Water stabilization of Zr₆-based metal–organic frameworks via solvent-assisted ligand incorporation†

Pravas Deria,‡ Yongchul G. Chung, Randall Q. Snurr, Joseph T. Hupp* and Omar K. Farha*ab

Water stability in metal–organic frameworks (MOFs) is critical for several practical applications. While water instability is mainly thought to stem from linker hydrolysis, MOFs with strong, hydrolysis-resistant metal-linker bonds can be susceptible to damage by capillary forces, which cause cavities and channels to collapse during activation from water. This study utilizes metal node functionalization as a strategy to create vapor-stable and recyclable MOFs.

Introduction

Porous metal–organic frameworks (MOFs) are hybrid solid-state compounds¹ that have attracted voluminous research efforts for applications relevant to gas storage and separation,² chemical catalysis,³ and optoelectronic⁴ and thermoelectric devices.⁵ However, many of these applications require recyclability, which is intimately tied to the stability of the corresponding frameworks in ambient conditions as well as in the presence of water.⁶ It is worth noting that only a modest fraction of the total known MOFs are both thermally and chemically stable. Stable MOFs include those derived from nodes with high valent metal ions⁷ and some imidazolate, pyrazolate, and multitopic carboxylate linkers. Among the more remarkable examples are ZIFs,⁷ pyrazolate-based frameworks,⁸ MILs,⁹ and ZrIV-based UiOs.⁷,10 While the microporous UiO-66 framework, consisting of 12-connected [Zr₆(OSi₃)₄(OH)₄]₄⁺ nodes linked through 1,4-benzenedicarboxylates (BDC),⁷ was established to be mechanically,⁷,11 thermally,⁷ hydrolytically and chemically stable,⁷,11 other Zr₆-based MOFs with larger pores, such as UiO-67, constructed with a longer linker (BPDC, 4,4’-biphenylDicarboxylate), have often been found to lose porosity upon exposure and activation from water.¹²,13 We recently showed that the instability of these Zr₆-based MOFs is not due to hydrolysis, which is an energetically disfavored process (ΔG = +26 kcal mol⁻¹ with ΔG‡ = +38 kcal mol⁻¹) owing to a strong ZrIV-carboxylate connectivity, but instead stems from capillary-force-driven channel collapse during removal of water.¹² The water-exposed MOF sample, when exchanged with a solvent exerting less capillary force, was shown to retain permanent porosity by thermal activation under dynamic vacuum.¹³,14 Some applications of MOFs, however, may not be compatible with stepwise solvent evacuation.⁶,¹³,¹⁶ The availability of an alternative stabilization method clearly would be desirable.

Given the breadth of applications specifically of Zr₆-based MOFs to condensed-phase chemistry,¹⁷ we have focused on understanding the origin of such capillary-force-driven collapse, while devising a strategy that would enable MOFs to be recycled (i.e. repetitively solvent evacuated) via simple thermal activation. The Zr₆-based mesoporous MOFs are intriguing, particularly those featuring eight-connected [Zr₆(OSi₃)₄(OH)₄(–OH)]₁⁺ nodes linked by tetra-carboxylate ligands (Fig. 1a) giving rise to framework structures with hydroxyl and aqua ligands protruding into mesoporous channels. These non-bonding hydroxyl and aqua ligands provide polar and tunable primary interaction points or ‘anchors’ for water clusters to attach and are, thus, ideal for studying the role of polar nodes in capillary-force-driven pore collapse.

We have established that the non-bonding hydroxyl and aqua ligands of the Zr₆-nodes of the mesoporous MOF NU-1000 [molecular formula Zr₆(µ₃-O)₄(µ₁-OH)₆(–OH)₄(–OH)₂(–OH)₂(TBAPy)₂; Fig. 1a; H₂TBAPy is the 1,3,6,8-tetakis[p-hydroxy-2H-pyrene]²⁻] can facilitate a variety of node-functionalization schemes including metalation via AIM (Atomic layer deposition In a Metal–organic framework)²⁷ and grafting organic chemical functionalities via SALI (Solvent-Assisted Ligand Incorporation; Fig. 1b).¹⁸,¹⁹ These modifications not only facilitate various chemical processes,¹⁹,²⁰ but also potentially provide a means to control the pore diameter and the accessibility of the nodes by condensed water. For example, a SALI-derived NU-1000 framework possesses a node coordination similar to that established in the UiO-66 structure,¹⁹ i.e. a [Zr₆(µ₃-O)₄(µ₁-OH)]₁²⁺ node
with twelve carboxylate ligands without any terminal hydroxyl or aqua ligand (Fig. 1b). Recent literature reports have established that in microporous UiO-66 and isostructural MOF-801 (with fumarate as the linker),\(^{5,6}\) the bridging hydroxo ligands on the Zr\(_6\)-oxo node play a key role as a primary interaction site for the guest water molecules, which form a hydrogen bonding network that facilitates further uptake of water molecules. Since SALI-decoration in NU-1000 replaces the terminal hydroxo and aqua ligands at the Zr\(_6\)-oxo nodes, the incoming carboxylate functional groups (CFGs) of various size not only reduce the pore volume but also potentially produce a ‘protected’ node that interacts less strongly with water molecules (Fig. 1b).\(^{5,6,21}\)

