83] the rate is given by

$$rate = k_1 P_{NH_3} \quad (40)$$

The experimental results [119] support these conclusions. At higher temperatures, the activation energy decreases to $E_a = 0$ and the order is $n = 1$ with respect to P_{NH_3} .

The rate of ammonia decomposition at elevated temperatures and under conditions far from thermodynamic equilibrium (where the ammonia formation can be neglected) can be simply described by

$$rate = -\frac{d[NH_3]}{dt} = k P_{NH_3} \quad (41)$$

If the reactions in both forward and backward directions are considered, rate expressions become

$$\begin{aligned} \frac{d[NH_3]}{dt} &= -k_F P_{NH_3} + k_B P_{H_2}^{3/2} P_{N_2}^{1/2} \\ \frac{d[H_2]}{dt} &= -k_B P_{H_2}^{3/2} P_{N_2}^{1/2} + k_F P_{NH_3} \\ \frac{d[N_2]}{dt} &= \frac{d[H_2]}{dt} \end{aligned} \quad (42)$$

where k_F is the rate constant of forward reaction in ammonia decomposition reaction (36), and k_B is the rate constant of the backward reaction. At an equilibrium condition we would expect the rate to be equal to

$$\frac{d[NH_3]}{dt} = 0 = -k_F P_{NH_3} + k_B P_{H_2}^{3/2} P_{N_2}^{1/2} \quad (43)$$

$$\text{or } \frac{k_F}{k_B} = \frac{P_{H_2}^{3/2} P_{N_2}^{1/2}}{P_{NH_3}} \quad (44)$$

This is the equilibrium constant K_{eq} of the reaction (36), and this sets the limit to the decomposition process.

$$K_{eq} = \frac{P_{H_2}^{3/2} P_{N_2}^{1/2}}{P_{NH_3}} \quad (45)$$

From the tabulated data [117], the equilibrium NH_3 concentrations can be calculated. Assuming the input gas is only ammonia, at 1500 K and a $N_2:H_2$ ratio of 1:3, a pressure of 200 Torr would yield an equilibrium NH_3 pressure of 2.4×10^{-3} Torr.

Using the simple rate description, Eq.(41), NH_3 decomposition rate in the GaN growth reaction tube was estimated roughly, in order to have an idea on the extend of decomposition and to understand the simulation results better.

Below set of equations shows how the distance along the tube was included in the equations.

$$\frac{dn_{NH_3}}{dt} = -k(n_{NH_3})_i \quad (46)$$

$$\frac{dn_{NH_3}}{dx} \frac{dx}{\underbrace{dt}_{\text{gas velocity, } v}} = -k(n_{NH_3})_i \quad (47)$$

$$\frac{dn_{NH_3}}{(n_{NH_3})_i} = -\frac{1}{v} k dx \quad (48)$$

$$\frac{n_{NH_3}}{(n_{NH_3})_i} = e^{-kx/v} \quad (49)$$

where x is the distance along the tube, n_{NH_3} is the ammonia concentration. Eq. (49) was calculated and plotted as shown in Fig. 30. In calculations, a gas velocity of 5 cm/sec and a temperature of 1250°C were used. Fig. 30 shows the extent of ammonia decomposition in the reactor where the majority of ammonia decomposes after traveling a few millimeters along the tube in the hot zone.

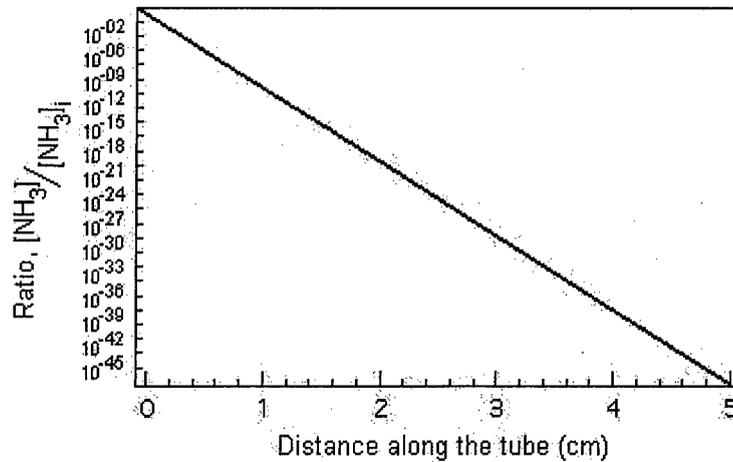


Figure 30. Ratio of ammonia concentration to the initial concentration versus distance.

REPORT DOCUMENTATION PAGE

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14. ABSTRACT GaN thin films were grown on (0001) sapphire substrates at 1050°C by controlled evaporation of Ga metal and reaction with NH ₃ . Transport of Ga species from source-to-substrate during this process and the influence of ammonia-liquid Ga reaction on Ga transport were investigated. Experimental results were compared to theoretical predictions to quantify the mechanism of transport in the vapor growth technique. In presence of NH ₃ , Ga transport far exceeded the predicted upper limit for the vapor phase transport. Visual observations confirmed that a significant amount of Ga left the source in a cluster rather than atomic form. Pure N ₂ was flowed directly above the molten Ga source. This prevented direct contact and reaction between the molten Ga and NH ₃ and prevented Ga spattering and GaN crust formation on the source surface. It enhanced Ga evaporation rate and enabled control of Ga transport and V/III ratio in the system. The results showed that the process was mass transport limited and the maximum growth rate was controlled by transport of both Ga and reactive ammonia species to the substrate surface.					
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