**Results and discussion**

Given that NU-1000 undergoes capillary-force-driven pore collapse when activated (thermal evacuation) from its liquid-water-filled form,\(^{13,14}\) we wanted to test its stability under water vapor conditions at various vapor partial pressures (or relative humidities, RH) using N\(_2\) as carrier gas at a constant total pressure of 1 bar.\(^{22}\) As can be seen in Fig. 2a, at RH \(\approx 70\%\), the mesoporous NU-1000 channels undergo capillary condensation and reach saturation with 1.14 mg mg\(^{-1}\) total uptake (corresponding to a water accessible pore volume of 1.14 cc g\(^{-1}\), where the N\(_2\) accessible pore volume \(= 1.4\) cc g\(^{-1}\)).\(^{23}\) Removal of the condensed water, during the desorption cycle, was achieved by progressive lowering of the vapor concentration or partial pressure with dry N\(_2\) carrier gas and occurs with considerable hysteresis. However, the second adsorption isotherm, under identical conditions, evinces only a 0.43 mg mg\(^{-1}\) saturation uptake. The decrease in the total uptake indicates a decrease in the porosity, which possibly commenced during the first desorption cycle. Likewise, the third adsorption shows 0.31 mg mg\(^{-1}\) of total uptake with a further decrease of the porosity. The powder X-ray diffraction (PXRD) pattern of the recovered sample revealed a lower crystallinity (Fig. 2c), as the diffraction peaks centered at 2.55, 5.1 and 7.4 degrees (2\(\theta\)) are significantly broadened and the peaks at higher angles less intense. Consistent with these findings, the N\(_2\) isotherm of the recovered (after the third cycle), thermally activated sample shows significant reductions of both the BET surface area (320 m\(^2\) g\(^{-1}\)) and the N\(_2\) accessible total pore volume down (0.15 cc g\(^{-1}\)) (Fig. 3a and Table 1).

To gain insight into the role of the polar hydroxyl and aqua ligands, we compared the results for NU-1000 with those obtained benzoate functionalized NU-1000 (termed SALI-BA), where the accessible terminal hydroxyl and aqua ligands are replaced by benzoate ligands (Fig. 1b). Similar to NU-1000, SALI-BA is subject to capillary condensation at RH \(\approx 70\%\) and reaches a water saturation uptake of 1.0 mg mg\(^{-1}\), which tracks with the N\(_2\) accessible pore volume (1.21 cc g\(^{-1}\); Fig. 2b; Table 1). SALI-BA exhibited a diminished uptake during the second desorption. However, the corresponding reduction of the pore volume is less severe (~44\%) than that observed for the unfunctionalized NU-1000 (~62\%). The magnitude of the pore volume reduction suggests that a benzoate functionalized node (1) provides a less effective ‘anchoring’ site and (2) engenders lower pore volume for the capillary-condensed water, rendering a lower degree of damage during the desorption. The corresponding PXRD pattern and the N\(_2\) isotherm of the recovered (after third cycle), thermally activated sample indicated better retention of crystallinity and porosity relative to the
unfunctionalized compound (see Fig. 2c and 3a). DRIFTS data of the recovered SALI-BA sample show retention of the characteristic bridging hydroxyl peak at 3671 cm⁻¹ whereas the recovered unfunctionalized NU-1000 displays a broad band at the expense of the sharp peaks at 3674 and 3671 cm⁻¹ associated with the terminal and bridging hydroxyl and the terminal aqua ligands, respectively (Fig. 2d). These data suggest node structure alteration and possibly a partially pore-collapsed framework. We now turn our attention to SALI-CFG materials featuring flexible functional groups that may protect the node further by inhibiting access of condensed water to the MOF node. Perfluorodecanoic-acid-functionalized NU-1000 (termed SALI-9) is an example of a highly-stable mesoporous MOF with a reduced pore volume, rendering a total water vapor uptake that is ~40% as great as that of unfunctionalized NU-1000 (Fig. 4, ESI-4a; Table 1). Remarkably, over the course of 20 cycles of water vapor sorption/desorption, the porosity and crystallinity of SALI-9 remain essentially unchanged. It is worth noting that the N₂-accessible pore volume of SALI-9 sample is measured to be 0.63 cc g⁻¹ i.e. somewhat lower than that of the parent material and consistent with the volumetric demands of the perfluoroalkane chains themselves. Even though functionalization significantly reduces the MOF pore volume, it is interesting to note that water condensation,³ in SALI-9, commences at essentially the same vapor pressure as for NU-1000 (Fig. 2). Corresponding contact angle measurements (Fig. ESI-7f), a bulk measurement of external-surface interactions with liquid water, indicate enhanced hydrophobicity for SALI-9 relative to unfunctionalized NU-1000. The recovered sample (after 20 cycles) after thermal activation reveals essentially no degradation of the crystallinity, porosity and Zr₆-oxo node structure (see Fig. 3, ESI-5 and 6f).

To understand if the acquired stability is due mainly to the hydrophobic nature of the perfluoroalkane decoration or to concurrent reduction of the pore size and, therefore, water cluster size, we studied the corresponding CF₃H₉ and the shorter CF₃F₁₁ chains in NU-1000 (SALI-9' and SALI-5, respectively). Due to its lower molar mass compared to the corresponding CF₃H₉ variant, SALI-9' shows higher gravimetric vapor uptake (ca. 68% of that of the unfunctionalize MOF; Fig. 4, ESI-4f). The stability of SALI-9' to repetitive infilling and removal of water is nearly as good as that of SALI-9, as can be seen from a modest 10% decrease in saturation water uptake after 20 cycles of vapor sorption and desorption.²⁴ The N₂ isotherm, PXRD, and DRIFTS data (Fig. 3 and ESI 5 and 6f) of the recovered thermally activated sample reveal essentially complete retention of the crystallinity, porosity, and node coordination environment. Likewise, the shorter perfluoroalkane variant, SALI-5, shows excellent performance: significantly higher water uptake (~63% of NU-1000) along with stability (Fig. 3 and 4; Table 1) similar to that of SALI-9.

Given that the pyrene macrocycle of the NU-1000 linker is hydrophobic, these experimental results led us to hypothesize that within the mesoporous MOF channel, the polar [Zr₆(OH)₄(μ₃-O)₆(μ₁-OH)₄(μ-OH)₄(μ-OH)₄(μ-OH)₄(μ-OH)₄] node strongly interacts

Table 1 BET surface areas, pore volumes, multi-cycle water uptake for NU-1000 and SALI-CFG samples

<table>
<thead>
<tr>
<th>MOF</th>
<th>Ligand</th>
<th>BET surface area (m² g⁻¹)</th>
<th>Pore volume (cc g⁻¹)</th>
<th>Multi-cycle saturation H₂O uptake (mg mg⁻¹)</th>
<th>BET surface area (m² g⁻¹)</th>
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<tr>
<td>NU-1000</td>
<td>OH⁻, H₂O</td>
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<td>1.14</td>
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<td>SALI-BA</td>
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<tr>
<td>SALI-5</td>
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<td>0.72</td>
<td>0.72</td>
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<td>SALI-9</td>
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<td>900</td>
<td>0.63</td>
<td>0.44</td>
<td>0.43</td>
</tr>
<tr>
<td>SALI-9'</td>
<td>CH₃(CH₂)₈CO₂⁻</td>
<td>1190</td>
<td>0.87</td>
<td>0.78</td>
<td>0.73</td>
</tr>
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</table>
with the condensed water. Thus during the desorption cycle, condensed water can exert significant capillary force, which in turn results in pore collapse. In the SALI-CFG samples, the terminal hydroxyl and aqua ligands were replaced with aromatic or alkyl carboxylates that limit the accessibility of the polar Zr₆-oxo nodes by water. Effective protection is tied to both the size and hydrophobicity of the CFG introduced by the SALI reaction.

To gain molecular insight into this hypothesis, we performed molecular dynamics (MD) simulations for water in NU-1000, SALI-BA and SALI-5. (see ESI section 7 for details.) The simulations predict that the number of water molecules sited within 15 Å of the Zr₆-nodes varies with MOF identity in the order NU-1000 > SALI-BA > SALI-5 (Table S4†). Likewise the density of water in the hexagonal channels (which could affect the capillary force) is ranked as NU-1000 > SALI-BA > SALI-5 (Table S3†). These results support our hypothesis that the improved stability in SALI-derived system is due to limited accessibility of water molecules near the Zr₆-nodes.

Conclusions

In conclusion, our results suggest that strong interactions of condensed water with the polar nodes within the framework of NU-1000 can exert strain during removal of the water from the pores. While replacement of the polar hydroxyl and aqua ligands with aprotic, non-polar organic carboxylates can reduce the strength of the interaction of water with the node, a flexible alkyl chain can further limit the accessibility by water vapor. MD simulations revealed that a longer and flexible carboxylate-based ligand indeed provides more limited access of water molecules to the Zr₆ nodes. Node functionalization, by methods such as SALI, can be an effective strategy for engendering MOF stability toward water removal and, therefore, MOF recyclability from liquid water without loss of permanent porosity.

Acknowledgements

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Notes and references


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Pore volumes were calculated from the saturation uptake (at 1013 Pa) of the recorded sorption isotherms following the BET model (Brunauer-Emmett-Teller). The micropore contribution was calculated with the t-plot method [28, 29]. The average particle size was estimated by applying the Sherrer formula, where the Bragg’s angle was calculated from the PXRD and N2-adsorption peaks.

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